



education

Department of
Education
FREE STATE PROVINCE

PHYSICAL SCIENCES
PAPER 2 (CHEMISTRY)
GRADE 12
TERMS & DEFINITIONS
QUESTIONS & ANSWERS
PER TOPIC

2017



This document consists of 114 pages.

CONTENTS

HOW TO USE THIS DOCUMENT	3
TERMS AND DEFINITIONS	4
QUESTIONS	12
ORGANIC MOLECULES: NOMENCLATURE.....	12
ORGANIC MOLECULES: PHYSICAL PROPERTIES.....	19
ORGANIC MOLECULES: ORGANIC REACTIONS.....	23
REACTION RATE AND ENERGY IN CHEMICAL REACTIONS.....	29
CHEMICAL EQUILIBRIUM.....	36
ACIDS AND BASES.....	42
GALVANIC CELLS.....	47
ELECTROLYTIC CELLS.....	53
FERTILISERS.....	58
ANSWERS TO QUESTIONS	63
ORGANIC MOLECULES: NOMENCLATURE.....	63
ORGANIC MOLECULES: PHYSICAL PROPERTIES.....	69
ORGANIC MOLECULES: ORGANIC REACTIONS.....	73
REACTION RATE AND ENERGY IN CHEMICAL REACTIONS.....	77
CHEMICAL EQUILIBRIUM.....	85
ACIDS AND BASES.....	95
GALVANIC CELLS.....	104
ELECTROLYTIC CELLS.....	108
FERTILISERS.....	111
BIBLIOGRAPHY	114



HOW TO USE THIS DOCUMENT


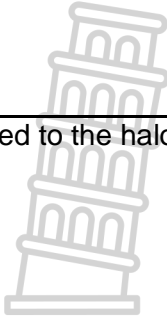
Dear Grade 12 learner

1. This document was compiled as an extra resource to help you to perform well in Physical Sciences.
2. Firstly you must make sure that you study the terms and definitions provided for each topic. Theory always forms part of any test or examination and you should ensure that you obtain full marks for ALL theory questions. Always be prepared to write a test on terms and definitions as soon as a topic is completed in class. Revise terms and definitions of topics already completed frequently so that you know them by the time you are sitting for a test or an examination.
3. Answer all the questions on a certain topic in your homework book as soon as the topic is completed. DO NOT look at the answers before attempting the questions. First try it yourself. Compare your answers with the answers at the back of the document. Mark your work with a pencil and do corrections for your incorrect answers. If you do not know how to answer a question, the answers are there to guide you. Acquaint yourself with the way in which a particular type of question should be answered. Answers supplied are from memoranda used to mark the questions in previous years.
4. Your teacher can, for example, give you two of the questions in this document as homework. The following day he/she will just check whether you answered them and whether you marked your answers. The teacher will only discuss those questions in which you do not understand the answers supplied in the document. Therefore a lot of time will be saved.
5. You are probably thinking about the point behind the answers at the back of the document. It is intended to help you to prepare for your tests and examinations. If you choose to copy answers into your homework book without trying them out yourself, you will be the losing party in the end! Not your teacher or anybody else!
6. Your teacher can also decide to give you a test on one of the questions given for homework. If you just copied the answers without any understanding, surely he/she will catch you! None of us want to be branded as dishonest, do we?
7. Work through all the questions and answers of a particular topic before you sit for an examination, even if you answered the questions before.
8. Any additional resource is only of help when used correctly. Ensure that you make use of all help provided in the correct way to enable you to be successful. All the best for 2017 and may you perform very well in Physical Sciences.

TERMS AND DEFINITIONS

MATTER AND MATERIALS: ORGANIC MOLECULES	
Addition reaction	A reaction in which a double bond in the starting material is broken and elements are added to it.
Addition polymer	A polymer formed when monomers (usually containing a double bond) combine through an addition reaction.
Addition polymerisation	A reaction in which small molecules join to form very large molecules by adding on double bonds.
Alcohol	An organic compound in which H atoms in an alkane have been substituted with hydroxyl groups (-OH groups). General formula: $C_nH_{2n+1}OH$
Aldehydes	Organic compounds having the general structure RCHO where R = H or alkyl. General formula: RCHO (R = alkyl group)
Alkane	An organic compound containing only C-H and C-C single bonds. General formula: C_nH_{2n+2}
Alkene	A compound of carbon and hydrogen that contains a carbon-carbon double bond. General formula: C_nH_{2n}
Alkyl group	A group formed by removing one H atom from an alkane.
Alkyne	A compound of carbon and hydrogen that contains a carbon-carbon triple bond.
Boiling point	The temperature at which the vapour pressure of a liquid equals atmospheric pressure.
Carbonyl group	Functional group of ketones ($>C=O$)
Carboxyl group	Functional group of carboxylic acids (-COOH)
Carboxylic acid	An organic compound containing a carboxyl group (-COOH group). General formula: $C_nH_{2n+1}COOH$ (or RCOOH)
Chain isomers	Compounds with the same molecular formula, but different types of chains.
Condensation polymer	A polymer formed by two monomers with different functional groups that are linked together in a condensation reaction in which a small molecule, usually water, is lost.
Condensation polymerisation	Molecules of two monomers with different functional groups undergo condensation reactions with the loss of small molecules, usually water
Condensed structural formula	A formula that shows the way in which atoms are bonded together in the molecule, but DOES NOT SHOW ALL bond lines.
Cracking	The chemical process in which longer chain hydrocarbon molecules are broken down to shorter more useful molecules.
Dehydration	Elimination of water from a compound usually such as an alcohol.
Dehydrohalogenation	The elimination of hydrogen and a halogen from a haloalkane.
Dipole-dipole force	Intermolecular forces found between polar molecules i.e. molecules in which there is an uneven distribution of charge so that the molecule has a positive and a negative side.
Elimination reaction	A reaction in which elements of the starting material are "lost" and a double bond is formed.
Esterification	The preparation of an ester from the reaction of a carboxylic acid with an alcohol.
Functional group	A bond or an atom or a group of atoms that determine(s) the physical and chemical properties of a group of organic compounds.
Functional isomers	Compounds with the same molecular formula, but different functional groups.
Haloalkane (Alkyl halide)	An organic compound in which one or more H atoms in an alkane have been replaced with halogen atoms. General formula: $C_nH_{2n+1}X$ (X = F, Cl, Br or I)

Halogenation	The reaction of a halogen (Br_2 , Cl_2) with a compound.
Homologous series	A series of organic compounds that can be described by the same general formula. OR A series of organic compounds in which one member differs from the next with a CH_2 group.
Hydration	The addition of water to a compound.
Hydrocarbon	Organic compounds that consist of hydrogen and carbon only.
Hydrogenation	The addition of hydrogen to an alkene
Hydrogen bond	A strong intermolecular force found between molecules in which an H atom is covalently bonded to either an N, O or F atom.
Hydrohalogenation	The addition of a hydrogen halide to an alkene.
Hydrolysis	The reaction of a compound with water.
Intermolecular force	Forces between molecules that determine physical properties of compounds.
IUPAC name	A chemical nomenclature (set of rules) created and developed by the International Union of Pure and Applied Chemistry (IUPAC) to generate systematic names for chemical compounds.
London force	A weak intermolecular force between non-polar molecules.
Macromolecule	A molecule that consists of a large number of atoms.
Melting point	The temperature at which the solid and liquid phases of a substance are at equilibrium.
Molecular formula	A chemical formula that indicates the type of atoms and the correct number of each in a molecule, e.g. CH_4 .
Monomer	Small organic molecules that can be covalently bonded to each other in a repeating pattern.
Organic chemistry	Chemistry of carbon compounds.
Polymer	A large molecule composed of smaller monomer units covalently bonded to each other in a repeating pattern
Polymerisation	A chemical reaction in which monomer molecules join to form a polymer
Positional isomer	Compounds with the same molecular formula, but different positions of the side chain, substituents or functional groups on the parent chain.
Primary alcohol	One C atom is bonded to the C atom bonded to hydroxyl group. Example: $\begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H}-\text{C}-\text{C}-\text{O}-\text{H} \\ \quad \\ \text{H} \quad \text{H} \end{array}$
Primary haloalkane	One C atom is bonded to the C atom bonded to the halogen. Example: $\begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H}-\text{C}-\text{C}-\text{Br} \\ \quad \\ \text{H} \quad \text{H} \end{array}$
Saturated compounds	Compounds in which there are no multiple bonds between C atoms in their hydrocarbon chains. OR Compounds with only single bonds between C atoms in their hydrocarbon chains.

<p>Secondary alcohol</p> 	<p>Two C atoms bonded to the C atom that is bonded to hydroxyl group. Example:</p> $ \begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H}-\text{C}-\text{C}-\text{O}-\text{H} \\ \quad \\ \text{H} \quad \text{H}-\text{C}-\text{H} \\ \quad \quad \\ \quad \quad \text{H} \end{array} $
<p>Secondary haloalkane</p>	<p>Two C atoms bonded to the C atom bonded to the halogen. Example:</p> $ \begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H}-\text{C}-\text{C}-\text{Br} \\ \quad \\ \text{H} \quad \text{H}-\text{C}-\text{H} \\ \quad \quad \\ \quad \quad \text{H} \end{array} $
<p>Structural formula</p>	<p>A structural formula of a compound shows which atoms are attached to which within the molecule. Atoms are represented by their chemical symbols and lines are used to represent ALL the bonds that hold the atoms together.</p>
<p>Structural isomer</p>	<p>Organic molecules with the same molecular formula, but different structural formulae.</p>
<p>Substituent (branch)</p>	<p>A group or branch attached to the longest continuous chain of C atoms in an organic compound.</p>
<p>Substitution reaction</p>	<p>A reaction in which an atom or a group of atoms in a molecule is replaced by another atom or group of atoms.</p>
<p>Tertiary alcohol</p>	<p>Three C atoms bonded to carbon that is bonded to hydroxyl group. Example:</p> $ \begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}-\text{H} \\ \\ \text{H} \quad \text{H} \\ \quad \\ \text{H}-\text{C}-\text{C}-\text{O}-\text{H} \\ \quad \\ \text{H} \quad \text{H}-\text{C}-\text{H} \\ \quad \quad \\ \quad \quad \text{H} \end{array} $
<p>Tertiary haloalkane</p>	<p>Three C atoms bonded to the C atom bonded to the halogen. Example:</p> $ \begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}-\text{H} \\ \\ \text{H} \quad \text{H} \\ \quad \\ \text{H}-\text{C}-\text{C}-\text{Br} \\ \quad \\ \text{H} \quad \text{H}-\text{C}-\text{H} \\ \quad \quad \\ \quad \quad \text{H} \end{array} $ 
<p>Unsaturated compounds</p>	<p>Compounds in which there are multiple bonds (double or triple bonds) between C atoms in their hydrocarbon chains.</p>
<p>Van der Waals forces</p>	<p>A combined name used for the different types of intermolecular forces.</p>
<p>Vapour pressure</p>	<p>The pressure exerted by a vapour at equilibrium with its liquid in a closed system.</p>

CHEMICAL CHANGE: QUANTITATIVE ASPECTS OF CHEMICAL CHANGE

Mole	One mole of a substance is the amount of substance having the same number of particles as there are atoms in 12 g carbon-12.
Molar gas volume at STP	The volume of one mole of a gas. (1 mole of any gas occupies 22,4 dm ³ at 0 °C (273 K) and 1 atmosphere (101,3 kPa).
Molar mass	The mass of one mole of a substance. Symbol: M Unit: g·mol ⁻¹
Avogadro's Law	Under the same conditions of temperature and pressure, the same number of moles of all gases occupy the same volume.
Concentration	The amount of solute per litre/cubic decimeter of solution. In symbols: $c = \frac{n}{V}$ Unit: mol·dm ⁻³
Empirical formula	The simplest positive integer ratio of atoms present in a compound.
Percentage yield	Yield is the amount of product obtained from a reaction. percentage yield = $\frac{\text{actual mass obtained}}{\text{calculated mass}} \times 100$
Percentage purity	percentage purity = $\frac{\text{mass of pure chemical}}{\text{total mass of sample}} \times 100$
Percentage composition	The percentage of each of the components in a substance. Percentage of component = $\frac{\text{mass contributed by component}}{\text{mass of all components}} \times 100$
Limiting reagents	The substance that is totally consumed when the chemical reaction is complete.

CHEMICAL CHANGE: ENERGY AND CHANGE

Heat of reaction (ΔH)	The energy absorbed or released in a chemical reaction.
Exothermic reactions	Reactions that release energy. ($\Delta H < 0$)
Endothermic reactions	Reactions that absorb energy. ($\Delta H > 0$)
Activation energy	The minimum energy needed for a reaction to take place.
Activated complex	The unstable transition state from reactants to products.

CHEMICAL CHANGE: RATE OF REACTION

Reaction rate	The change in concentration of reactants or products per unit time. Rate = $\frac{\Delta c}{\Delta t}$ Unit: mol·dm ⁻³ ·s ⁻¹
Collision theory	a model that explains reaction rate as the result of particles colliding with a certain minimum energy.
Catalyst	A substance that increases the rate of a chemical reaction without itself undergoing a permanent change. (A catalyst increases the rate of a reaction by providing an alternative path of lower activation energy. It therefore decreases the net/total activation energy.)
Factors that affect reaction rate	Nature of reacting substances, surface area, concentration (pressure for gases), temperature and the presence of a catalyst.


CHEMICAL CHANGE: CHEMICAL EQUILIBRIUM	
Open system	A system which continuously interacts with the environment – it exchanges matter and energy with its environment.
Closed system	A system that only exchanges energy with its surroundings, but it does not exchange matter with its surroundings.
Reversible reaction	A reaction is reversible when products can be converted back to reactants.
Chemical equilibrium	Dynamic equilibrium when the rate of the forward reaction equals the rate of the reverse reaction.
Factors that influence the equilibrium position	Pressure (gases only), concentration and temperature.
Le Chatelier's principle	When the equilibrium in a closed system is disturbed, the system will re-instate a new equilibrium by favouring the reaction that will oppose the disturbance.

CHEMICAL CHANGE: ACIDS AND BASES	
Acid-base indicator	A dye used to distinguish between acidic and basic solutions by means of the colour changes it undergoes in these solutions.
Amphiprotic substance/ampholyte	A substance that can act as either an acid or a base.
Arrhenius theory	An acid is a substance that produces hydrogen ions (H^+)/ hydronium ions (H_3O^+) when it dissolves in water. A base is a substance that produces hydroxide ions (OH^-) when it dissolves in water.
Auto-ionisation of water	A reaction in which water reacts with itself to form ions (hydronium ions and hydroxide ions).
Concentrated acids/bases	Contain a large amount (number of moles) of acid/base in proportion to the volume of water.
Conjugate acid-base pair	A pair of compounds or ions that differ by the presence of one H^+ ion. Example: CO_3^{2-} and HCO_3^- OR HCl and Cl^-
Conjugate acid and base	A conjugate acid has one H^+ ion more than its conjugate base. Example: HCO_3^- is the conjugate acid of base CO_3^{2-} . CO_3^{2-} is the conjugate base of acid HCO_3^- .
Dilute acids/bases	Contain a small amount (number of moles) of acid/base in proportion to the volume of water.
Diprotic acid	An acid that can donate two protons. Example: H_2SO_4
Dissociation	The process in which ionic compounds split into ions.
Endpoint	The point in a titration where the indicator changes colour.
Equivalence point	The point in a reaction where equivalent amounts of acid and base have reacted completely.
Hydrolysis	The reaction of a salt with water. OR The reaction of an ion with water to produce the conjugate acid and a hydroxide ion or the conjugate base and a hydronium ion.
Ionisation	The process in which ions are formed during a chemical reaction.
Ion product of water	The product of the ions formed during auto-ionisation of water i.e. $[H_3O^+][OH^-]$ at 25 °C.
Ionisation constant of water (K_w)	The equilibrium value of the ion product $[H_3O^+][OH^-]$ at 25 °C.
K_a value	Ionisation constant for an acid.
K_b value	Dissociation or ionisation constant for a base.

Lowry-Brønsted theory	An acid is a proton (H^+ ion) donor. A base is a proton (H^+ ion) acceptor.
Monoprotic acid	An acid that can donate one proton. Example: HCl
Neutralisation	The reaction of an acid with a base to form a salt (ionic compound) and water.
pH	The negative of the logarithm of the hydronium ion concentration in $mol \cdot dm^{-3}$. In symbols: $pH = -\log[H_3O^+]$ Unit: None
pH scale	A scale from 0 – 14 used as a measure of the acidity and basicity of solutions where $pH = 7$ is neutral, $pH > 7$ is basic and $pH < 7$ is acidic.
Salt	The ionic compound that is the product of a neutralisation reaction.
Standard solution	A solution of precisely known concentration.
Strong bases	Dissociate completely in water to form a high concentration of OH^- ions. Examples: sodium hydroxide ($NaOH$) and potassium hydroxide (KOH)
Strong acids	Ionise completely in water to form a high concentration of H_3O^+ ions. Examples: hydrochloric acid (HCl), sulphuric acid (H_2SO_4) and nitric acid (HNO_3)
Titration	The procedure for determining the amount of acid (or base) in a solution by determining the volume of base (or acid) of known concentration that will completely react with it.
Weak acids	Ionise incompletely in water to form a low concentration of H_3O^+ ions. Examples: ethanoic acid (CH_3COOH) and oxalic acid ($(COOH)_2$)
Weak bases	Dissociate/ionise incompletely in water to form a low concentration of OH^- ions. Examples: ammonia (NH_3), sodium hydrogen carbonate ($NaHCO_3$), sodium carbonate (Na_2CO_3), potassium carbonate (K_2CO_3), calcium carbonate ($CaCO_3$)


CHEMICAL CHANGE: ELECTROCHEMICAL REACTIONS

Galvanic cell	A cell in which chemical energy is converted into electrical energy. A galvanic (voltaic) cell has self-sustaining electrode reactions.
Electrolytic cell	A cell in which electrical energy is converted into chemical energy.
Redox reaction	A reaction in which an electron transfer takes place.
Oxidation	A loss of electrons./An increase in oxidation number.
Reduction	A gain of electrons./A decrease in oxidation number.
Oxidising agent	A substance that is reduced/gains electrons/whose oxidation number decreases.
Reducing agent	A substance that is oxidised/loses electrons/whose oxidation number increases.
Anode	The electrode where oxidation takes place.
Cathode	The electrode where reduction takes place.
Electrolyte	A solution that conducts electricity through the movement of ions.
Electrolysis	The chemical process in which electrical energy is converted to chemical energy OR the use of electrical energy to produce a chemical change.
Salt bridge	The connection between two half-cells needed to ensure electrical neutrality in the cell. OR A component used in a galvanic cell to complete the circuit.
Electrodes	An electrical conductor used in a galvanic cell to make contact with a non-metallic part of the circuit e.g. the electrolyte.

 <p>Cell notation</p>	<p>A short way to represent a galvanic cell.</p> <p>When writing cell notation, the following convention should be used:</p> <ul style="list-style-type: none"> ○ The $\text{H}_2 \text{H}^+$ half-cell is treated just like any other half-cell. ○ Cell terminals (electrodes) are written on the outside of the cell notation. ○ Active electrodes: reducing agent oxidised species oxidising agent reduced species ○ Inert electrodes (usually Pt or C): Pt reducing agent oxidised species oxidising agent reduced species Pt <p>Example: $\text{Pt} \text{Cl}^-(\text{aq}) \text{Cl}_2(\text{g}) \text{F}_2(\text{g}) \text{F}^-(\text{aq}) \text{Pt}$</p>
Overall cell reaction	The reaction obtained by combining two half-reactions.
Positive value of the standard emf	The reaction is spontaneous under standard conditions.
Standard conditions for a galvanic cell	Temperature: $25\text{ }^\circ\text{C} / 298\text{ K}$ Concentration: $1\text{ mol}\cdot\text{dm}^{-3}$ Pressure (gases only): $101,3\text{ kPa} / 1\text{ atmosphere}$
Standard hydrogen electrode	<p>The reference electrode used to compile the Table of Standard Reduction Potentials. The hydrogen half-cell was given a standard reduction potential of 0 V.</p> <p>Half-cell notation: $\text{Pt} \text{H}_2(\text{g}) \text{H}^+(\text{aq})$</p> <p>Half-reaction: $2\text{H}^+ + 2\text{e}^- \rightleftharpoons \text{H}_2$</p>
Electroplating	The covering of an object with a metal by making it the cathode in an electrolytic cell.
Bauxite	The ore from which aluminium is recovered.
Cryolite	An aluminium compound in which aluminium oxide is dissolved to reduce the cost of the extraction of aluminium. Cryolite has a lower melting point than aluminium oxide.

CHEMICAL SYSTEMS: FERTILISERS

Non-mineral nutrients in plants	<p>C, H and O</p> <p>Obtained from the atmosphere (CO_2) and rain (H_2O)</p>
Primary nutrients for plants	<p>Nitrogen (N), phosphorous (P) and potassium (K)</p> <p>Function:</p> <p>N: essential for plant growth, green leaves</p> <p>P: development of roots, stems and seeds</p> <p>K: resistance against diseases and production of flowers/fruits</p>
NPK fertilisers	Fertilisers containing the three primary nutrients for plants. They contain ammonium nitrate (NH_4NO_3), ammonium phosphate [$(\text{NH}_4)_3\text{PO}_4$] and potassium chloride (KCl).
N : P : K ratio	The percentage nitrogen, phosphorous and potassium in a fertiliser.
Fractional distillation of liquid air	Industrial preparation of nitrogen from air.
Steam reforming	<p>Preparation of hydrogen from earth gas (methane)</p> <p>Reaction: $\text{CH}_4 + \text{H}_2\text{O} \rightarrow 3\text{H}_2 + \text{CO}$</p>
Haber process	<p>Industrial preparation of ammonia</p> <p>Reaction: $\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$</p> <p>Iron catalyst</p>

<p>Ostwald process</p> 	<p>Industrial preparation of nitric acid</p> <p>Reactions: Catalytic oxidation of ammonia; catalyst: Pt $4\text{NH}_3 + 5\text{O}_2 \rightleftharpoons 4\text{NO} + 6\text{H}_2\text{O}$</p> <p>$2\text{NO} + \text{O}_2 \rightarrow 2\text{NO}_2$ $3\text{NO}_2 + \text{H}_2\text{O} \rightarrow 2\text{HNO}_3 + \text{NO}$ OR $4\text{NO}_2 + 2\text{H}_2\text{O} + \text{O}_2 \rightarrow 4\text{HNO}_3$</p>
<p>Contact process</p>	<p>Industrial preparation of sulphuric acid</p> <p>Reactions: $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$ $2\text{SO}_2 + \text{O}_2 \rightarrow 2\text{SO}_3$ Contact catalyst: vanadium pentoxide (V_2O_5) $\text{SO}_3 + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2\text{S}_2\text{O}_7$ ($\text{H}_2\text{S}_2\text{O}_7$: fuming sulphuric acid OR pyro sulphuric acid OR oleum) $\text{H}_2\text{S}_2\text{O}_7 + \text{H}_2\text{O} \rightarrow 2\text{H}_2\text{SO}_4$</p>
<p>Preparation of ammonium sulphate</p>	<p>ammonia + sulphuric acid \rightarrow ammonium sulphate $2\text{NH}_3 + \text{H}_2\text{SO}_4 \rightarrow (\text{NH}_4)_2\text{SO}_4$</p>
<p>Preparation of ammonium nitrate</p>	<p>ammonia + nitric acid \rightarrow ammonium nitrate $\text{NH}_3 + \text{HNO}_3 \rightarrow \text{NH}_4\text{NO}_3$</p>
<p>Eutrophication</p>	<p>The process by which an ecosystem, e.g. a river or dam, becomes enriched with inorganic plant nutrients, especially phosphorus and nitrogen, resulting in excessive plant growth. As plant growth becomes excessive, the amount of dead and decaying plant material increases rapidly.</p>



QUESTIONS

ORGANIC MOLECULES: NOMENCLATURE

QUESTION 1

Consider the organic compounds represented by the letters **A** to **F** in the table below.

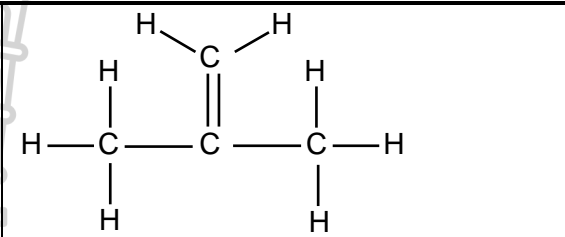
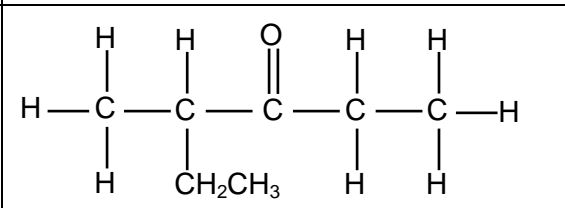
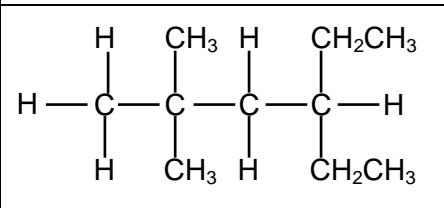
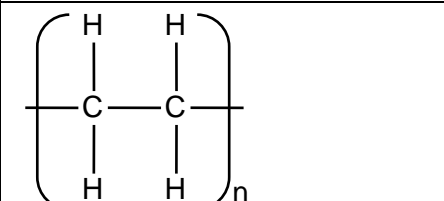
A	2,2,4-trimethylhexane	B	CH ₃ CH ₂ CH ₂ CH ₂ CHO
C		D	
E		F	Pentan-2-one

- 1.1 Write down the LETTER that represents the following:
- 1.1.1 An aldehyde (1)
 - 1.1.2 A condensation polymer (1)
 - 1.1.3 A compound which has a carbonyl group bonded to two carbon atoms as its functional group (1)
- 1.2 Write down the IUPAC name of:
- 1.2.1 Compound **C** (3)
 - 1.2.2 The monomer of compound **D** (1)
- 1.3 Write down the structural formula of:
- 1.3.1 Compound **A** (2)
 - 1.3.2 Compound **F** (2)
- 1.4 The table contains compounds which are functional isomers.
- 1.4.1 Define the term *functional isomer*. (2)
 - 1.4.2 Write down the LETTERS that represent two compounds that are functional isomers. (1)

[14]

QUESTION 2

The letters **A** to **F** in the table below represent six organic compounds.

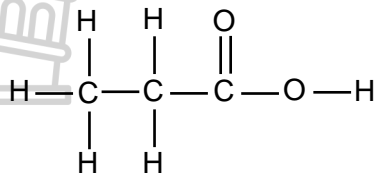
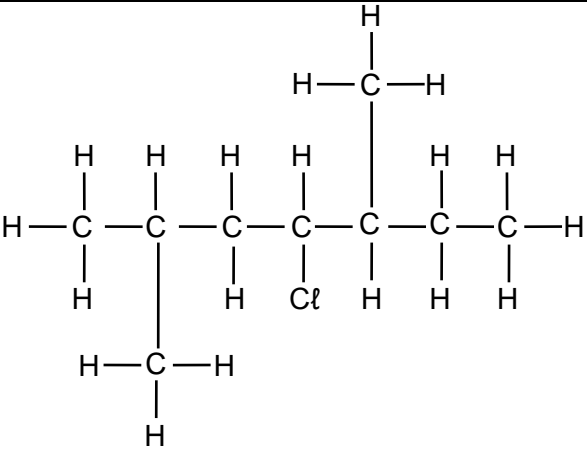
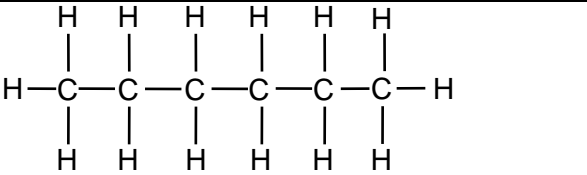
A		B	2-methylbutanoic acid
C		D	
E	But-2-ene	F	

- 2.1 Write down the:
- 2.1.1 NAME of the functional group of compound **B** (1)
 - 2.1.2 Homologous series to which compound **C** belongs (1)
 - 2.1.3 Type of polymerisation reaction that produces compound **F** (1)
- 2.2 Write down the IUPAC name of:
- 2.2.1 The monomer used to prepare compound **F** (1)
 - 2.2.2 Compound **C** (2)
 - 2.2.3 Compound **D** (2)
- 2.3 Write down the NAME or FORMULA of each product formed during the complete combustion of compound **D**. (2)
- 2.4 Write down the structural formula of:
- 2.4.1 Compound **B** (2)
 - 2.4.2 A CHAIN ISOMER of compound **A** (2)
- 2.5 A laboratory assistant uses bromine water to distinguish between compounds **D** and **E**. She adds bromine water to a sample of each in two different test tubes. She observes that the one compound decolourises the bromine water immediately, whilst the other one only reacts after placing the test tube in direct sunlight. Write down the:
- 2.5.1 Letter (**D** or **E**) of the compound that will immediately decolourise the bromine water (1)
 - 2.5.2 Name of the type of reaction that takes place in the test tube containing compound **D** (1)
 - 2.5.3 Structural formula of the organic product formed in the test tube containing compound **E** (2)

[18]

QUESTION 3

The letters **A** to **F** in the table below represent six organic compounds.

<p>A</p> 		<p>B</p> 	
<p>C</p>	<p>C_4H_8</p>	<p>D</p>	<p>$CH_3CH_2COCH_3$</p>
<p>E</p>	<p>$CH_3CH(CH_3)CH_2OH$</p>	<p>F</p> 	

Use the information in the table (where applicable) to answer the questions that follow.

3.1 Write down the LETTER that represents a compound that:
(A compound may be used more than once.)

- 3.1.1 Is a haloalkane (1)
- 3.1.2 Has a hydroxyl group as functional group (1)
- 3.1.3 Belongs to the same homologous series as ethanoic acid (1)

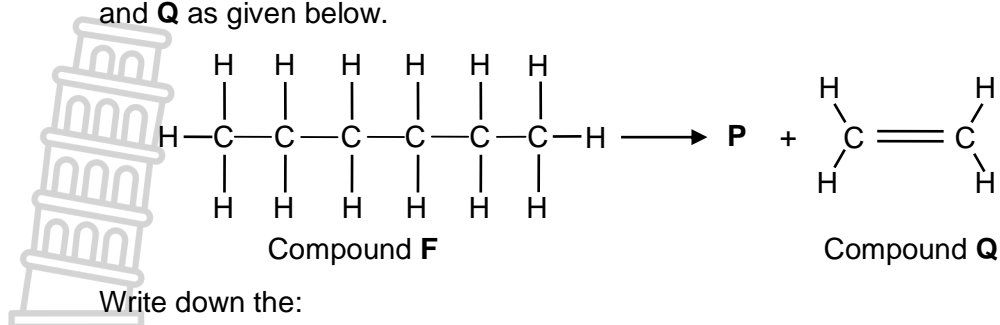
3.2 Write down the:

- 3.2.1 IUPAC name of compound **B** (3)
- 3.2.2 IUPAC name of compound **E** (2)
- 3.2.3 Structural formula of the functional group of compound **D** (1)

3.3 Compound **C** has CHAIN and POSITIONAL isomers.

- 3.3.1 Define the term *positional isomer*. (2)
- 3.3.2 Write down the IUPAC name of each of the TWO positional isomers of compound **C**. (4)
- 3.3.3 Write down the structural formula of a chain isomer of compound **C**. (2)

- 3.4 Compound **F** reacts at high pressure and high temperature to form compounds **P** and **Q** as given below.



Write down the:

- 3.4.1 Type of reaction that takes place (1)
- 3.4.2 IUPAC name of compound **Q** (1)
- 3.4.3 Molecular formula of compound **P** (1)

Compound **Q** is the monomer of a polymer used to make plastic bags.

- 3.4.4 Write down the NAME and CONDENSED FORMULA of this polymer. (3) **[23]**

QUESTION 4

The letters **A** to **D** in the table below represent four organic compounds.

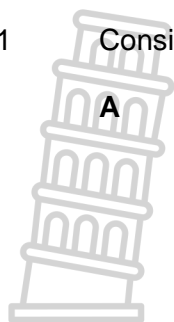
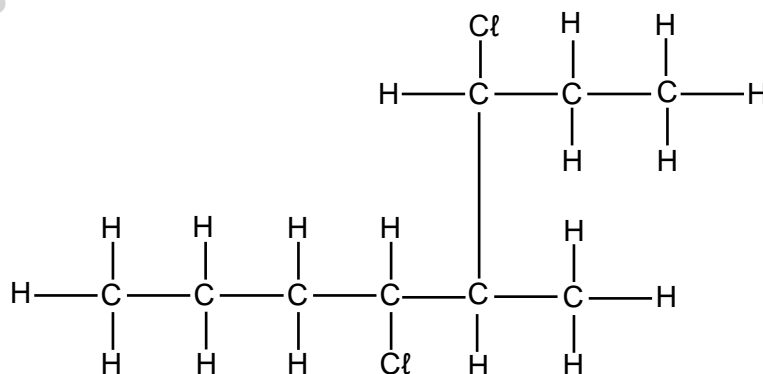
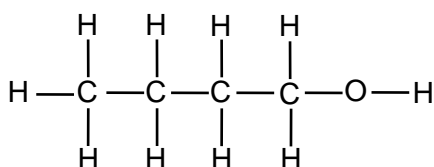
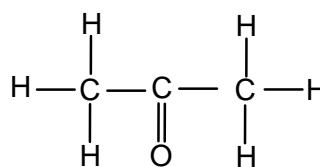
A	$ \begin{array}{ccccccc} & \text{H} & & \text{CH}_3 & & \text{H} & & \text{H} \\ & & & & & & & \\ \text{H} & -\text{C} & - & \text{C} & - & \text{C} & - & \text{C} \equiv \text{C} & - & \text{C} & - & \text{H} \\ & & & & & & & & & \\ & \text{CH}_3 & & \text{H} & & \text{CH}_2\text{CH}_3 & & & & \text{H} \end{array} $	B	$ \begin{array}{ccccccc} & \text{H} & & \text{H} & & \text{O} & & \text{H} \\ & & & & & & & \\ \text{H} & -\text{C} & - & \text{C} & - & \text{C} & - & \text{C} & - & \text{H} \\ & & & & & & & \\ & \text{H} & & \text{H} & & & & \text{H} \end{array} $
C	$\text{CH}_3\text{CH}_2\text{CHO}$	D	Butane

- 4.1 Write down the:
- 4.1.1 Letter that represents a ketone (1)
- 4.1.2 Structural formula of the functional group of compound **C** (1)
- 4.1.3 General formula of the homologous series to which compound **A** belongs (1)
- 4.1.4 IUPAC name of compound **A** (3)
- 4.1.5 IUPAC name of compound **B** (2)
- 4.2 Compound **D** is a gas used in cigarette lighters.
- 4.2.1 To which homologous series does compound **D** belong? (1)
- 4.2.2 Write down the STRUCTURAL FORMULA and IUPAC NAME of a structural isomer of compound **D**. (4)
- 4.2.3 Is the isomer in QUESTION 4.2.2 a CHAIN, POSITIONAL or FUNCTIONAL isomer? (1)
- 4.3 Compound **D** reacts with bromine (Br_2) to form 2-bromobutane. Write down the name of the:
- 4.3.1 Homologous series to which 2-bromobutane belongs (1)
- 4.3.2 Type of reaction that takes place (1)

[16]

QUESTION 5

5.1 Consider the organic compounds represented by the letters **A** to **C** below.

**A****B****C**

Write down the:

5.1.1 Name of the homologous series to which compound **C** belongs (1)

5.1.2 IUPAC name of compound **A** (3)

5.1.3 Structural formula of a tertiary alcohol that is a structural isomer of compound **B** (2)

5.2 An alcohol and methanoic acid are heated in the presence of concentrated sulphuric acid to form an ester.

5.2.1 What is the role of the concentrated sulphuric acid in this reaction? (1)

5.2.2 Write down the NAME or FORMULA of the inorganic product formed. (1)

The ester contains 6,67% hydrogen (H), 40% carbon (C) and 53,33% oxygen (O). The molar mass of the ester is $60 \text{ g}\cdot\text{mol}^{-1}$. Use a calculation to determine its:

5.2.3 Empirical formula (5)

5.2.4 Molecular formula (3)

Write down the:

5.2.5 Structural formula of methanoic acid (1)

5.2.6 IUPAC name of the ester (2)

[19]

QUESTION 6

Consider the organic compounds **A** to **F** below.

A	$\begin{array}{c} \text{CH}_3 \quad \quad \text{H} \\ \quad \quad \diagdown \quad \diagup \\ \quad \quad \text{C} = \text{C} \\ \quad \quad \diagup \quad \diagdown \\ \text{CH}_3 \quad \quad \text{CH}_3 \end{array}$	B	$\begin{array}{c} \text{H} \quad \quad \quad \text{H} \\ \quad \quad \quad \\ \text{H}-\text{C}-\text{C}\equiv\text{C}-\text{C}-\text{H} \\ \quad \quad \quad \\ \text{H} \quad \quad \quad \text{CH}_3 \end{array}$
C	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$	D	2,2-dimethylpropane
E	$\begin{array}{c} \text{H} \quad \quad \quad \text{O} \quad \quad \text{H} \quad \quad \text{H} \\ \quad \quad \quad \quad \quad \quad \quad \\ \text{H}-\text{C}-\text{O}-\text{C}-\text{C}-\text{C}-\text{H} \\ \quad \quad \quad \quad \quad \\ \text{H} \quad \quad \quad \text{H} \quad \quad \text{H} \end{array}$	F	$\text{CH}_3\text{CH}(\text{CH}_3)_2$

6.1 Write down the LETTER that represents a compound that:

6.1.1 Has a carbonyl group (1)

6.1.2 Is an alcohol (1)

6.1.3 Is a CHAIN ISOMER of $\text{CH}_3(\text{CH}_2)_3\text{CH}_3$ (1)

6.2 Write down the:

6.2.1 IUPAC name of compound **B** (2)

6.2.2 Structural formula of compound **F** (2)

6.2.3 IUPAC name of a POSITIONAL isomer of compound **A** (3)

6.3 Compound **E** is formed when a carboxylic acid reacts with another organic compound.

Write down the:

6.3.1 Homologous series to which compound **E** belongs (1)

6.3.2 NAME or FORMULA of the catalyst used for the preparation of compound **E** (1)

6.3.3 IUPAC name of compound **E** (2)

[14]

QUESTION 7

The letters **A** to **F** in the table below represent six organic compounds.

A	$ \begin{array}{c} \text{H} \qquad \qquad \text{H} \\ \qquad \qquad \\ \text{H}-\text{C}-\text{C}=\text{C}-\text{H} \\ \qquad \qquad \\ \text{H} \qquad \qquad \text{H}-\text{C}-\text{H} \\ \\ \text{H} \end{array} $	B	Ethyl ethanoate
C	2,3-dibromo-3-methylpentane	D	Polyethene
E	$ \begin{array}{c} \text{H} \qquad \text{O} \qquad \text{H} \\ \qquad \qquad \\ \text{H}-\text{C}-\text{C}-\text{O}-\text{C}-\text{H} \\ \qquad \qquad \\ \text{H} \qquad \qquad \text{H} \end{array} $	F	$ \begin{array}{c} \text{H} \qquad \text{H} \qquad \text{H} \qquad \text{O}-\text{H} \\ \qquad \qquad \qquad \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{C}=\text{O} \\ \qquad \qquad \\ \text{H} \qquad \text{H} \qquad \text{H} \end{array} $

7.1 Write down the LETTER that represents the following:

- 7.1.1 A hydrocarbon (1)
- 7.1.2 A functional isomer of compound **F** (1)
- 7.1.3 A compound which belongs to the same homologous series as compound **B** (1)
- 7.1.4 A plastic (1)

7.2 Write down the STRUCTURAL FORMULA of EACH of the following:

- 7.2.1 Compound **C** (3)
- 7.2.2 The acid used to prepare compound **B** (2)
- 7.2.3 The monomer used to make compound **D** (2)

7.3 Compound **A** reacts with an unknown reactant, **X**, to form 2-methylpropane.

Write down the:

- 7.3.1 NAME of reactant **X** (1)
- 7.3.2 Type of reaction that takes place (1)

[13]

ORGANIC MOLECULES: PHYSICAL PROPERTIES

QUESTION 1

- 1.1 Give a reason why alkanes are *saturated* hydrocarbons. (1)
- 1.2 Write down the structural formula of:
- 1.2.1 The functional group of alcohols (1)
- 1.2.2 A tertiary alcohol that is a structural isomer of butan-1-ol (2)
- 1.3 Learners investigate factors that influence the boiling points of alkanes and alcohols. In one of the investigations they determine the boiling points of the first three alkanes.
- 1.3.1 Write down an investigative question for this investigation. (2)
- 1.3.2 Fully explain why the boiling point increases from methane to propane. (3)
- 1.4 The learners find that the boiling point of propan-1-ol is higher than that of propane. Explain this observation by referring to the TYPE of INTERMOLECULAR FORCES present in each of these compounds. (3)
- [12]**

QUESTION 2

Learners use compounds **A** to **C**, shown in the table below, to investigate a factor which influences the boiling point of organic compounds.

A	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$
B	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$
C	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$

- 2.1 Which ONE of the compounds (**A**, **B** or **C**) has the highest boiling point? (1)
- 2.2 For this investigation, write down the:
- 2.2.1 Independent variable (1)
- 2.2.2 Dependent variable (1)
- 2.3 Write down the name of the type of Van der Waals force that occurs between the molecules of compound **B**. (1)
- 2.4 How will the vapour pressure of 2-methylpentane compare to that of compound **C**? Write down only HIGHER THAN, LOWER THAN or EQUAL TO. (1)

The learners now compare the boiling points of compounds **D** and **E**, shown in the table below.

D	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$
E	$\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$

- 2.5 How does the boiling point of compound **D** compare to that of compound **E**? Write down HIGHER THAN, LOWER THAN or EQUAL TO. Fully explain the answer. (4)
- [9]**

QUESTION 3

The table below shows five organic compounds represented by the letters **A** to **E**.

A	CH ₄
B	CH ₃ CH ₃
C	CH ₃ CH ₂ CH ₃
D	CH ₃ CH ₂ CH ₂ CH ₃
E	CH ₃ CH ₂ OH

3.1 Is compound **B** SATURATED or UNSATURATED? Give a reason for the answer. (2)

Consider the boiling points of compounds **A** to **E** given in random order below and use them, where applicable, to answer the questions that follow.

0 °C	- 162 °C	- 42 °C	- 89 °C	78 °C
------	----------	---------	---------	-------

3.2 Write down the boiling point of:

3.2.1 Compound **C** (1)

3.2.2 Compound **E** (1)

3.3 Explain the difference in boiling points of compounds **C** and **E** by referring to the TYPE of intermolecular forces present in EACH of these compounds. (3)

3.4 Does vapour pressure INCREASE or DECREASE from compounds **A** to **D**? Fully explain the answer. (4)

3.5 How will the vapour pressure of 2-methylpropane compare to the vapour pressure of compound **D**? Write down only HIGHER THAN, LOWER THAN or EQUAL TO. (1)
[12]

QUESTION 4

Four compounds of comparable molecular mass are used to investigate the effect of functional groups on vapour pressure. The results obtained are shown in the table below.

COMPOUND		VAPOUR PRESSURE (kPa at 20 °C)
A	Butane	204
B	Propan-2-one	24,6
C	Propan-1-ol	2
D	Ethanoic acid	1,6

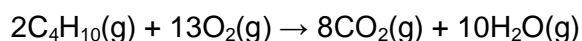
4.1 Define the term *functional group* of an organic compound. (2)

4.2 Which ONE of the compounds (**A**, **B**, **C** or **D**) in the table has the:

4.2.1 Highest boiling point
(Refer to the vapour pressures in the table to give a reason for the answer.) (2)

4.2.2 Weakest intermolecular forces (1)

- 4.3 Refer to the type of intermolecular forces to explain the difference between the vapour pressure of compound **A** and compound **B**. (3)
- 4.4 The vapour pressures of compounds **C** and **D** are much lower than those of compounds **A** and **B**. Name the type of intermolecular force in **A** and **B** that is responsible for this difference. (1)
- 4.5 Briefly explain the difference in vapour pressure between compound **C** and compound **D**. (2)
- 4.6 During a combustion reaction in a closed container of adjustable volume, 8 cm^3 of compound **A** (butane) reacts in excess oxygen according to the following balanced equation:



If the initial volume of the oxygen in the container was 60 cm^3 , calculate the TOTAL volume of the gases that are present in the container at the end of the reaction. All the gases in the container are at the same temperature and pressure. (5)

[16]

QUESTION 5

- 5.1 Define the term *boiling point*. (2)
- 5.2 What is the relationship between strength of intermolecular forces and boiling point? (1)

The relationship between strength of intermolecular forces and boiling point is investigated using four organic compounds from different homologous series. The compounds and their boiling points are given in the table below.

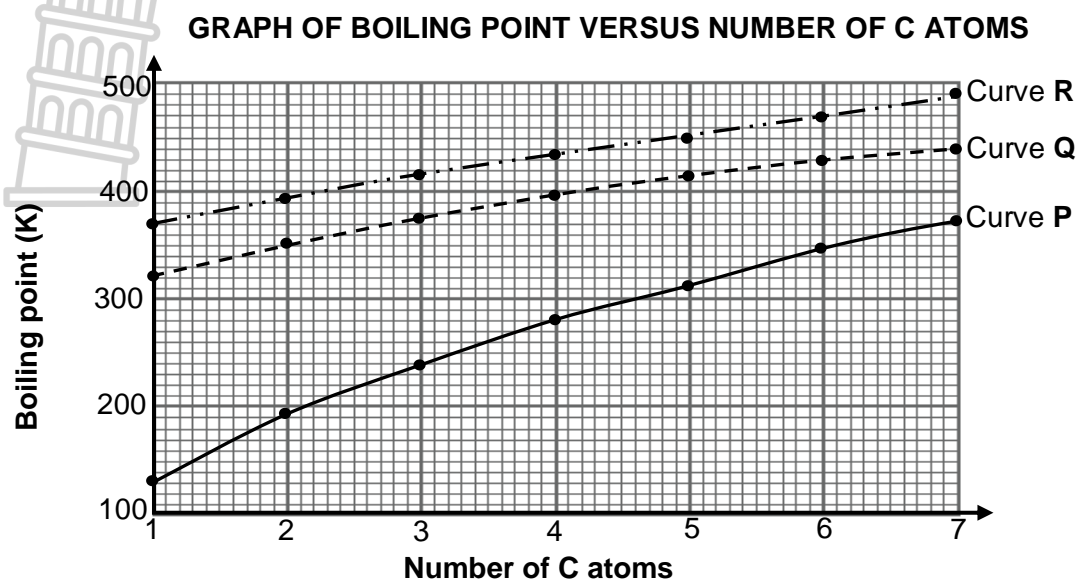
COMPOUND		BOILING POINT (°C)
A	Propane	-42
B	Propan-2-one	56
C	Propan-1-ol	97
D	Propanoic acid	141

- 5.3 Refer to the TYPE and the STRENGTH of intermolecular forces to explain the difference in boiling points between:
- 5.3.1 Compounds **A** and **B** (3)
- 5.3.2 Compounds **C** and **D** (3)
- 5.4 Is compound **B** a GAS or a LIQUID at room temperature? (1)

[10]

QUESTION 6

The relationship between boiling point and the number of carbon atoms in straight chain molecules of alkanes, carboxylic acids and alcohols is investigated. Curves **P**, **Q** and **R** are obtained.



- 6.1 Define the term *boiling point*. (2)
- 6.2 For curve **P**, write down a conclusion that can be drawn from the above results. (2)
- 6.3 Identify the curve (**P**, **Q** or **R**) that represents each of the following:
- 6.3.1 Alkanes (1)
- 6.3.2 Carboxylic acids (1)
- 6.4 Explain the answer to QUESTION 6.3.2 by referring to the:
- Types of intermolecular forces present in alkanes, carboxylic acids and alcohols
 - Relative strengths of these intermolecular forces
 - Energy needed
- (5)
[11]

QUESTION 7

The boiling points of three isomers are given in the table below.

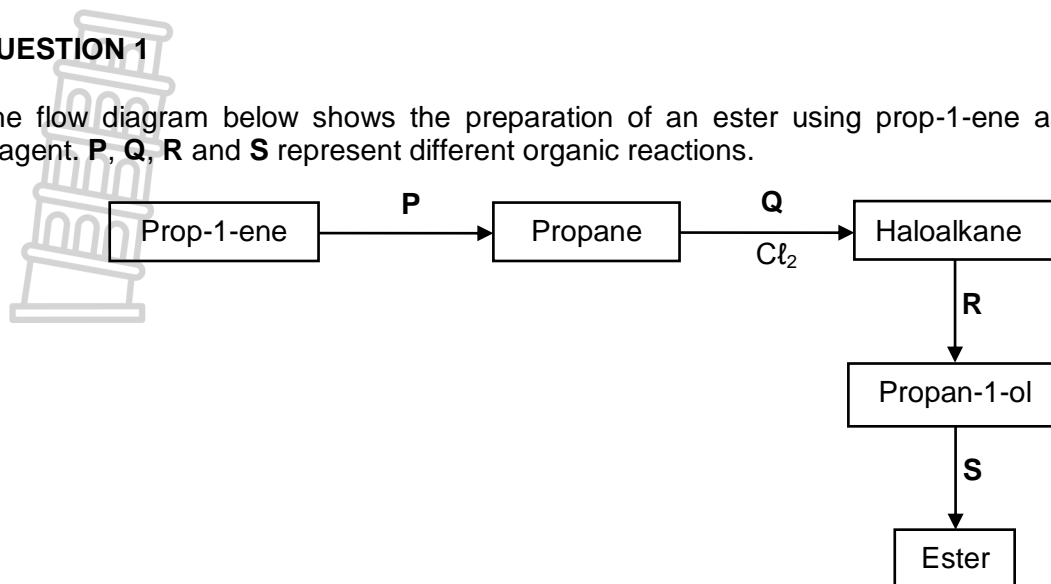
	ISOMERS	BOILING POINT (°C)
A	2,2-dimethylpropane	9
B	2-methylbutane	28
C	pentane	36

- 7.1 Define the term *structural isomer*. (2)
- 7.2 What type of isomers (POSITIONAL, CHAIN or FUNCTIONAL) are these three compounds? (1)
- 7.3 Explain the trend in the boiling points from compound **A** to compound **C**. (3)
- 7.4 Which ONE of the three compounds (**A**, **B** or **C**) has the highest vapour pressure? Refer to the data in the table to give a reason for the answer. (2)
- 7.5 Use MOLECULAR FORMULAE and write down a balanced equation for the complete combustion of compound **B**. (3)
- [11]**

ORGANIC MOLECULES: ORGANIC REACTIONS

QUESTION 1

The flow diagram below shows the preparation of an ester using prop-1-ene as a starting reagent. **P**, **Q**, **R** and **S** represent different organic reactions.

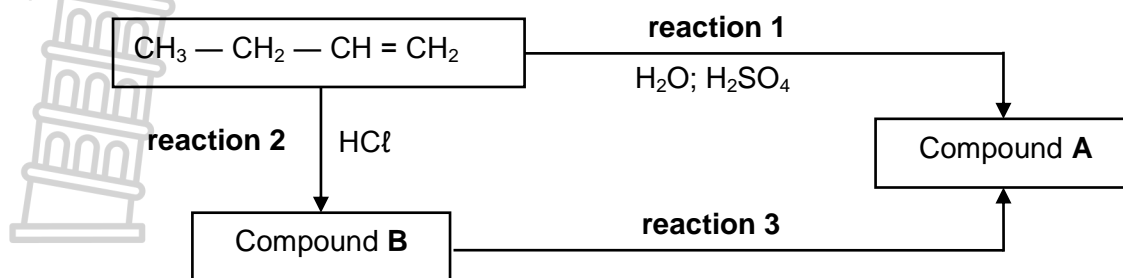


- 1.1 Write down the type of reaction represented by:
- 1.1.1 **Q** (1)
- 1.1.2 **R** (1)
- 1.2 For reaction **P** write down the:
- 1.2.1 Type of addition reaction (1)
- 1.2.2 Balanced equation using structural formulae (3)
- 1.3 Write down the structural formula of the haloalkane formed in reaction **Q**. (2)
- 1.4 In reaction **S** propan-1-ol reacts with ethanoic acid to form the ester.
- For this reaction write down the:
- 1.4.1 Name of the reaction that takes place (1)
- 1.4.2 FORMULA or NAME of the catalyst needed (1)
- 1.4.3 Structural formula of the ester formed (2)
- 1.4.4 IUPAC name of the ester formed (2)
- 1.5 The propan-1-ol formed in reaction **R** can be converted to prop-1-ene. Write down the FORMULA or NAME of the inorganic reagent needed. (1)

[15]

QUESTION 2

In the flow diagram below, but-1-ene is used as starting material in the preparation of compound **A**.

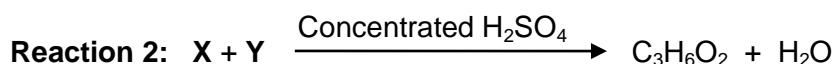
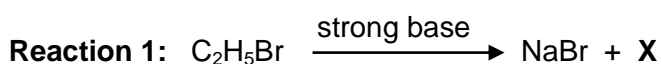


- 2.1 Is but-1-ene a SATURATED or UNSATURATED compound? Give a reason for the answer. (2)
- 2.2 Compound **A** is the major product formed in **reaction 1**.
Write down the:
- 2.2.1 Structural formula of compound **A** (2)
- 2.2.2 Type of reaction that takes place (1)
- 2.3 For compound **B**, write down the:
- 2.3.1 IUPAC name (2)
- 2.3.2 Structural formula of the positional isomer (2)
- 2.4 For **reaction 3**, write down:
- 2.4.1 TWO reaction conditions needed (2)
- 2.4.2 The type of reaction that occurs (1)
- 2.4.3 A balanced equation, using molecular formulae (3)
- [15]**

QUESTION 3

Consider the incomplete equations of two reactions below.

X represents the organic product formed in **reaction 1**, which is a SUBSTITUTION REACTION. In **reaction 2**, **X** reacts with reactant **Y** as shown.



- 3.1 Consider **reaction 1**. Write down the:
- 3.1.1 Type of substitution reaction that takes place (1)
- 3.1.2 TWO reaction conditions (2)
- 3.1.3 IUPAC name of compound **X** (1)

3.2 Consider **reaction 2**. Write down the:

3.2.1 Type of reaction that takes place (1)

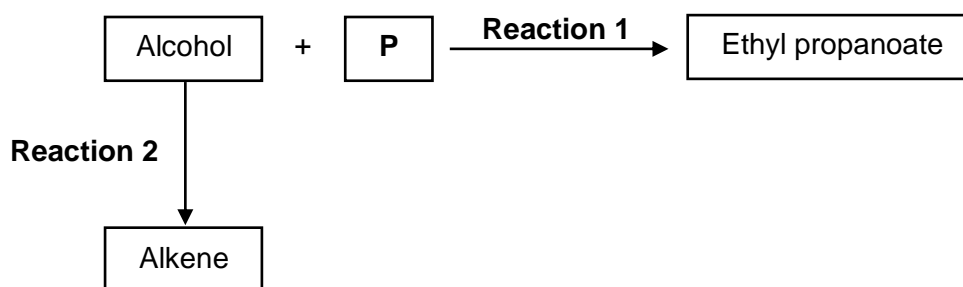
3.2.2 Structural formula of compound **Y** (2)

3.2.3 IUPAC name of the organic product (2)

[9]

QUESTION 4

4.1 The flow diagram below shows two organic reactions. The letter **P** represents an organic compound.



Use the information in the flow diagram to answer the questions that follow.

Write down the:

4.1.1 Type of reaction of which **Reaction 1** is an example (1)

4.1.2 STRUCTURAL FORMULA of the functional group of ethyl propanoate (1)

4.1.3 IUPAC name of compound **P** (1)

Reaction 2 takes place in the presence of an acid catalyst and heat.

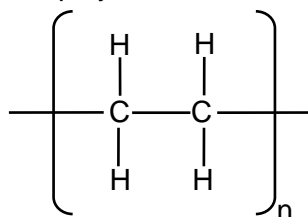
Write down the:

4.1.4 Type of reaction of which **Reaction 2** is an example (1)

4.1.5 NAME or FORMULA of the acid catalyst (1)

4.1.6 STRUCTURAL FORMULA of the alkene (2)

4.2 The condensed formula of a polymer is shown below.



Write down the:

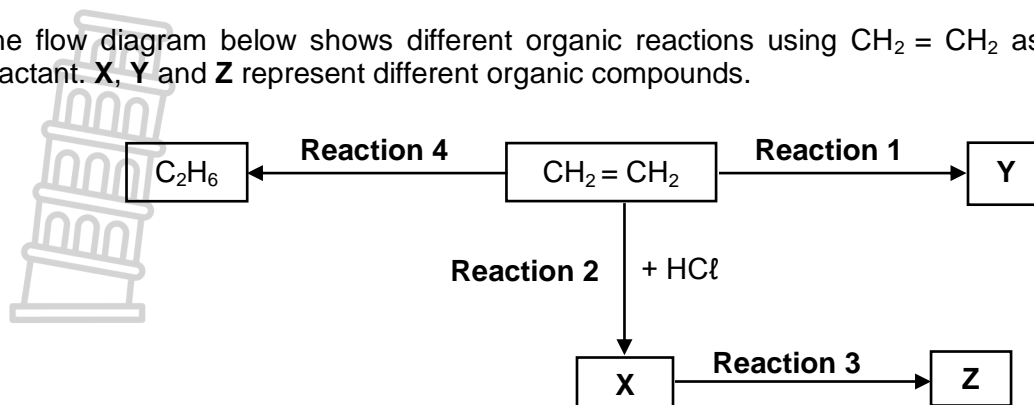
4.2.1 STRUCTURAL FORMULA of the monomer that is used to prepare the above polymer (2)

4.2.2 Type of polymerisation reaction (ADDITION or CONDENSATION) that is used to prepare this polymer (1)

[10]

QUESTION 5

The flow diagram below shows different organic reactions using $\text{CH}_2 = \text{CH}_2$ as the starting reactant. **X**, **Y** and **Z** represent different organic compounds.



5.1 During **Reaction 1**, $\text{CH}_2 = \text{CH}_2$ undergoes polymerisation to form compound **Y**.

For this reaction, write down the:

5.1.1 Type of polymerisation (1)

5.1.2 NAME of compound **Y** (1)

5.2 For **Reaction 2**, write down the:

5.2.1 IUPAC name of compound **X** (2)

5.2.2 Type of addition reaction of which this is an example (1)

5.3 During **Reaction 3**, compound **X** reacts with excess hot water.

Write down the:

5.3.1 STRUCTURAL FORMULA of compound **Z** (2)

5.3.2 NAME or FORMULA of the INORGANIC product (1)

5.4 **Reaction 4** is an addition reaction.

5.4.1 Is C_2H_6 a SATURATED or an UNSATURATED compound? Give a reason for the answer. (2)

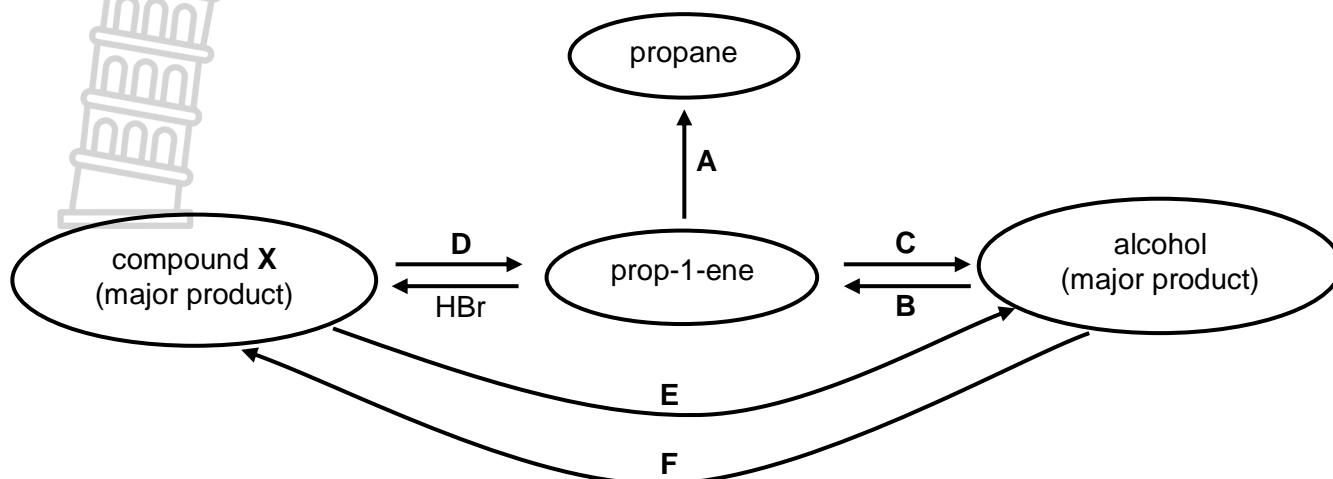
5.4.2 Write down the NAME or FORMULA of the INORGANIC reactant needed for this reaction. (1)

5.4.3 Using molecular formulae, write down a balanced equation for the complete combustion of C_2H_6 . (3)

[14]

QUESTION 6

The flow diagram below shows how prop-1-ene can be used to prepare other organic compounds.



6.1 Write down the type of reaction represented by:

6.1.1 **A** (1)

6.1.2 **D** (1)

6.1.3 **F** (1)

6.2 Write down the:

6.2.1 NAME or FORMULA of the catalyst needed for reaction **A** (1)

6.2.2 NAME or FORMULA of the inorganic reagent needed for reaction **B** (1)

6.2.3 Type of addition reaction represented by reaction **C** (1)

6.2.4 IUPAC name of compound **X** (2)

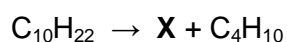
6.3 Use structural formulae to write down a balanced equation for reaction **B**. (5)

6.4 Both reactions **D** and **E** take place in the presence of a strong base. State TWO conditions that will favour reaction **D** over reaction **E**. (2)

[15]

QUESTION 7

Butane (C_4H_{10}) is produced in industry by the THERMAL cracking of long-chain hydrocarbon molecules, as shown in the equation below. **X** represents an organic compound that is produced.



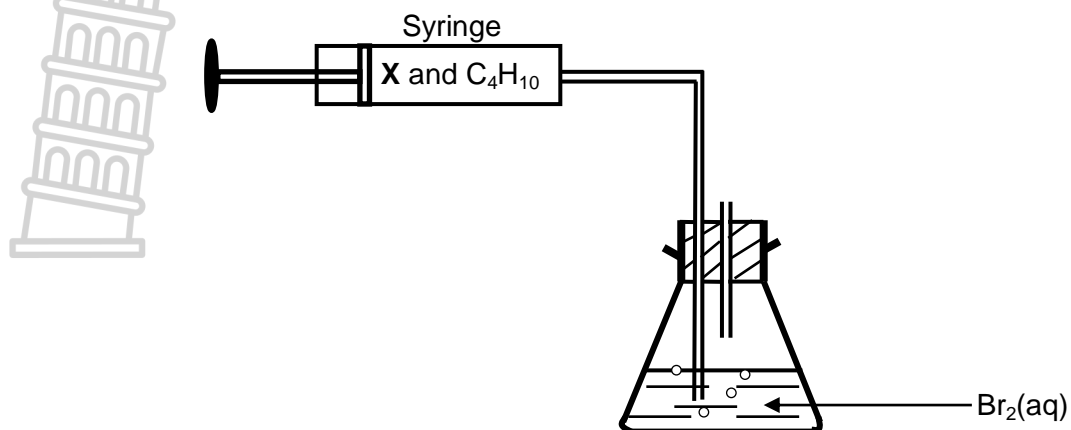
7.1 Write down:

7.1.1 ONE condition required for THERMAL cracking to take place (1)

7.1.2 The molecular formula of compound **X** (1)

7.1.3 The homologous series to which compound **X** belongs (1)

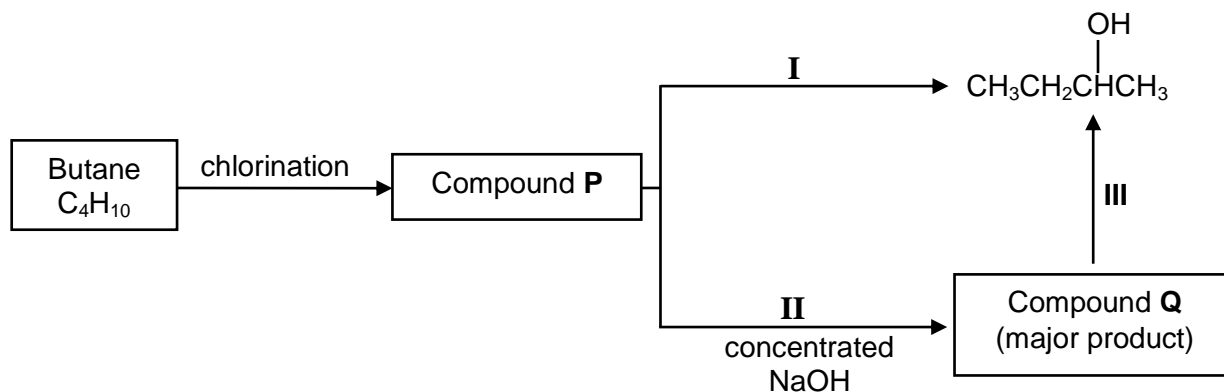
- 7.2 A mixture of the two gases, compound **X** and butane, is bubbled through bromine water, $\text{Br}_2(\text{aq})$, in a conical flask, as illustrated below. THE REACTION IS CARRIED OUT IN A DARKENED ROOM.



The colour of the bromine water changes from reddish brown to colourless when the mixture of the two gases is bubbled through it.

Which ONE of the gases (**X** or BUTANE) decolorises the bromine water? Explain the answer. (4)

- 7.3 Study the flow diagram below, which represents various organic reactions, and answer the questions that follow.



Write down the:

- 7.3.1 IUPAC name of compound **P** (2)
- 7.3.2 Type of reaction labelled **I** (1)
- 7.3.3 Structural formula of compound **Q** (2)
- 7.3.4 The type of addition reaction represented by reaction **III** (1)

[13]

REACTION RATE AND ENERGY IN CHEMICAL REACTIONS

QUESTION 1

- 1.1 Define the term *reaction rate* in words. (2)

Learners use the reaction between IMPURE POWDERED calcium carbonate and excess hydrochloric acid to investigate reaction rate. The balanced equation for the reaction is:



They perform four experiments under different conditions of concentration, mass and temperature as shown in the table below. They use identical apparatus in the four experiments and measure the volume of gas released in each experiment.

	EXPERIMENT			
	1	2	3	4
Concentration of acid ($\text{mol}\cdot\text{dm}^{-3}$)	1	0,5	1	1
Mass of impure calcium carbonate (g)	15	15	15	25
Initial temperature of acid ($^{\circ}\text{C}$)	30	30	40	40

- 1.2 The results of experiments **1** and **3** are compared in the investigation.

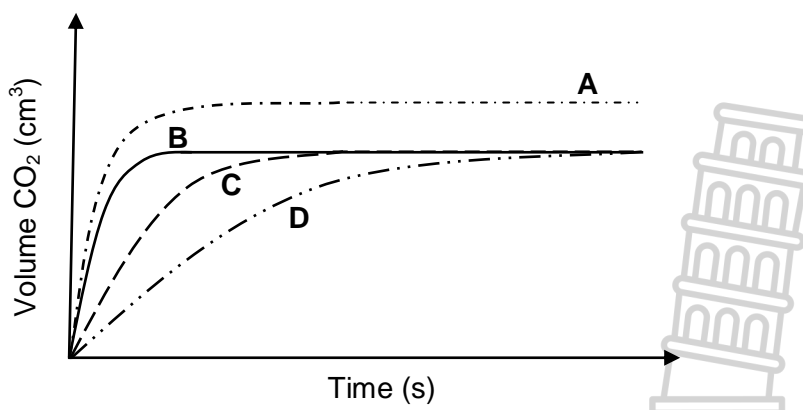
Write down the:

- 1.2.1 Independent variable (1)

- 1.2.2 Dependent variable (1)

- 1.3 Use the collision theory to explain why the reaction rate in experiment **4** will be higher than that in experiment **3**. (3)

The learners obtain graphs **A**, **B**, **C** and **D** below from their results.



- 1.4 Which ONE of the graphs (**A**, **B**, **C** or **D**) represents experiment **1**? Fully explain the answer by comparing experiment **1** with experiments **2**, **3** and **4**. (6)

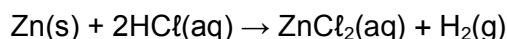
- 1.5 When the reaction in experiment **4** reaches completion, the volume of gas formed is $4,5 \text{ dm}^3$. Assume that the molar gas volume at 40°C is equal to $25,7 \text{ dm}^3$.

- Calculate the mass of the impurities present in the calcium carbonate. (5)

[18]

QUESTION 2

A group of learners uses the reaction of EXCESS hydrochloric acid (HCl) with zinc (Zn) to investigate factors which influence reaction rate. The balanced equation for the reaction is:



They use the same volume of hydrochloric acid and 1,2 g of zinc in each of five experiments. The reaction conditions and temperature readings before and after completion of the reaction in each experiment are summarised in the table below.

Experiment	REACTION CONDITIONS			Time (s)	
	Concentration of HCl (mol·dm ⁻³)	Temperature (°C)			
		Before	After		
1	0,5	20	34	granules	50
2	0,5	20	35	powder	10
3	0,8	20	36	powder	6
4	0,5	35	50	granules	8
5	0,5	20	34	granules	11

2.1 Is the reaction between hydrochloric acid and zinc EXOTHERMIC or ENDOTHERMIC? Give a reason for the answer by referring to the data in the table. (2)

2.2 Give a reason for the difference in reaction rate observed for **Experiments 1** and **2**. (1)

2.3 The learners compare the results of **Experiments 1** and **3** to draw a conclusion regarding the effect of concentration on reaction rate. Give a reason why this is not a fair comparison. (1)

2.4 How does the rate of the reaction in **Experiment 5** compare to that in **Experiment 1**? Write down FASTER THAN, SLOWER THAN or EQUAL TO.

Write down the factor responsible for the difference in the rate of reaction and fully explain, by referring to the collision theory, how this factor affects reaction rate. (5)

2.5 Calculate the rate at which the hydrochloric acid reacts in **Experiment 4** in mol·s⁻¹. (6)
[15]

QUESTION 3

A group of learners uses the reaction of clean magnesium ribbon with dilute hydrochloric acid to investigate factors that influence reaction rate. The balanced equation for the reaction is:



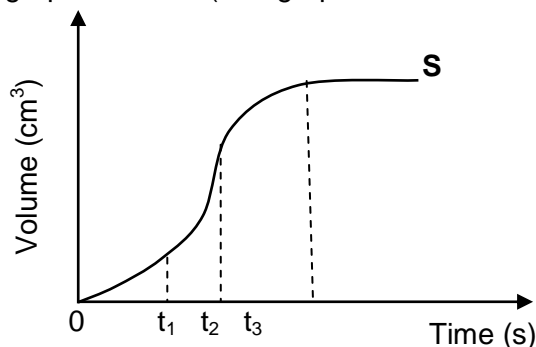
3.1 Is the above reaction EXOTHERMIC or ENDOTHERMIC? Give a reason for the answer. (2)

3.2 In one of the experiments 5 g magnesium ribbon was added to the hydrochloric acid solution.

3.2.1 If 30 cm³ dilute hydrochloric acid solution of concentration 1,5 mol·dm⁻³ is USED UP in 1 minute, calculate the average reaction rate in mol·s⁻¹. (5)



The volume of hydrogen gas produced as a function of time in this experiment is represented by graph **S** below. (The graph is NOT drawn to scale.)

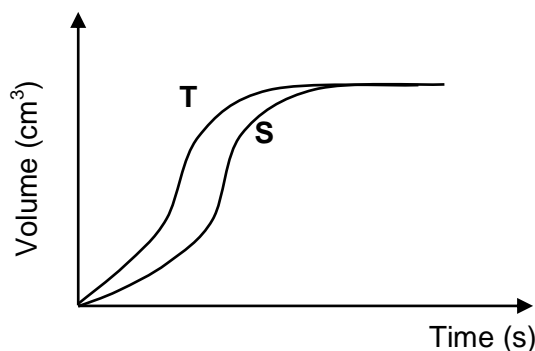


3.2.2 How does the rate of the reaction change between:
(Write down INCREASES, DECREASES or NO CHANGE.)

(a) t_1 and t_2
Use the collision theory to explain the answer. (4)

(b) t_2 and t_3
Give a reason for the answer without referring to the graph. (2)

3.3 In another experiment they add 5 g of magnesium to 30 cm³ of dilute hydrochloric acid of concentration 1,5 mol·dm⁻³. They obtained graph **T** below. (The graph is NOT drawn to scale.)



Give TWO possible reasons why graph **T** differs from graph **S**.

(2)
[15]

QUESTION 4

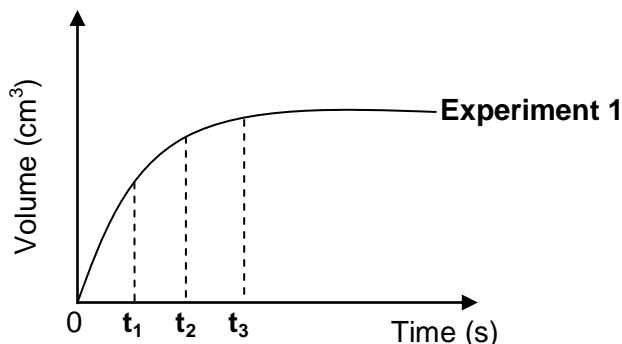
Dilute acids, indicated in the table below, react with EXCESS zinc in each of the three experiments to produce hydrogen gas. The zinc is completely covered with the acid in each experiment.

EXPERIMENT	DILUTE ACID
1	100 cm ³ of 0,1 mol·dm ⁻³ H ₂ SO ₄
2	50 cm ³ of 0,2 mol·dm ⁻³ H ₂ SO ₄
3	100 cm ³ of 0,1 mol·dm ⁻³ HCl

The volume of hydrogen gas produced is measured in each experiment.

4.1 Name TWO essential apparatuses needed to determine the rate of hydrogen production. (2)

The graph below was obtained for **Experiment 1**.



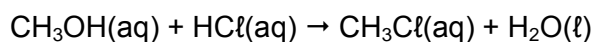
Use this graph and answer the questions that follow.

- 4.2 At which time (t_1 , t_2 or t_3) is the:
- 4.2.1 Reaction rate the highest (1)
- 4.2.2 Mass of zinc present in the flask the smallest (1)
- 4.3 In which time interval, **between t_1 and t_2** OR **between t_2 and t_3** , does the largest volume of hydrogen gas form per second? (1)
- 4.4 Redraw the graph for **Experiment 1** in the ANSWER BOOK. On the same set of axes, sketch the graphs that will be obtained for **Experiments 2** and **3**. Clearly label the three graphs as **EXPERIMENT 1**, **EXPERIMENT 2** and **EXPERIMENT 3**. (4)
- 4.5 The initial mass of zinc used in each experiment is 0,8 g. The balanced equation for the reaction in **Experiment 3** is:
- $$\text{Zn(s)} + 2\text{HCl(aq)} \rightarrow \text{ZnCl}_2\text{(aq)} + \text{H}_2\text{(g)}$$
- 4.5.1 Calculate the mass of zinc present in the flask after completion of the reaction in **Experiment 3**. (5)
- 4.5.2 How will the mass of zinc present in the flask after completion of the reaction in **Experiment 2** compare to the answer to QUESTION 5.5.1? Write down only LARGER THAN, SMALLER THAN or EQUAL TO. (1)

[15]

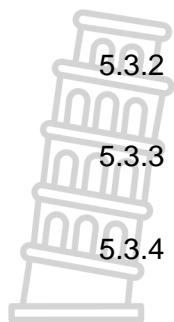
QUESTION 5

Methanol and hydrochloric acid react according to the following balanced equation:



- 5.1 State TWO factors that can INCREASE the rate of this reaction. (2)
- 5.2 Define the term *reaction rate*. (2)
- 5.3 The rate of the reaction between methanol and hydrochloric acid is investigated. The concentration of HCl(aq) was measured at different time intervals. The following results were obtained:

TIME (MINUTES)	HCl CONCENTRATION (mol·dm ⁻³)
0	1,90
15	1,45
55	1,10
100	0,85
215	0,60



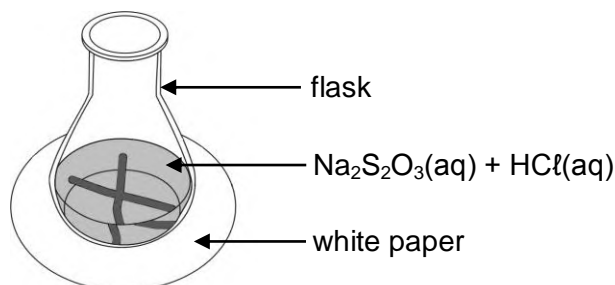
- 5.3.1 Calculate the average reaction rate, in $(\text{mol}\cdot\text{dm}^{-3})\cdot\text{min}^{-1}$ during the first 15 minutes. (3)
- 5.3.2 Use the data in the table to draw a graph of concentration versus time on a graph paper. (3)
- 5.3.3 From the graph, determine the concentration of $\text{HCl}(\text{aq})$ at the 40th minute. (1)
- 5.3.4 Use the collision theory to explain why the reaction rate decreases with time. Assume that the temperature remains constant. (3)
- 5.3.5 Calculate the mass of $\text{CH}_3\text{Cl}(\text{aq})$ in the flask at the 215th minute. The volume of the reagents remains 60 cm^3 during the reaction. (5)
- [19]**

QUESTION 6

The reaction between dilute hydrochloric acid and sodium thiosulphate ($\text{Na}_2\text{S}_2\text{O}_3$) is used to investigate one of the factors that influences reaction rate. The balanced equation for the reaction is:



The hydrochloric acid solution is added to the sodium thiosulphate solution in a flask. The flask is placed over a cross drawn on a sheet of white paper, as shown in the diagram below. The time that it takes for the cross to become invisible is measured to determine the reaction rate.



Four experiments, **A** to **D**, are conducted during this investigation. The volumes of reactants used in each of the four experiments and the times of the reactions are summarised in the table below.

Experiment	Volume of $\text{Na}_2\text{S}_2\text{O}_3(\text{aq})$ (cm^3)	Volume of $\text{H}_2\text{O}(\text{l})$ (cm^3)	Volume of $\text{HCl}(\text{aq})$ (cm^3)	Time (s)
A	25	0	5	50,0
B	20	5	5	62,5
C	15	10	5	83,3
D	10	15	5	125,0

- 6.1 State TWO factors that can influence the rate of the reaction above. (2)
- 6.2 Write down the NAME or FORMULA of the product that causes the cross to become invisible. (1)
- 6.3 Give a reason why water is added to the reaction mixture in experiments **B** to **D**. (1)
- 6.4 Write down an investigative question for this investigation. (2)
- 6.5 In which experiment (**A**, **B**, **C** or **D**) is the reaction rate the highest? (1)

6.6 Use the collision theory to explain the difference in reaction rate between experiments **B** and **D**. (3)

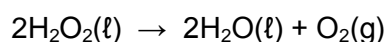
6.7 The original $\text{Na}_2\text{S}_2\text{O}_3$ solution was prepared by dissolving 62,50 g $\text{Na}_2\text{S}_2\text{O}_3$ crystals in distilled water in a 250 cm^3 volumetric flask.

Calculate the mass of sulphur, S, that will form in experiment **D** if $\text{Na}_2\text{S}_2\text{O}_3$ is the limiting reactant. (7)

[17]

QUESTION 7

Hydrogen peroxide, H_2O_2 , decomposes to produce water and oxygen according to the following balanced equation:



7.1 The activation energy (E_A) for this reaction is 75 kJ and the heat of reaction (ΔH) is –196 kJ.

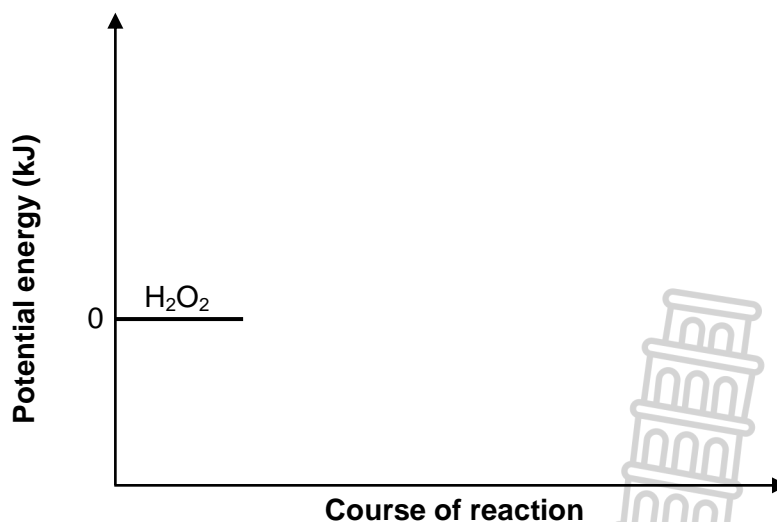
7.1.1 Define the term *activation energy*. (2)

7.1.2 Redraw the set of axes below in your ANSWER BOOK and then complete the potential energy diagram for this reaction.

Indicate the value of the potential energy of the following on the y-axis:

- Activated complex
- Products

(The graph does NOT have to be drawn to scale.)



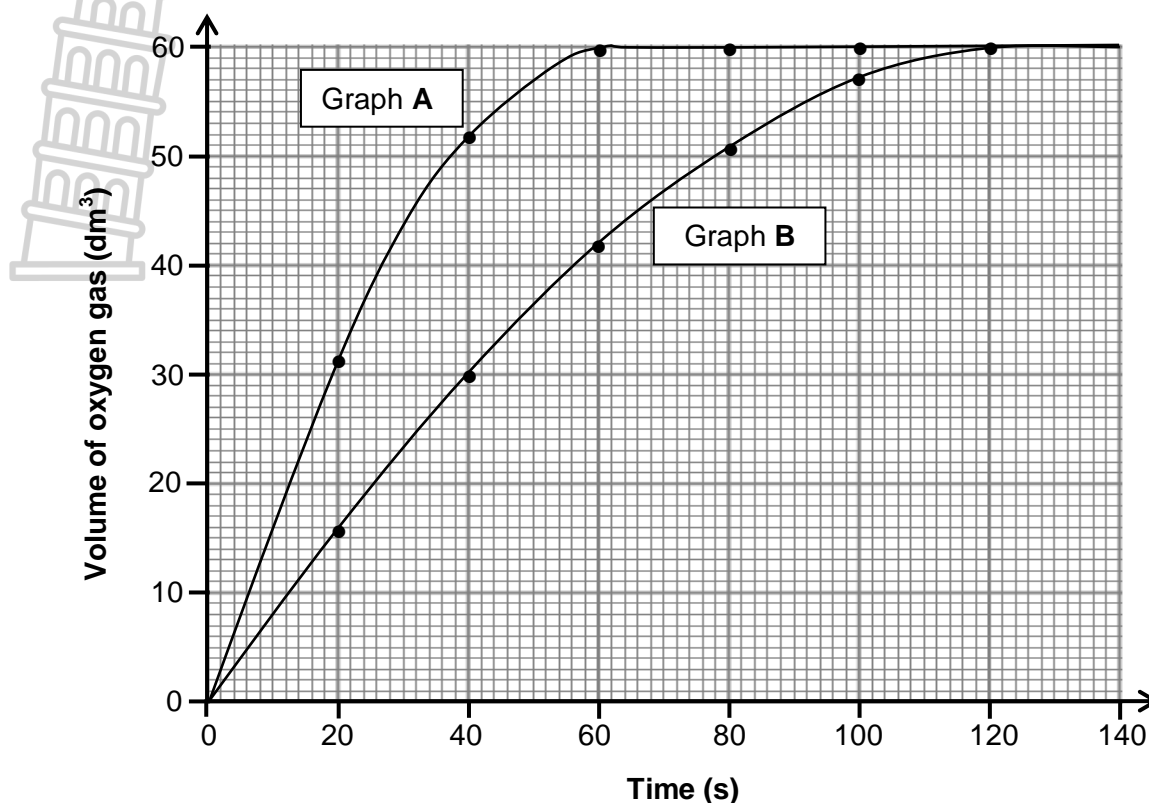
(3)

When powdered manganese dioxide is added to the reaction mixture, the rate of the reaction increases.

7.1.3 On the graph drawn for QUESTION 7.1.2, use broken lines to show the path of the reaction when the manganese dioxide is added. (2)

7.1.4 Use the collision theory to explain how manganese dioxide influences the rate of decomposition of hydrogen peroxide. (3)

7.2 Graphs **A** and **B** below were obtained for the volume of oxygen produced over time under different conditions.

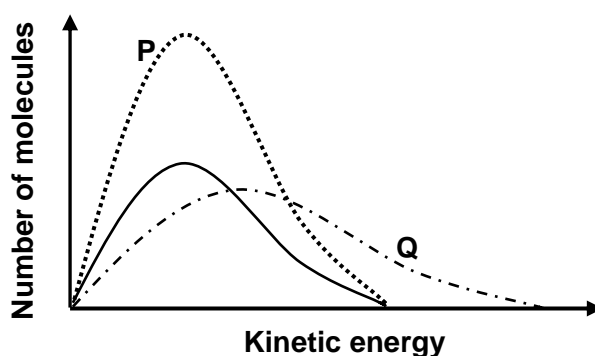


7.2.1 Calculate the average rate of the reaction (in $\text{dm}^3 \cdot \text{s}^{-1}$) between $t = 10 \text{ s}$ and $t = 40 \text{ s}$ for graph **A**. (3)

7.2.2 Use the information in graph **A** to calculate the mass of hydrogen peroxide used in the reaction. Assume that all the hydrogen peroxide decomposed. Use $24 \text{ dm}^3 \cdot \text{mol}^{-1}$ as the molar volume of oxygen. (4)

7.2.3 How does the mass of hydrogen peroxide used to obtain graph **B** compare to that used to obtain graph **A**? Choose from GREATER THAN, SMALLER THAN or EQUAL TO. (1)

7.3 Three energy distribution curves for the oxygen gas produced under different conditions are shown in the graph below. The curve with the solid line represents 1 mol of oxygen gas at 90°C .



Choose the curve (**P** or **Q**) that best represents EACH of the following situations:

7.3.1 1 mol of oxygen gas produced at 120°C (1)

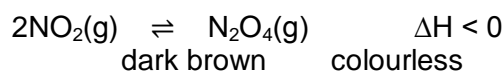
7.3.2 2 moles of oxygen gas produced at 90°C (1)

[20]

CHEMICAL EQUILIBRIUM

QUESTION 1

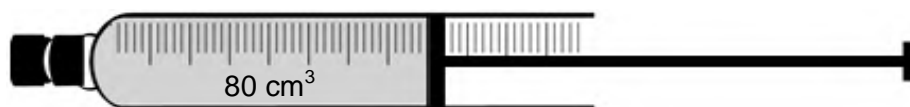
A certain amount of nitrogen dioxide gas (NO_2) is sealed in a gas syringe at 25°C . When equilibrium is reached, the volume occupied by the reaction mixture in the gas syringe is 80 cm^3 . The balanced chemical equation for the reaction taking place is:



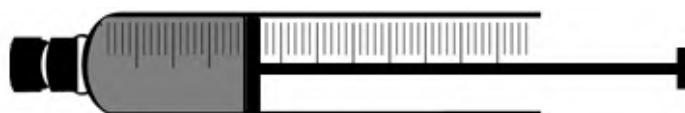
- 1.1 Define the term *chemical equilibrium*. (2)
- 1.2 At equilibrium the concentration of the $\text{NO}_2(\text{g})$ is $0,2\text{ mol}\cdot\text{dm}^{-3}$. The equilibrium constant for the reaction is 171 at 25°C .

Calculate the initial number of moles of $\text{NO}_2(\text{g})$ placed in the gas syringe. (8)

- 1.3 The diagram below shows the reaction mixture in the gas syringe after equilibrium is established.



The pressure is now increased by decreasing the volume of the gas syringe at constant temperature as illustrated in the diagram below.



- 1.3.1 IMMEDIATELY after increasing the pressure, the colour of the reaction mixture in the gas syringe appears darker than before. Give a reason for this observation. (1)

After a while a new equilibrium is established as illustrated below. The colour of the reaction mixture in the gas syringe now appears lighter than the initial colour.

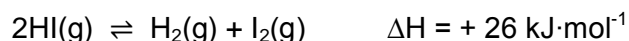


- 1.3.2 Use Le Chatelier's principle to explain the colour change observed in the gas syringe. (3)
- 1.4 The temperature of the reaction mixture in the gas syringe is now increased and a new equilibrium is established. How will each of the following be affected?
- 1.4.1 Colour of the reaction mixture
Write down only DARKER, LIGHTER or REMAINS THE SAME. (1)
- 1.4.2 Value of the equilibrium constant (K_c)
Write down only INCREASES, DECREASES or REMAINS THE SAME. (1)

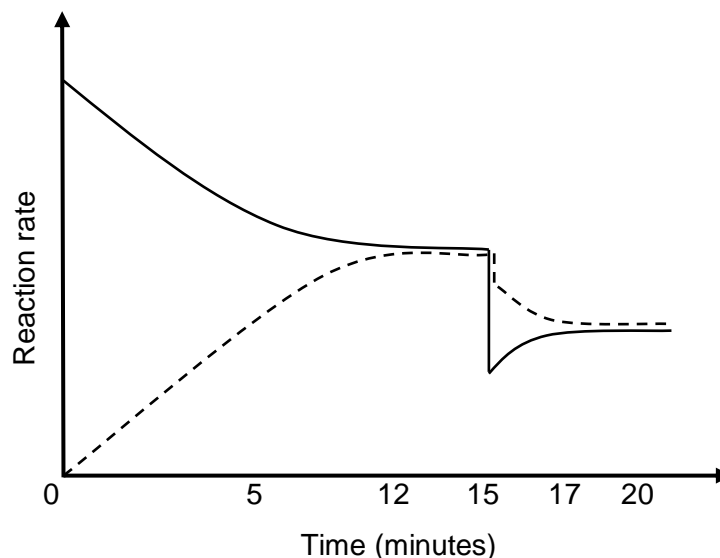
[16]

QUESTION 2

Pure hydrogen iodide, sealed in a 2 dm³ container at 721 K, decomposes according to the following balanced equation:



The graph below shows how reaction rate changes with time for this reversible reaction.



- 2.1 Write down the meaning of the term *reversible reaction*. (1)
- 2.2 How does the concentration of the reactant change between the 12th and the 15th minute? Write down only INCREASES, DECREASES or NO CHANGE. (1)
- 2.3 The rates of both the forward and the reverse reactions suddenly change at t = 15 minutes.
- 2.3.1 Give a reason for the sudden change in reaction rate. (1)
- 2.3.2 Fully explain how you arrived at the answer to QUESTION 2.3.1. (3)

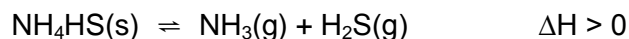
The equilibrium constant (K_c) for the forward reaction is 0,02 at 721 K.

- 2.4 At equilibrium it is found that 0,04 mol HI(g) is present in the container. Calculate the concentration of H₂(g) at equilibrium. (6)
- 2.5 Calculate the equilibrium constant for the reverse reaction. (1)
- 2.6 The temperature is now increased to 800 K. How will the value of the equilibrium constant (K_c) for the forward reaction change? Write down only INCREASES, DECREASES or REMAINS THE SAME. (1)

[14]

QUESTION 3

Initially excess $\text{NH}_4\text{HS}(\text{s})$ is placed in a 5 dm^3 container at $218 \text{ }^\circ\text{C}$. The container is sealed and the reaction is allowed to reach equilibrium according to the following balanced equation:



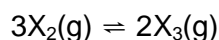
- 3.1 State Le Chatelier's principle. (2)
- 3.2 What effect will each of the following changes have on the amount of $\text{NH}_3(\text{g})$ at equilibrium? Write down only INCREASES, DECREASES or REMAINS THE SAME.
- 3.2.1 More $\text{NH}_4\text{HS}(\text{s})$ is added (1)
- 3.2.2 The temperature is increased (1)
- 3.3 The equilibrium constant for this reaction at $218 \text{ }^\circ\text{C}$ is $1,2 \times 10^{-4}$. Calculate the minimum mass of $\text{NH}_4\text{HS}(\text{s})$ that must be sealed in the container to obtain equilibrium. (6)

The pressure in the container is now increased by decreasing the volume of the container at constant temperature.

- 3.4 How will this change affect the number of moles of $\text{H}_2\text{S}(\text{g})$ produced? Fully explain the answer. (3)
- [13]**

QUESTION 4

An unknown gas, $\text{X}_2(\text{g})$, is sealed in a container and allowed to form $\text{X}_3(\text{g})$ at $300 \text{ }^\circ\text{C}$. The reaction reaches equilibrium according to the following balanced equation:



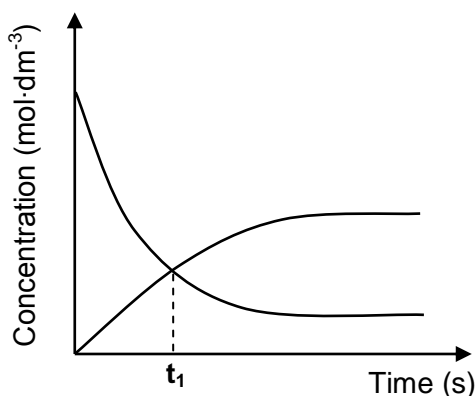
- 4.1 How will the rate of formation of $\text{X}_3(\text{g})$ compare to the rate of formation of $\text{X}_2(\text{g})$ at equilibrium? Write down only HIGHER THAN, LOWER THAN or EQUAL TO. (1)

The reaction mixture is analysed at regular time intervals. The results obtained are shown in the table below.

TIME (s)	$[\text{X}_2]$ ($\text{mol}\cdot\text{dm}^{-3}$)	$[\text{X}_3]$ ($\text{mol}\cdot\text{dm}^{-3}$)
0	0,4	0
2	0,22	0,120
4	0,08	0,213
6	0,06	0,226
8	0,06	0,226
10	0,06	0,226

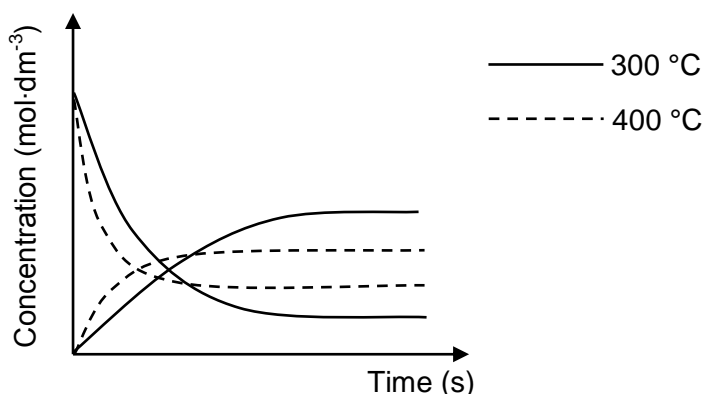
- 4.2 Calculate the equilibrium constant, K_c , for this reaction at $300 \text{ }^\circ\text{C}$. (4)
- 4.3 More $\text{X}_3(\text{g})$ is now added to the container.
- 4.3.1 How will this change affect the amount of $\text{X}_2(\text{g})$? Write down INCREASES, DECREASES or REMAINS THE SAME. (1)
- 4.3.2 Use Le Chatelier's principle to explain the answer to QUESTION 4.3.1. (2)

The curves on the set of axes below (not drawn to scale) was obtained from the results in the table on page 10.



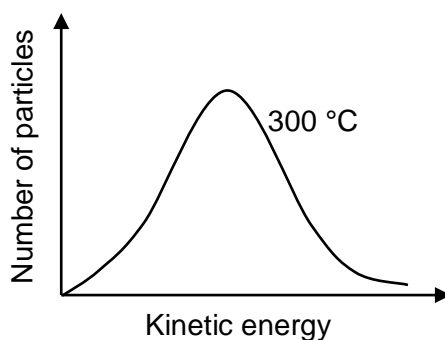
4.4 How does the rate of the forward reaction compare to that of the reverse reaction at t_1 ? Write down only HIGHER THAN, LOWER THAN or EQUAL TO. (1)

The reaction is now repeated at a temperature of 400 °C. The curves indicated by the dotted lines below were obtained at this temperature.



4.5 Is the forward reaction EXOTHERMIC or ENDOTHERMIC? Fully explain how you arrived at the answer. (4)

The Maxwell-Boltzmann distribution curve below represents the number of particles against kinetic energy at 300 °C.



4.6 Redraw this curve in the ANSWER BOOK. On the same set of axes, sketch the curve that will be obtained at 400 °C. Clearly label the curves as 300 °C and 400 °C respectively. (2)

[15]

QUESTION 5

Initially, 2,2 g of pure $\text{CO}_2(\text{g})$ is sealed in an empty 5 dm^3 container at $900 \text{ }^\circ\text{C}$.

- 5.1 Calculate the initial concentration of $\text{CO}_2(\text{g})$. (4)
- 5.2 Give a reason why equilibrium will not be established. (1)

$\text{CaCO}_3(\text{s})$ is now added to the 2,2 g $\text{CO}_2(\text{g})$ in the container and after a while equilibrium is established at $900 \text{ }^\circ\text{C}$ according to the following balanced equation:



The equilibrium constant for this reaction at $900 \text{ }^\circ\text{C}$ is 0,0108.

- 5.3 Give a reason why this reaction will only reach equilibrium in a SEALED container. (1)
- 5.4 Calculate the minimum mass of $\text{CaCO}_3(\text{s})$ that must be added to the container to achieve equilibrium. (7)
- 5.5 How will EACH of the following changes affect the amount of $\text{CO}_2(\text{g})$? Write down only INCREASES, DECREASES or REMAINS THE SAME. (1)
- 5.5.1 More $\text{CaCO}_3(\text{s})$ is added at $900 \text{ }^\circ\text{C}$ (1)
- 5.5.2 The pressure is increased (1)
- 5.6 It is found that the equilibrium constant (K_c) for this reaction is $2,6 \times 10^{-6}$ at $727 \text{ }^\circ\text{C}$. Is the reaction EXOTHERMIC or ENDOTHERMIC? Fully explain how you arrived at the answer. (4)
- [19]**

QUESTION 6

Carbon dioxide reacts with carbon in a closed system to produce carbon monoxide, $\text{CO}(\text{g})$, according to the following balanced equation:



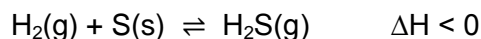
- 6.1 What does the double arrow indicate in the equation above? (1)
- 6.2 Is the above reaction an EXOTHERMIC reaction or an ENDOTHERMIC reaction? Give a reason for the answer. (2)

Initially an unknown amount of carbon dioxide is exposed to hot carbon at $800 \text{ }^\circ\text{C}$ in a sealed 2 dm^3 container. The equilibrium constant, K_c , for the reaction at this temperature is 14. At equilibrium it is found that 168,00 g carbon monoxide is present.

- 6.3 How will the equilibrium concentration of the product compare to that of the reactants? Choose from LARGER THAN, SMALLER THAN or EQUAL TO. (2)
- Give a reason for the answer. (No calculation is required.) (2)
- 6.4 Calculate the initial amount (in moles) of $\text{CO}_2(\text{g})$ present. (9)
- 6.5 State how EACH of the following will affect the yield of $\text{CO}(\text{g})$ at equilibrium. Choose from INCREASES, DECREASES or REMAINS THE SAME. (1)
- 6.5.1 More carbon is added at constant temperature. (1)
- 6.5.2 The pressure is increased. (1)
- 6.5.3 The temperature is increased. (1)
- [17]**

QUESTION 7

Hydrogen gas, $\text{H}_2(\text{g})$, reacts with sulphur powder, $\text{S}(\text{s})$, according to the following balanced equation:



The system reaches equilibrium at 90°C .

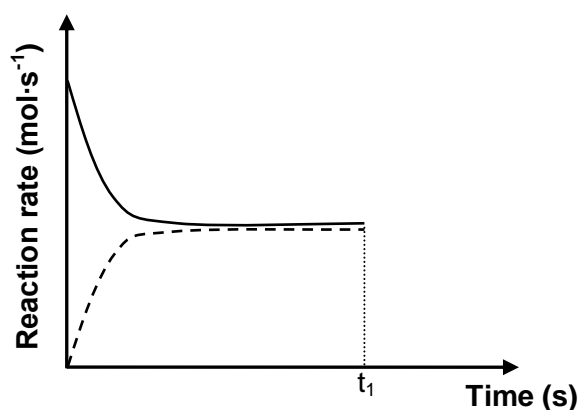
7.1 Define the term *chemical equilibrium*. (2)

7.2 How will EACH of the following changes affect the number of moles of $\text{H}_2\text{S}(\text{g})$ at equilibrium? Choose from INCREASES, DECREASES or REMAINS THE SAME.

7.2.1 The addition of more sulphur (1)

7.2.2 An increase in temperature
Use Le Chatelier's principle to explain the answer. (4)

7.3 The sketch graph below was obtained for the equilibrium mixture.



A catalyst is added to the equilibrium mixture at time t_1 .

Redraw the graph above in your ANSWER BOOK. On the same set of axes, complete the graph showing the effect of the catalyst on the reaction rates. (2)

Initially $0,16 \text{ mol H}_2(\text{g})$ and excess $\text{S}(\text{s})$ are sealed in a 2 dm^3 container and the system is allowed to reach equilibrium at 90°C .

An exact amount of $\text{Pb}(\text{NO}_3)_2$ solution is now added to the container so that ALL the $\text{H}_2\text{S}(\text{g})$ present in the container at EQUILIBRIUM is converted to $\text{PbS}(\text{s})$ according to the following balanced equation: $\text{Pb}(\text{NO}_3)_2(\text{aq}) + \text{H}_2\text{S}(\text{g}) \rightarrow \text{PbS}(\text{s}) + 2\text{HNO}_3(\text{aq})$

The mass of the PbS precipitate is $2,39 \text{ g}$.

7.4 Calculate the equilibrium constant K_c for the reaction $\text{H}_2(\text{g}) + \text{S}(\text{s}) \rightleftharpoons \text{H}_2\text{S}(\text{g})$ at 90°C . (9)

[18]

ACIDS AND BASES

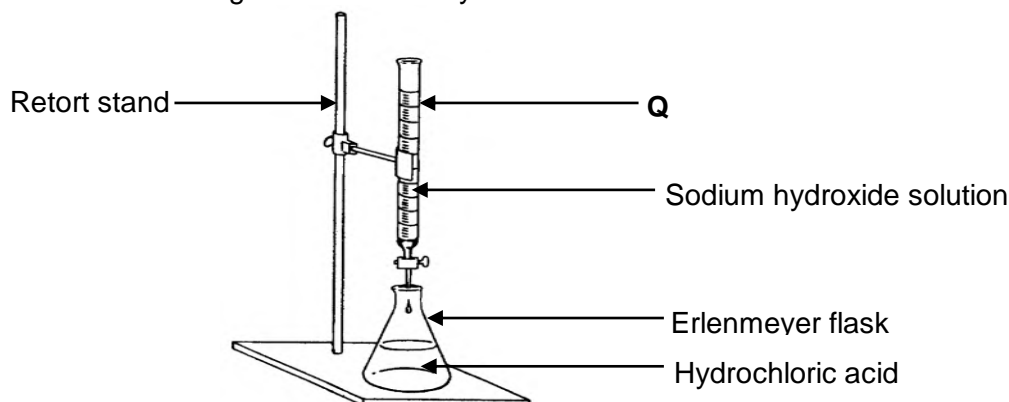
QUESTION 1

- 1.1 Nitric acid (HNO_3), an important acid used in industry, is a strong acid.
- 1.1.1 Give a reason why nitric acid is classified as a strong acid. (1)
- 1.1.2 Write down the NAME or FORMULA of the conjugate base of nitric acid. (1)
- 1.1.3 Calculate the pH of a $0,3 \text{ mol}\cdot\text{dm}^{-3}$ nitric acid solution. (3)

- 1.2 A laboratory technician wants to determine the percentage purity of magnesium oxide. He dissolves a $4,5 \text{ g}$ sample of the magnesium oxide in 100 cm^3 hydrochloric acid of concentration $2 \text{ mol}\cdot\text{dm}^{-3}$.

- 1.2.1 Calculate the number of moles of hydrochloric acid added to the magnesium oxide. (3)

He then uses the apparatus below to titrate the EXCESS hydrochloric acid in the above solution against a sodium hydroxide solution.



- 1.2.2 Write down the name of apparatus **Q** in the above diagram. (1)
- 1.2.3 The following indicators are available for the titration:

INDICATOR	pH RANGE
A	3,1 – 4,4
B	6,0 – 7,6
C	8,3 – 10,0

- Which ONE of the above indicators (**A**, **B** or **C**) is most suitable to indicate the exact endpoint in this titration? Give a reason for the answer. (3)
- 1.2.4 During the titration, the technician uses distilled water to wash any sodium hydroxide spilled against the sides of the Erlenmeyer flask into the solution. Give a reason why the addition of distilled water to the Erlenmeyer flask will not influence the results. (1)
- 1.2.5 At the endpoint of the titration he finds that 21 cm^3 of a $0,2 \text{ mol}\cdot\text{dm}^{-3}$ sodium hydroxide solution has neutralised the EXCESS hydrochloric acid. Calculate the number of moles of hydrochloric acid in excess. (3)
- 1.2.6 The balanced equation for the reaction between hydrochloric acid and magnesium oxide is: $\text{MgO}(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{MgCl}_2(\text{aq}) + \text{H}_2\text{O}(\text{l})$

Calculate the percentage purity of the magnesium oxide. Assume that only the magnesium oxide in the $4,5 \text{ g}$ sample reacted with the acid. (5)

[21]

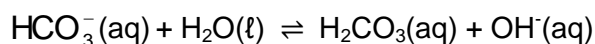
QUESTION 2

2.1 Sulphuric acid is a diprotic acid.

2.1.1 Define an *acid* in terms of the Lowry-Brønsted theory. (2)

2.1.2 Give a reason why sulphuric acid is referred to as a *diprotic acid*. (1)

2.2 The hydrogen carbonate ion can act as both an acid and a base. It reacts with water according to the following balanced equation:



2.2.1 Write down ONE word for the underlined phrase. (1)

2.2.2 $\text{HCO}_3^-(\text{aq})$ acts as base in the above reaction. Write down the formula of the conjugate acid of $\text{HCO}_3^-(\text{aq})$. (1)

2.3 A learner accidentally spills some sulphuric acid of concentration $6 \text{ mol}\cdot\text{dm}^{-3}$ from a flask on the laboratory bench. Her teacher tells her to neutralise the spilled acid by sprinkling sodium hydrogen carbonate powder onto it. The reaction that takes place is: (Assume that the H_2SO_4 ionises completely.)



The fizzing, due to the formation of carbon dioxide, stops after the learner has added 27 g sodium hydrogen carbonate to the spilled acid.

2.3.1 Calculate the volume of sulphuric acid that spilled. Assume that all the sodium hydrogen carbonate reacts with all the acid. (6)

The learner now dilutes some of the $6 \text{ mol}\cdot\text{dm}^{-3}$ sulphuric acid solution in the flask to $0,1 \text{ mol}\cdot\text{dm}^{-3}$.

2.3.2 Calculate the volume of the $6 \text{ mol}\cdot\text{dm}^{-3}$ sulphuric acid solution needed to prepare 1 dm^3 of the dilute acid. (2)

During a titration 25 cm^3 of the $0,1 \text{ mol}\cdot\text{dm}^{-3}$ sulphuric acid solution is added to an Erlenmeyer flask and titrated with a $0,1 \text{ mol}\cdot\text{dm}^{-3}$ sodium hydroxide solution.

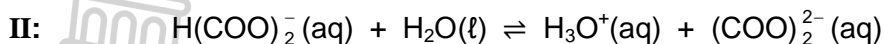
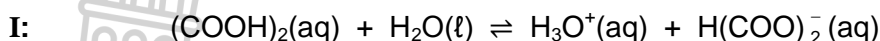
2.3.3 The learner uses bromothymol blue as indicator. What is the purpose of this indicator? (1)

2.3.4 Calculate the pH of the solution in the flask after the addition of 30 cm^3 of sodium hydroxide. The endpoint of the titration is not yet reached at this point. (8)

[22]

QUESTION 3

Anhydrous oxalic acid is an example of an acid that can donate two protons and thus ionises in two steps as represented by the equations below:



3.1 Write down:

3.1.1 ONE word for the underlined phrase in the above sentence (1)

3.1.2 The FORMULA of each of the TWO bases in **reaction II** (2)

3.1.3 The FORMULA of the substance that acts as ampholyte in **reactions I and II**. Give a reason for the answer. (2)

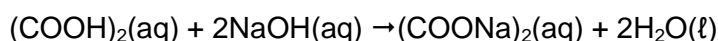
3.2 Give a reason why oxalic acid is a weak acid. (1)

3.3 A standard solution of $(\text{COOH})_2$ of concentration $0,20 \text{ mol}\cdot\text{dm}^{-3}$ is prepared by dissolving a certain amount of $(\text{COOH})_2$ in water in a 250 cm^3 volumetric flask.

Calculate the mass of $(\text{COOH})_2$ needed to prepare the standard solution. (4)

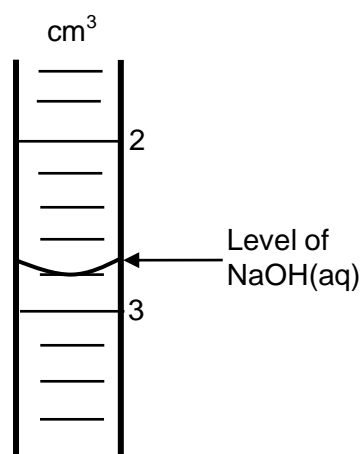
3.4 During a titration 25 cm^3 of the standard solution of $(\text{COOH})_2$ prepared in QUESTION 7.3 is neutralised by a sodium hydroxide solution from a burette.

The balanced equation for the reaction is:

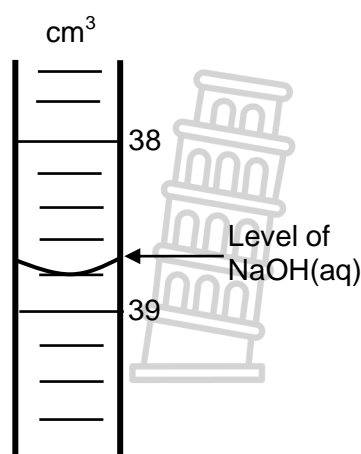


The diagrams below show the burette readings before the titration commenced and at the endpoint respectively.

Before the titration



At the endpoint



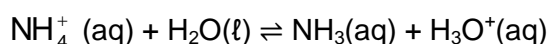
3.4.1 Use the burette readings and calculate the concentration of the sodium hydroxide solution. (5)

3.4.2 Write down a balanced equation that explains why the solution has a pH greater than 7 at the endpoint. (3)

[18]

QUESTION 4

- 4.1 Ammonium chloride crystals, $\text{NH}_4\text{Cl}(\text{s})$, dissolve in water to form ammonium and chloride ions. The ammonium ions react with water according to the balanced equation below:



4.1.1 Write down the name of the process described by the underlined sentence. (1)

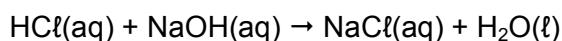
4.1.2 Is ammonium chloride ACIDIC or BASIC in aqueous solution? Give a reason for the answer. (2)

- 4.2 A certain fertiliser consists of 92% ammonium chloride. A sample of mass x g of this fertiliser is dissolved in 100 cm^3 of a $0,10 \text{ mol}\cdot\text{dm}^{-3}$ sodium hydroxide solution, $\text{NaOH}(\text{aq})$. The NaOH is in excess. The balanced equation for the reaction is:



4.2.1 Calculate the number of moles of sodium hydroxide in which the sample is dissolved. (3)

During a titration, 25 cm^3 of the excess sodium hydroxide solution is titrated with a $0,11 \text{ mol}\cdot\text{dm}^{-3}$ hydrochloric acid solution, $\text{HCl}(\text{aq})$. At the endpoint it is found that $14,55 \text{ cm}^3$ of the hydrochloric acid was used to neutralise the sodium hydroxide solution according to the following balanced equation:



4.2.2 Calculate the mass x (in grams) of the fertiliser sample used. (8)

- 4.3 Calculate the pH of a $0,5 \text{ mol}\cdot\text{dm}^{-3}$ sodium hydroxide solution at $25 \text{ }^\circ\text{C}$. (4) **[18]**

QUESTION 5

5.1 Define an *acid* in terms of the Lowry-Brønsted theory. (2)

5.2 Carbonated water is an aqueous solution of carbonic acid, H_2CO_3 . $\text{H}_2\text{CO}_3(\text{aq})$ ionises in two steps when it dissolves in water.

5.2.1 Write down the FORMULA of the conjugate base of $\text{H}_2\text{CO}_3(\text{aq})$. (1)

5.2.2 Write down a balanced equation for the first step in the ionisation of carbonic acid. (3)

5.2.3 The pH of a carbonic acid solution at $25 \text{ }^\circ\text{C}$ is 3,4. Calculate the hydroxide ion concentration in the solution. (5)

5.3 **X** is a monoprotic acid.

5.3.1 State the meaning of the term *monoprotic*. (1)

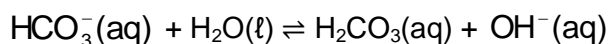
5.3.2 A sample of acid **X** is titrated with a standard sodium hydroxide solution using a suitable indicator. At the endpoint it is found that 25 cm^3 of acid **X** is neutralised by $27,5 \text{ cm}^3$ of the sodium hydroxide solution of concentration $0,1 \text{ mol}\cdot\text{dm}^{-3}$. Calculate the concentration of acid **X**. (5)

5.3.3 The concentration of H_3O^+ ions in the sample of acid **X** is $2,4 \times 10^{-4} \text{ mol}\cdot\text{dm}^{-3}$. Is acid **X** a WEAK or a STRONG acid? Explain the answer by referring to the answer in QUESTION 5.3.2. (3)

[20]

QUESTION 6

- 6.1 Hydrogen carbonate ions react with water according to the following balanced equation:

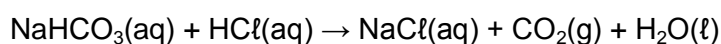


6.1.1 Define an *acid* according to the Lowry-Brønsted theory. (2)

6.1.2 Write down the FORMULAE of the two acids in the equation above. (2)

6.1.3 Write down the formula of a substance in the reaction above that can act as an ampholyte. (1)

- 6.2 During an experiment 0,50 dm³ of a 0,10 mol·dm⁻³ HCl solution is added to 0,80 dm³ of a NaHCO₃ solution of concentration 0,25 mol·dm⁻³. The balanced equation for the reaction is:



Calculate the:

6.2.1 Concentration of the hydroxide ions in the solution on completion of the reaction (8)

6.2.2 pH of the solution on completion of the reaction (4)
[17]

QUESTION 7

- 7.1 A learner dissolves ammonium chloride (NH₄Cl) crystals in water and measures the pH of the solution.

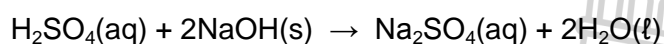
7.1.1 Define the term *hydrolysis* of a salt. (2)

7.1.2 Will the pH of the solution be GREATER THAN, SMALLER THAN or EQUAL TO 7? Write a relevant equation to support your answer. (3)

- 7.2 A sulphuric acid solution is prepared by dissolving 7,35 g of H₂SO₄(ℓ) in 500 cm³ of water.

7.2.1 Calculate the number of moles of H₂SO₄ present in this solution. (2)

Sodium hydroxide (NaOH) pellets are added to the 500 cm³ H₂SO₄ solution. The balanced equation for the reaction is:



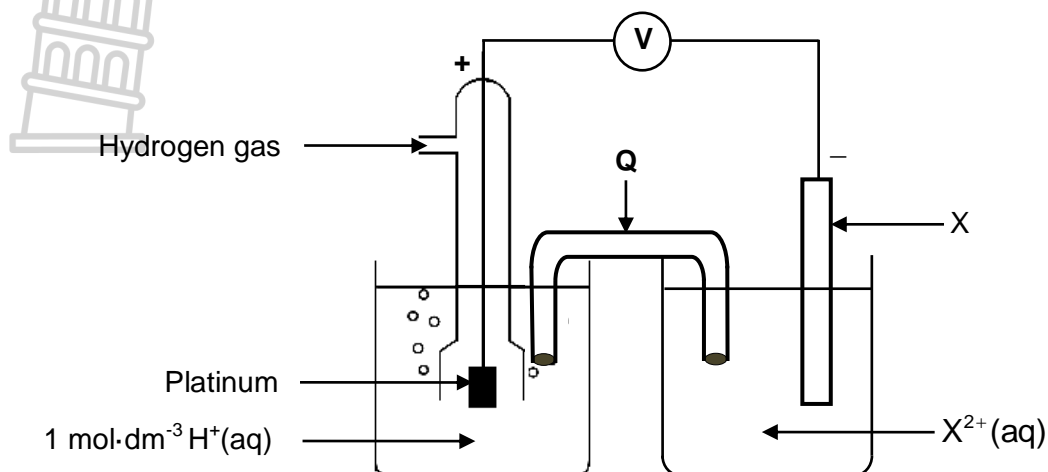
After completion of the reaction, the pH of the solution was found to be 1,3. Assume complete ionisation of H₂SO₄.

7.2.2 Calculate the mass of NaOH added to the H₂SO₄ solution. Assume that the volume of the solution does not change. (9)
[16]

GALVANIC CELLS

QUESTION 1

A standard electrochemical cell is set up using a standard hydrogen half-cell and a standard $X|X^{2+}$ half-cell as shown below. A voltmeter connected across the cell, initially registers 0,31 V.



- 1.1 Besides concentration write down TWO conditions needed for the hydrogen half-cell to function under standard conditions. (2)
- 1.2 Give TWO reasons, besides being a solid, why platinum is suitable to be used as electrode in the above cell. (2)
- 1.3 Write down the:
 - 1.3.1 NAME of component Q (1)
 - 1.3.2 Standard reduction potential of the $X|X^{2+}$ half-cell (1)
 - 1.3.3 Half-reaction that takes place at the cathode of this cell (2)
- 1.4 The hydrogen half-cell is now replaced by a $M|M^{2+}$ half-cell. The cell notation of this cell is:

$$M(s) | M^{2+}(aq) || X^{2+}(aq) | X(s)$$

The initial reading on the voltmeter is now 2,05 V.

 - 1.4.1 Identify metal M. Show how you arrived at the answer. (5)
 - 1.4.2 Is the cell reaction EXOTHERMIC or ENDOTHERMIC? (1)
- 1.5 The reading on the voltmeter becomes zero after using this cell for several hours. Give a reason for this reading by referring to the cell reaction. (1)

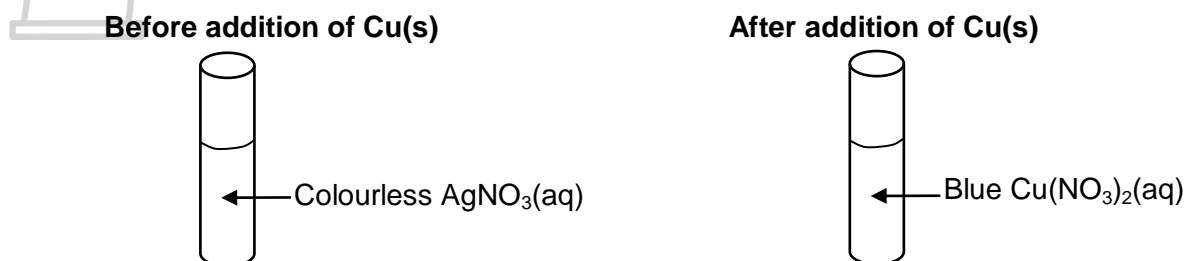
[15]

QUESTION 2

A learner conducts two experiments to investigate the reaction between copper (Cu) and a silver nitrate solution, $\text{AgNO}_3(\text{aq})$.

EXPERIMENT 1

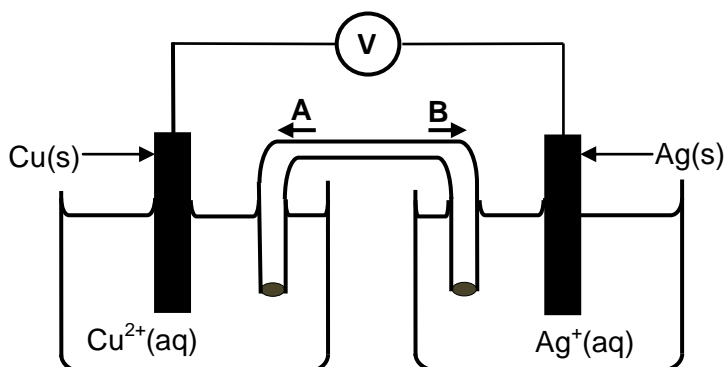
The learner adds a small amount of copper (Cu) powder to a test tube containing silver nitrate solution, $\text{AgNO}_3(\text{aq})$. The solution changes from colourless to blue after a while.



- 2.1 Define the term *oxidising agent*. (2)
- 2.2 Explain why the solution turns blue by referring to the relative strength of oxidising agents. (4)

EXPERIMENT 2

The learner now sets up a galvanic cell as shown below. The cell functions under standard conditions.

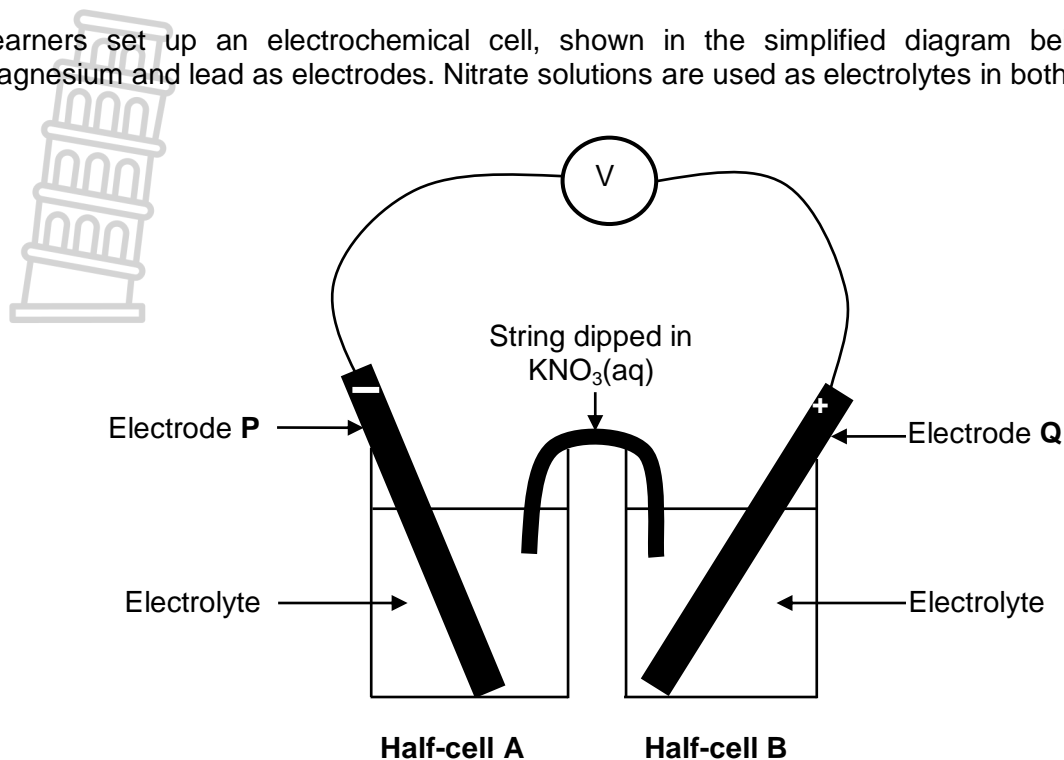


- 2.3 Write down the energy conversion that takes place in this cell. (1)
- 2.4 In which direction (**A** or **B**) will ANIONS move in the salt bridge? (1)
- 2.5 Calculate the emf of the above cell under standard conditions. (4)
- 2.6 Write down the balanced equation for the net cell reaction that takes place in this cell. (3)
- 2.7 How will the addition of 100 cm^3 of a $1 \text{ mol} \cdot \text{dm}^{-3}$ silver nitrate solution to the silver half-cell influence the initial emf of this cell? Write down only INCREASES, DECREASES or REMAINS THE SAME. (1)

[16]

QUESTION 3

Learners set up an electrochemical cell, shown in the simplified diagram below, using magnesium and lead as electrodes. Nitrate solutions are used as electrolytes in both half-cells.



- 3.1 What type of reaction (NEUTRALISATION, REDOX or PRECIPITATION) takes place in this cell? (1)
- 3.2 Which electrode, **P** or **Q**, is magnesium? Give a reason for the answer. (2)
- 3.3 Write down the:
 - 3.3.1 Standard conditions under which this cell functions (2)
 - 3.3.2 Cell notation for this cell (3)
 - 3.3.3 NAME or FORMULA of the oxidising agent in the cell (1)
- 3.4 Calculate the initial emf of the cell above under standard conditions. (4)
- 3.5 How will the voltmeter reading change if the: (Write down only INCREASES, DECREASES or REMAINS THE SAME.)
 - 3.5.1 Size of electrode **P** is increased (1)
 - 3.5.2 Initial concentration of the electrolyte in half-cell **B** is increased (1)

[15]

QUESTION 4

Learners are given the following two unknown half-cells:

Half-cell 1: $Q^{2+}(aq) | Q(s)$

Half-cell 2: $Pt | R_2(g) | R^-(aq)$

During an investigation to identify the two half-cells, the learners connect each half-cell alternately to a $Cd^{2+}(aq) | Cd(s)$ half-cell under standard conditions. For each combination of two half-cells, they write down the net cell reaction and measure the cell potential. The results obtained for the two half-cell combinations are given in the table below.

COMBINATION	NET CELL REACTION	CELL POTENTIAL
I	$Q^{2+}(aq) + Cd(s) \rightarrow Cd^{2+}(aq) + Q(s)$	0,13 V
II	$R_2(g) + Cd(s) \rightarrow Cd^{2+}(aq) + 2R^-(aq)$	1,76 V

- 4.1 Write down THREE conditions needed for these cells to function as standard cells. (3)
- 4.2 For **Combination I**, identify:
- 4.2.1 The anode of the cell (1)
- 4.2.2 **Q** by using a calculation (5)
- 4.3 For **Combination II**, write down the:
- 4.3.1 Oxidation half-reaction (2)
- 4.3.2 NAME or FORMULA of the metal used in the cathode compartment (1)
- 4.4 Arrange the following species in order of INCREASING oxidising ability:
 Q^{2+} ; R_2 ; Cd^{2+}
- Explain fully how you arrived at the answer. A calculation is NOT required. (4)

[16]**QUESTION 5**

An electrochemical cell consisting of half-cells **A** and **B** is assembled under standard conditions as shown below.

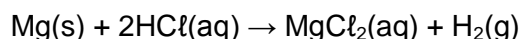
Half-cell A	$Pt, Cl_2(101,3 \text{ kPa}) Cl^-(1 \text{ mol}\cdot\text{dm}^{-3})$
Half-cell B	$Mg^{2+}(1 \text{ mol}\cdot\text{dm}^{-3}) Mg(s)$

- 5.1 At which half-cell, **A** or **B**, are electrons released into the external circuit? (1)
- 5.2 Write down the:
- 5.2.1 Reduction half-reaction that takes place in this cell (2)
- 5.2.2 NAME or FORMULA of the substance whose oxidation number DECREASES (1)
- 5.3 Calculate the initial cell potential of this cell when it is in operation. (4)
- 5.4 Write down an observation that will be made in half-cell **B** as the cell operates. Give a reason for the answer. (2)

[10]

QUESTION 6

Magnesium (Mg) reacts with a dilute hydrochloric acid solution, HCl(aq), according to the following balanced equation:



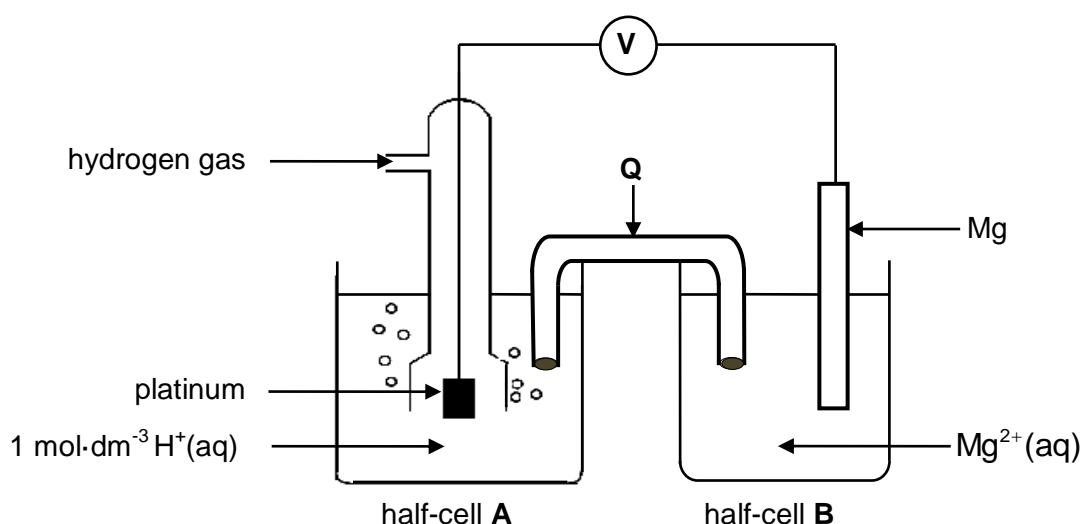
6.1 Give a reason why the reaction above is a redox reaction. (1)

6.2 Write down the FORMULA of the oxidising agent in the reaction above. (1)

It is found that silver does not react with the hydrochloric acid solution.

6.3 Refer to the relative strengths of reducing agents to explain this observation. (3)

The reaction of magnesium with hydrochloric acid is used in an electrochemical cell, as shown in the diagram below. The cell functions under standard conditions.



6.4 What is the function of platinum in the cell above? (1)

6.5 Write down the:

6.5.1 Energy conversion that takes place in this cell (1)

6.5.2 Function of Q (1)

6.5.3 Half-reaction that takes place at the cathode (2)

6.5.4 Cell notation of this cell (3)

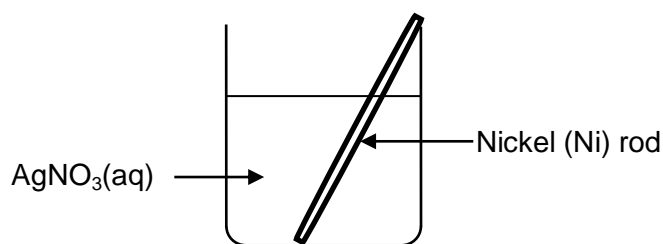
6.6 Calculate the initial emf of this cell. (4)

6.7 How will the addition of concentrated acid to half-cell A influence the answer to QUESTION 6.6? Choose from INCREASES, DECREASES or REMAINS THE SAME. (1)

[18]

QUESTION 7

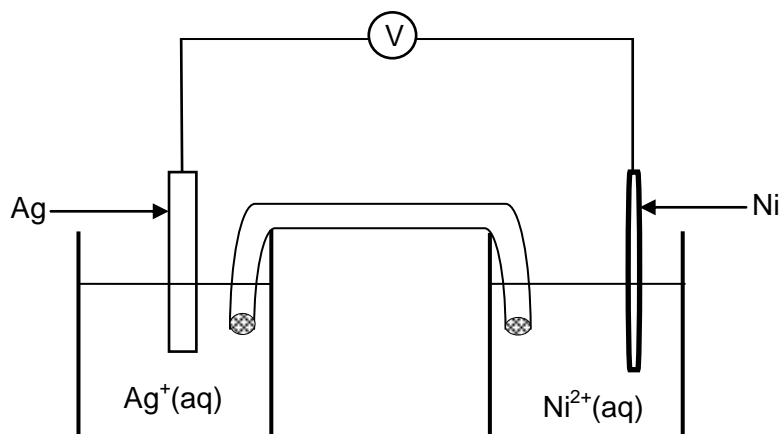
7.1 A nickel (Ni) rod is placed in a beaker containing a silver nitrate solution, $\text{AgNO}_3(\text{aq})$ and a reaction takes place.



Write down the:

- 7.1.1 NAME or FORMULA of the electrolyte (1)
- 7.1.2 Oxidation half-reaction that takes place (2)
- 7.1.3 Balanced equation for the net (overall) redox reaction that takes place (3)

7.2 A galvanic cell is now set up using a nickel half-cell and a silver half-cell.



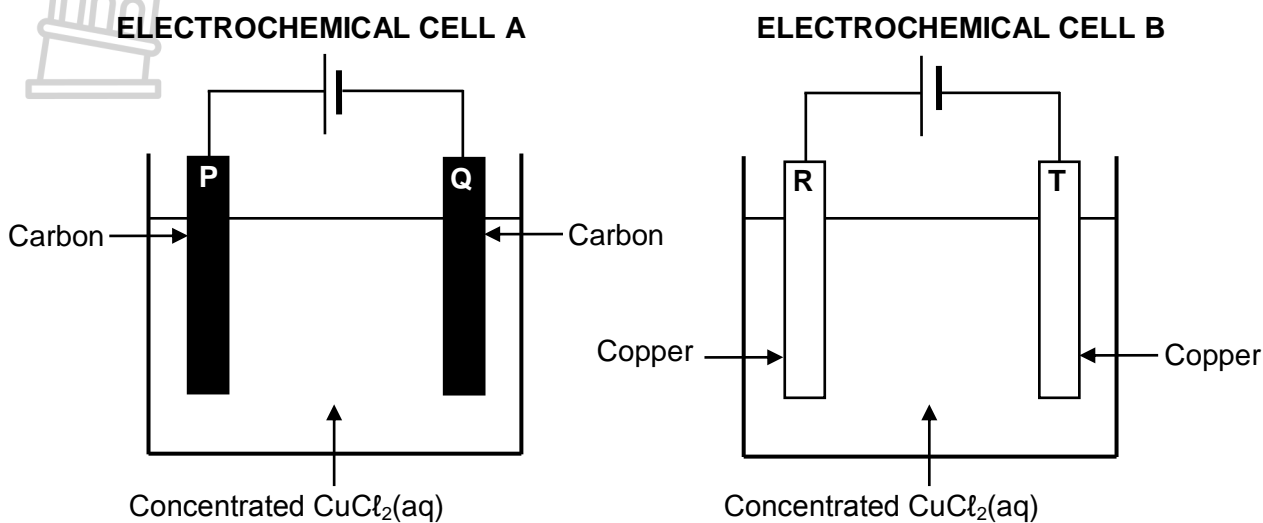
- 7.2.1 Which electrode (**Ni** or **Ag**) must be connected to the negative terminal of the voltmeter? Give a reason for the answer. (2)
- 7.2.2 Write down the cell notation for the galvanic cell above. (3)
- 7.2.3 Calculate the initial reading on the voltmeter if the cell functions under standard conditions. (4)
- 7.2.4 How will the voltmeter reading in QUESTION 7.2.3 be affected if the concentration of the silver ions is increased? Choose from INCREASES, DECREASES or REMAINS THE SAME. (1)

[16]

ELECTROLYTIC CELLS

QUESTION 1

The simplified diagrams below represent two electrochemical cells, **A** and **B**. A concentrated copper(II) chloride solution is used as electrolyte in both cells.

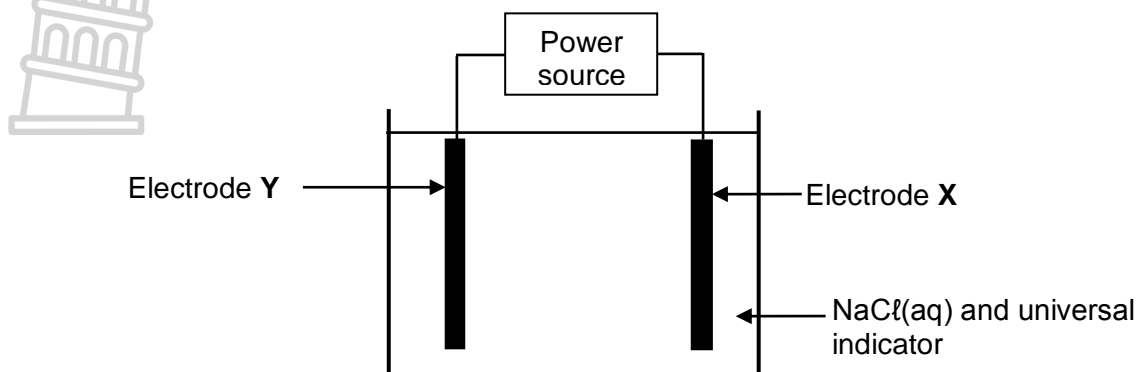


- 1.1 Are **A** and **B** ELECTROLYTIC or GALVANIC cells? (1)
 - 1.2 Which of the electrodes (**P**, **Q**, **R** or **T**) will show a mass increase? Write down a half-reaction to motivate the answer. (4)
 - 1.3 Write down the NAME or FORMULA of the product formed at:
 - 1.3.1 Electrode **P** (1)
 - 1.3.2 Electrode **R** (1)
 - 1.4 Fully explain the answer to QUESTION 1.3.2 by referring to the relative strengths of the reducing agents involved. (3)
- [10]**



QUESTION 2

The apparatus below is used to demonstrate the electrolysis of a concentrated sodium chloride solution. Both electrodes are made of carbon. A few drops of universal indicator are added to the electrolyte. The equation for the net cell reaction is:



Initially the solution has a green colour. Universal indicator becomes red in acidic solutions and purple in alkaline solutions.

2.1 Define the term *electrolyte*. (2)

When the power source is switched on, the colour of the electrolyte around electrode Y changes from green to purple.

2.2 Write down the:

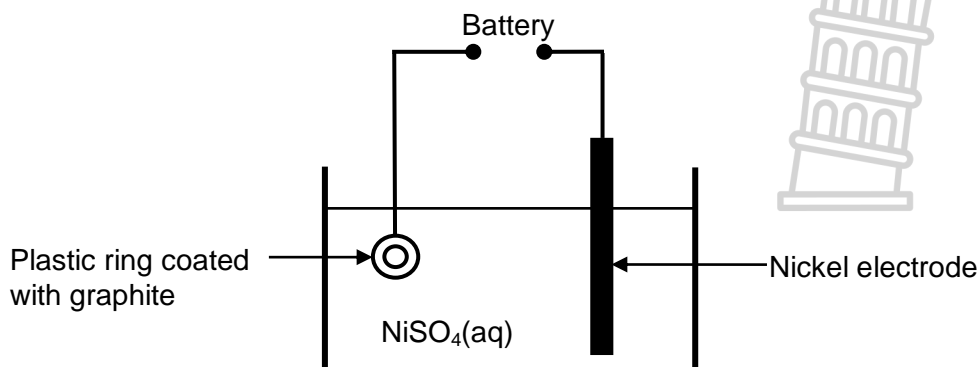
2.2.1 Half-reaction that takes place at electrode Y (2)

2.2.2 NAME or FORMULA of the gas released at electrode X (1)

2.3 Refer to the Table of Standard Reduction Potentials to explain why hydrogen gas, and not sodium, is formed at the cathode of this cell. (2) [7]

QUESTION 3

The diagram below shows a simplified electrolytic cell that can be used to electroplate a plastic ring with nickel. Prior to electroplating the ring is covered with a graphite layer.



3.1 Define the term *electrolyte*. (2)

3.2 Give ONE reason why the plastic ring must be coated with graphite prior to electroplating. (1)

3.3 Write down the:

3.3.1 Half-reaction that occurs at the plastic ring (2)

3.3.2 NAME or FORMULA of the reducing agent in the cell.
Give a reason for the answer. (2)

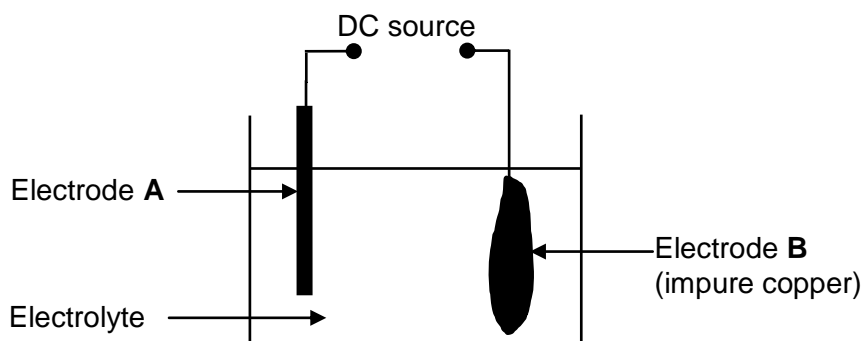
3.4 Which electrode, the **RING** or **NICKEL**, is the cathode? Give a reason for the answer. (2)

The nickel electrode is now replaced with a carbon rod.

3.5 How will the concentration of the electrolyte change during electroplating? Write down only INCREASES, DECREASES or NO CHANGE. Give a reason for the answer. (2)
[11]

QUESTION 4

The simplified diagram below represents an electrochemical cell used for the purification of copper.



4.1 Define the term *electrolysis*. (2)

4.2 Give a reason why a direct-current (DC) source is used in this experiment. (1)

4.3 Write down the half-reaction which takes place at electrode **A**. (2)

4.4 Due to small amounts of zinc impurities in the impure copper, the electrolyte becomes contaminated with Zn^{2+} ions.

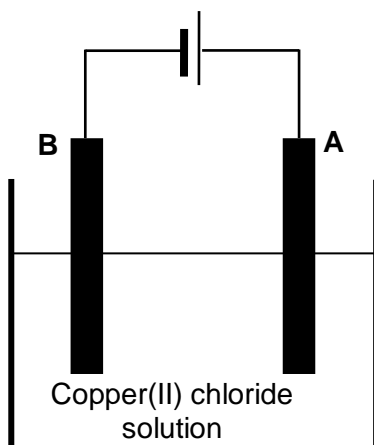
Refer to the attached Table of Standard Reduction Potentials to explain why the Zn^{2+} ions will not influence the purity of the copper obtained during this process. (3)

4.5 After the purification of the impure copper was completed, it was found that $2,85 \times 10^{-2}$ moles of copper were formed.

The initial mass of electrode **B** was 2,0 g. Calculate the percentage of copper that was initially present in electrode **B**. (4)
[12]

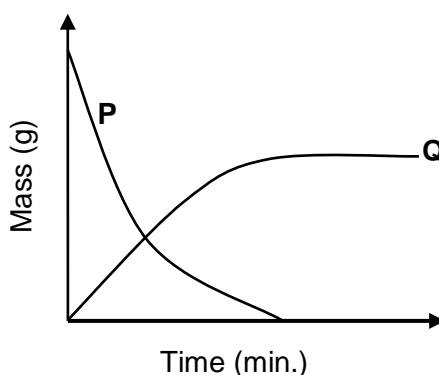
QUESTION 5

The electrochemical cell below is set up to demonstrate the purification of copper.



5.1 Write down the type of electrochemical cell illustrated above. (1)

The graphs below show the change in mass of the electrodes whilst the cell is in operation.



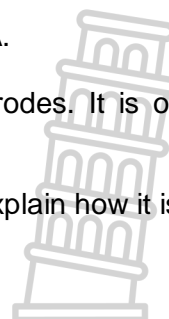
5.2 Define a *reducing agent* in terms of electron transfer. (2)

5.3 Which graph represents the change in mass of electrode **A**? (1)

5.4 Write down the half-reaction that takes place at electrode **A**. (2)

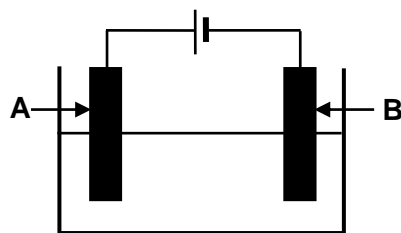
5.5 Electrodes **A** and **B** are now replaced by graphite electrodes. It is observed that chlorine gas (Cl_2) is released at one of the electrodes.

At which electrode (**A** or **B**) is chlorine gas formed? Fully explain how it is formed. (3) [9]



QUESTION 6

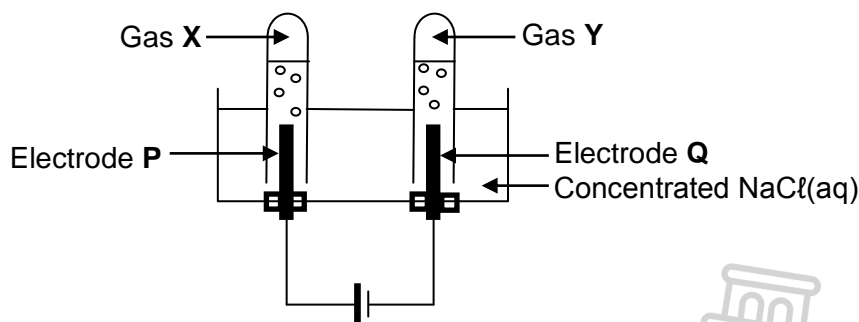
The diagram below shows an electrochemical cell used to purify copper. A solution that conducts electricity is used in the cell.



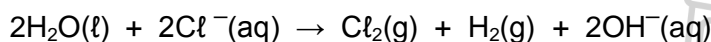
- 6.1 Write down:
- 6.1.1 ONE word for the underlined phrase above the diagram (1)
 - 6.1.2 The type of electrochemical cell illustrated above (1)
- 6.2 In which direction (**from A to B** or **from B to A**) will electrons flow in the external circuit? (1)
- 6.3 Which electrode (**A** or **B**) is the:
- 6.3.1 Cathode (1)
 - 6.3.2 Impure copper (1)
- 6.4 How will the mass of electrode **A** change as the reaction proceeds? Choose from INCREASES, DECREASES or REMAINS THE SAME. Give a reason for the answer. (2)
- [7]**

QUESTION 7

In the electrochemical cell below, carbon electrodes are used during the electrolysis of a concentrated sodium chloride solution.



The balanced equation for the net (overall) cell reaction is:



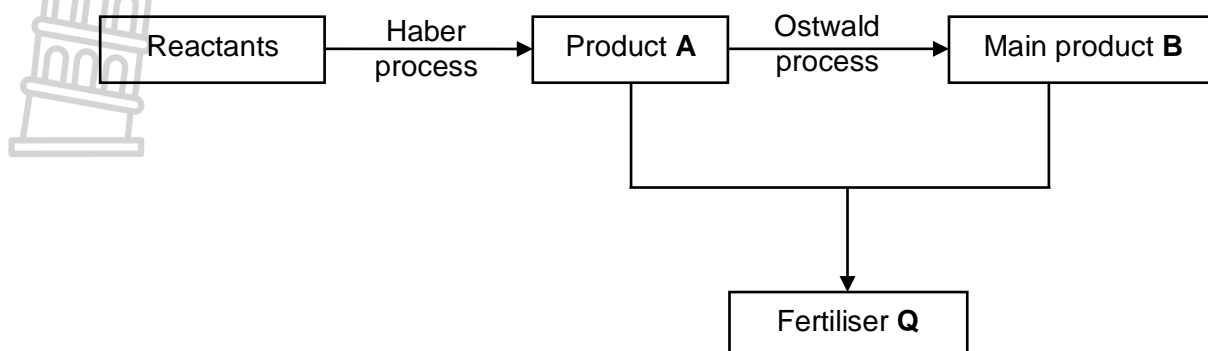
- 7.1 Is the reaction EXOTHERMIC or ENDOTHERMIC? (1)
 - 7.2 Is electrode **P** the ANODE or the CATHODE? Give a reason for the answer. (2)
 - 7.3 Write down the:
 - 7.3.1 NAME or FORMULA of gas **X** (1)
 - 7.3.2 NAME or FORMULA of gas **Y** (1)
 - 7.3.3 Reduction half-reaction (2)
 - 7.4 Is the solution in the cell ACIDIC or ALKALINE (BASIC) after completion of the reaction? Give a reason for the answer. (2)
- [9]**



FERTILISERS

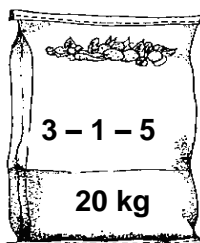
QUESTION 1

- 1.1 The flow diagram below shows the processes involved in the industrial preparation of fertiliser **Q**.



Write down the:

- 1.1.1 NAMES or FORMULAE of the reactants used in the Haber process (2)
- 1.1.2 Balanced equation for the formation of fertiliser **Q** (3)
- 1.2 The diagram below shows a bag of NPK fertiliser.

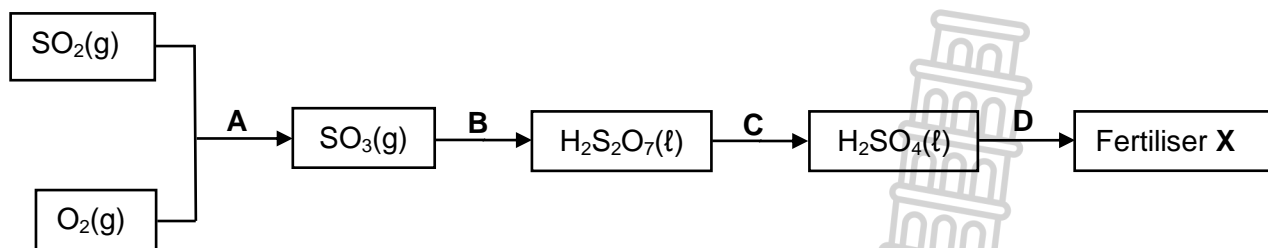


Calculate the mass of nitrogen in the bag.

(4)
[9]

QUESTION 2

Reactions **A**, **B**, **C** and **D** in the flow diagram below represent the manufacturing of Fertiliser **X**.



- 2.1 Write down the name of the industrial preparation of sulphuric acid. (1)
- 2.2 Write down the:
- 2.2.1 NAME or FORMULA of the catalyst used in reaction **A** (1)
- 2.2.2 Balanced equation for reaction **C** (3)
- 2.3 Ammonia is one of the reactants used in reaction **D** to make Fertiliser **X**. Write down:
- 2.3.1 A balanced equation for reaction **D** (3)
- 2.3.2 The NAME of Fertiliser **X** (1)

2.4 Two 50 kg bags, containing fertilisers **P** and **Q** respectively, are labelled as follows:

Fertiliser **P**: 5 : 2 : 3 (25)
 Fertiliser **Q**: 1 : 3 : 4 (20)

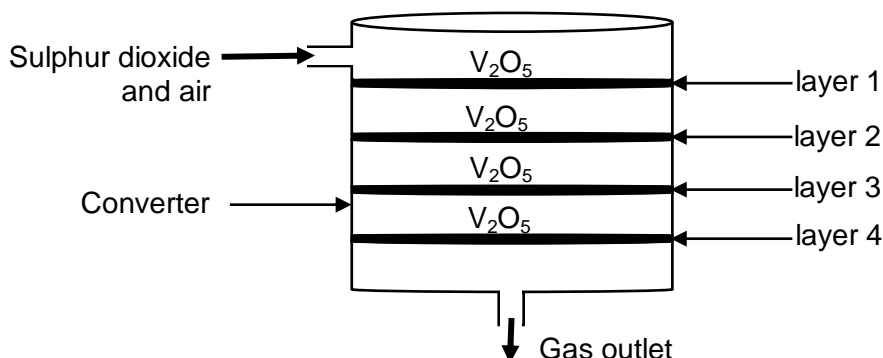
2.4.1 What do the numbers (25) and (20) on the labels represent? (1)

2.4.2 Using calculations, determine which fertiliser (**P** or **Q**) contains the greater mass of potassium. (4)
[14]

QUESTION 3

The industrial process for the preparation of sulphuric acid involves a series of stages.

The second stage in this process involves the conversion of sulphur dioxide into sulphur trioxide in a converter as illustrated below. In the converter the gases are passed over vanadium pentoxide (V_2O_5) placed in layers as shown below.



3.1 Write down the:

3.1.1 Balanced equation for the reaction taking place in the converter (3)

3.1.2 Function of the vanadium pentoxide (1)

The table below shows data obtained during the second stage.

VANADIUM PENTOXIDE LAYER	TEMPERATURE OF GAS BEFORE THE REACTION (°C)	TEMPERATURE OF GAS AFTER THE REACTION (°C)	PERCENTAGE OF REACTANT CONVERTED TO PRODUCT
1	450	600	66
2	450	518	85
3	450	475	93
4	450	460	99,5

3.2 Is the reaction in the second stage EXOTHERMIC or ENDOTHERMIC? Refer to the data in the table to give a reason for the answer. (2)

3.3 After the conversion at each layer the gases are cooled down to 450 °C. Fully explain why the gases must be cooled to this temperature. (3)

3.4 During the third stage sulphur trioxide is dissolved in sulphuric acid rather than in water to produce oleum.

3.4.1 Write down the FORMULA of oleum. (1)

3.4.2 Give a reason why sulphur trioxide is not dissolved in water. (1)

3.5 Sulphuric acid reacts with ammonia to form a fertiliser. Write down a balanced equation for this reaction. (3)
[14]

QUESTION 4

Ammonia is an important fertiliser. Large amounts are prepared from hydrogen and nitrogen in industry.

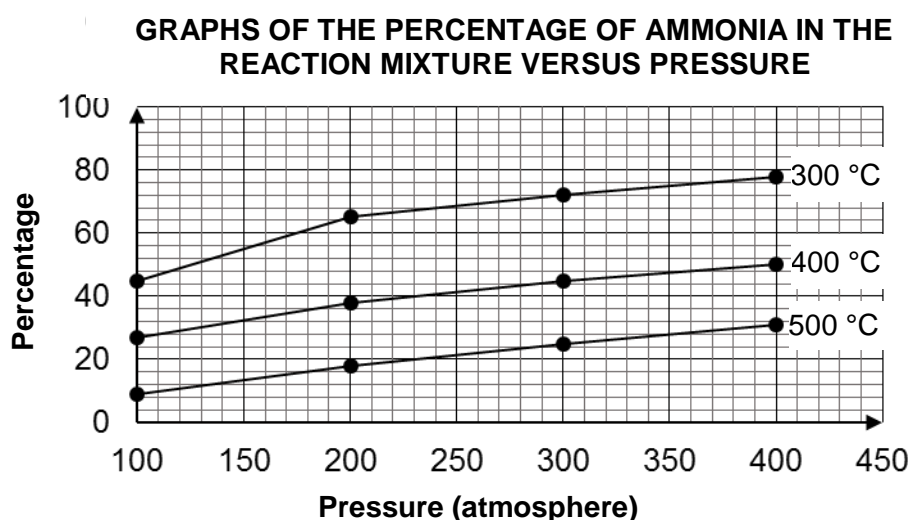
4.1 For the industrial preparation of ammonia, write down:

4.1.1 The name of the process used (1)

4.1.2 A balanced equation for the reaction that occurs (3)

4.1.3 The source of nitrogen (1)

4.2 The yield of ammonia changes with temperature and pressure during its industrial preparation. The graphs below show how the percentage of ammonia in the reaction mixture that leaves the reaction vessel varies under different conditions.



4.2.1 Use the appropriate graph to estimate the percentage of ammonia present in the reaction mixture at 240 atmosphere and 400 °C. (1)

4.2.2 State TWO advantages of using high pressure in the preparation of ammonia. (2)

4.2.3 The advantage of using a low temperature is the large percentage of ammonia formed. What is the disadvantage of using a low temperature? (1)

4.3 Ammonia is also used in the preparation of other fertilisers such as ammonium nitrate. Calculate the mass of nitrogen in a 50 kg bag of pure ammonium nitrate fertiliser. (3)

[12]

QUESTION 5

Ammonium nitrate is an important fertiliser. It is produced by reacting nitric acid with ammonia. Both nitric acid and ammonia are prepared on a large scale in industry.

5.1 Write down the name of the industrial preparation of nitric acid. (1)

5.2 The catalytic oxidation of ammonia is one of the steps in the process named in QUESTION 5.1. Write down the NAMES or FORMULAE of the TWO products formed in this step. (2)

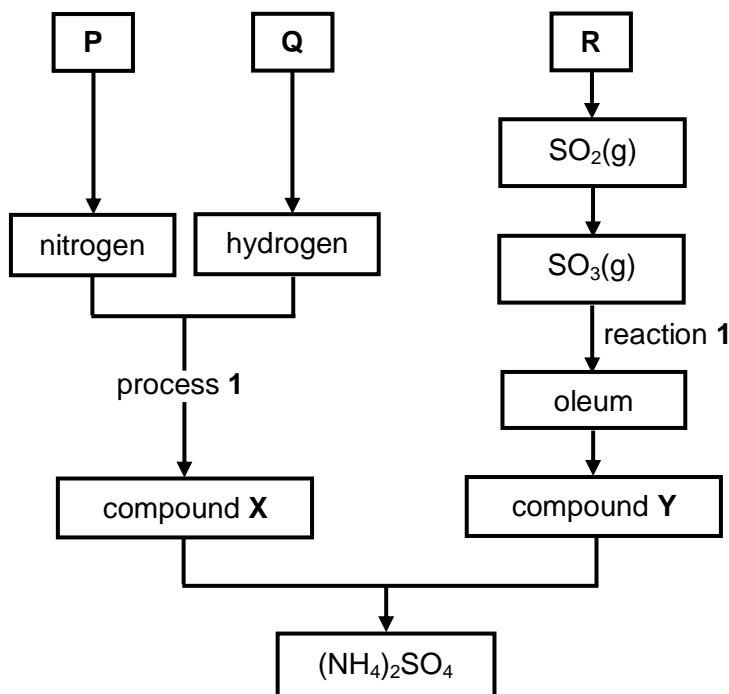
5.3 Write down a balanced equation for the preparation of ammonium nitrate. (3)

5.4 Calculate the mass, in kilogram, of ammonium nitrate that can be made from $6,8 \times 10^4$ kg of ammonia and excess nitric acid.
(One mole of ammonia produces one mole of ammonium nitrate.) (3)

5.5 Ammonium nitrate is often mixed with potassium chloride and ammonium phosphate.
Give a reason why it is mixed with these compounds. (1)
[10]

QUESTION 6

A chemical company produces ammonium sulphate, $(\text{NH}_4)_2\text{SO}_4$, starting from the raw materials **P**, **Q** and **R**, as shown in the flow diagram below.



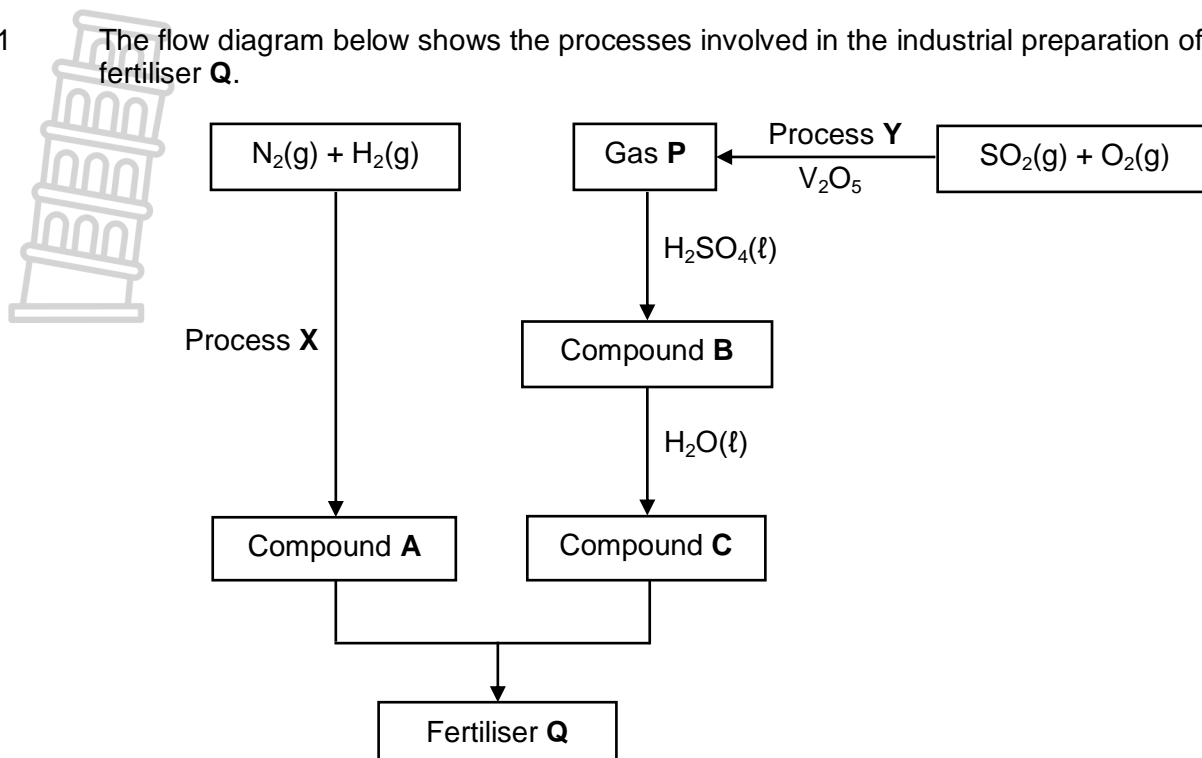
- 6.1 Write down the NAME of raw material:
- 6.1.1 **P** (1)
 - 6.1.2 **Q** (1)
 - 6.1.3 **R** (1)

- 6.2 Write down the:
- 6.2.1 NAME of process 1 (1)
 - 6.2.2 NAME of compound X (1)
 - 6.2.3 FORMULA of compound Y (1)
 - 6.2.4 Balanced equation for reaction 1 (3)

- 6.3 The company compares the nitrogen content of ammonium sulphate with that of ammonium nitrate, NH_4NO_3 .
- 6.3.1 Determine, by performing the necessary calculations, which ONE of the two fertilisers has the higher percentage of nitrogen per mass. (4)
 - 6.3.2 Write down the name of the process that should be included in the flow diagram above if the company wants to prepare ammonium nitrate instead of ammonium sulphate. (1)
- [14]**

QUESTION 7

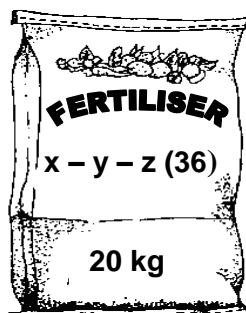
7.1 The flow diagram below shows the processes involved in the industrial preparation of fertiliser **Q**.



Write down the:

- 7.1.1 Name of process **X** (1)
- 7.1.2 Name of process **Y** (1)
- 7.1.3 NAME or FORMULA of gas **P** (1)
- 7.1.4 Balanced equation for the formation of compound **B** (3)
- 7.1.5 Balanced equation for the formation of fertiliser **Q** (4)

7.2 The diagram below shows a bag of NPK fertiliser of which the NPK ratio is unknown. It is found that the mass of nitrogen in the bag is 4,11 kg and the mass of phosphorus is 0,51 kg.



Calculate the NPK ratio of the fertiliser.

(4)
[14]

ANSWERS TO QUESTIONS

ORGANIC MOLECULES: NOMENCLATURE

QUESTION 1

2.1

2.1.1 B ✓

(1)

2.1.2 E ✓

(1)

2.1.3 F ✓

(1)

2.2

2.2.1 2-bromo-3-chloro-4-methylpentane

Marking criteria:

- Correct stem i.e. pentane. ✓
- All substituents correctly identified. ✓
- Substituents correctly numbered, in alphabetical order, hyphens and commas correctly used. ✓

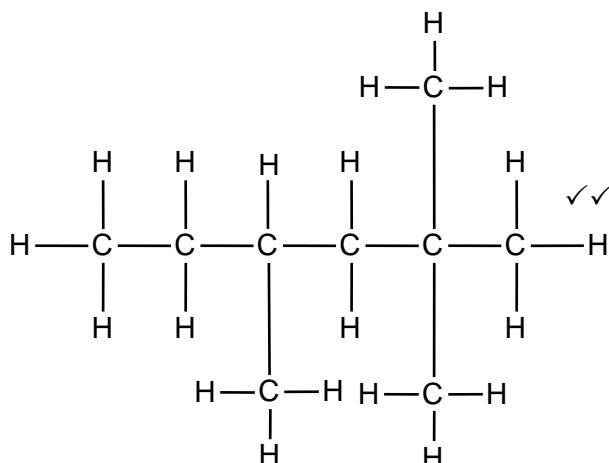
(3)

2.2.2 Ethene ✓

(1)

2.3

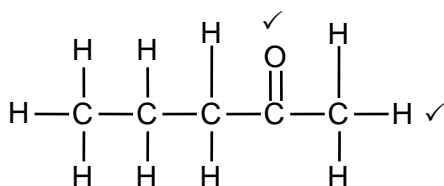
2.3.1

**Marking criteria:**

- Six saturated C atoms in longest chain i.e. hexane. ✓.
- Three methyl substituents on 2nd C and 4th C atom. ✓

(2)

2.3.2

**Marking criteria:**

- Whole structure correct: ✓✓
- Only functional group correct: ✓

(2)

2.4

2.4.1 Compounds with the same molecular formula ✓ but DIFFERENT functional groups. ✓

(2)

2.4.2 B & F ✓

(1)

[14]

QUESTION 2

2.1

2.1.1 Carboxyl group ✓

(1)

2.1.2 Ketones ✓

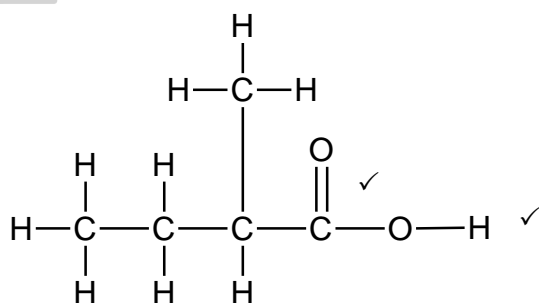
(1)

2.1.3 Addition ✓

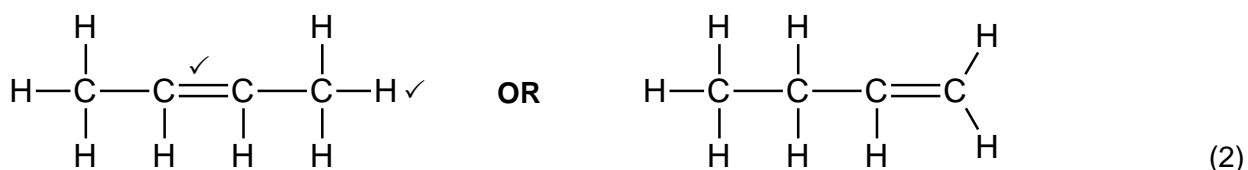
(1)

- 2.2
 2.2.1 Ethene ✓ (1)
 2.2.2 4-methylhexan-3-one ✓ (2)
 2.2.3 4-ethyl-2,2-dimethylhexane ✓ (2)
 2.3 Carbon dioxide / CO₂ ✓
 Water / H₂O ✓ (2)

2.4
 2.4.1



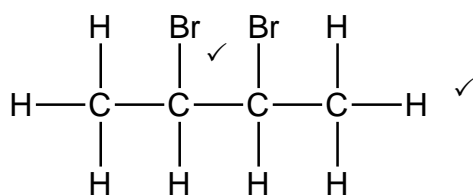
2.4.2



2.5

- 2.5.1 E ✓ (1)
 2.5.2 Substitution / halogenation / bromination ✓ (1)

2.5.3



[18]

QUESTION 3

- 3.1
 3.1.1 B ✓ (1)
 3.1.2 E ✓ (1)
 3.1.3 A ✓ (1)
 3.2
 3.2.1 4-chloro-2,5-dimethylheptane

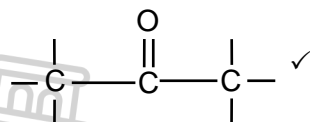
Marking criteria:

- Correct stem i.e. heptane. ✓
- All substituents (chloro and dimethyl) correctly identified. ✓ Substituents correctly numbered, in alphabetical order, hyphens and commas correctly used. ✓

(3)

- 3.2.2 2-methyl propan-1-ol ✓ (2)

3.2.3



(1)

3.3

3.3.1 Compounds with the same molecular formula ✓ but different POSITIONS of the functional group / side chain / substituents on parent chain. ✓

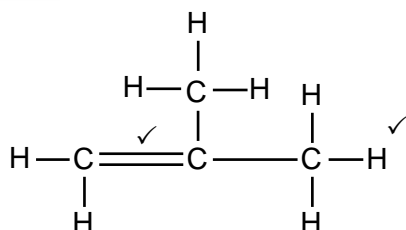
(2)

3.3.2 But-1-ene ✓✓ **AND** But-2-ene ✓✓

(4)

KKH

3.3.3



Marking criteria:

- Whole structure correct: ✓✓
- Only functional group correct: ✓

(2)

3.4

3.4.1 Cracking / elimination ✓

(1)

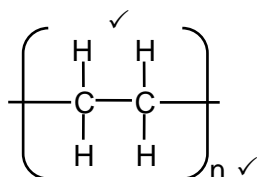
3.4.2 Ethene ✓

(1)

3.4.3 C_4H_{10} ✓

(1)

3.4.4 Polyethene ✓



(3)

[23]

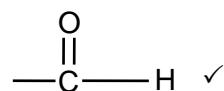
QUESTION 4

4.1

4.1.1 B ✓

(1)

4.1.2



(1)

4.1.3 C_nH_{2n-2} ✓

(1)

4.1.4 4-ethyl-5-methylhept-2-yne
OR
4-ethyl-5-methyl-2-heptyne

Marking criteria:

- 4-ethyl ✓
- 5-methyl ✓
- hept-2-yne / 2-heptyne ✓

(3)

4.1.5 Butan-2-one / 2-butanone / butanone

Marking criteria:

- Functional group ✓
- Whole name correct ✓

(2)

4.2

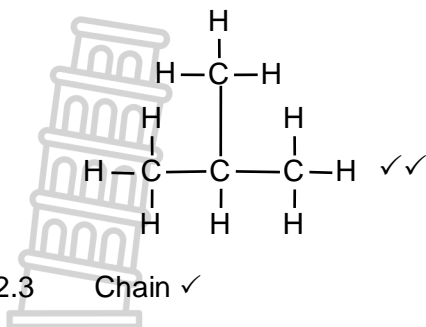
4.2.1 Alkanes ✓

(1)

4.2.2 2-methylpropane ✓

OR

methylpropane ✓



Marking criteria structural formula:

- Three carbons in longest chain. ✓
- Methyl group on second carbon. ✓

(4)

4.2.3 Chain ✓

(1)

4.3

4.3.1 Haloalkanes / Alkyl halides ✓

(1)

4.3.2 Substitution / halogenation / bromination ✓

(1)

[16]

QUESTION 5

5.1

5.1.1 Ketones ✓

(1)

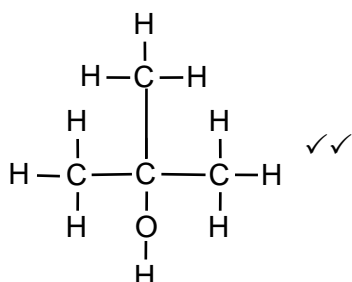
5.1.2 3,5-dichloro-4-methyloctane ✓

Marking criteria

- 3,5-dichloro ✓
- -4-methyl ✓
- octane ✓

(3)

5.1.3



Marking criteria

- Functional group (-OH) on second C atom. ✓
- Whole structure correct ✓

(2)

5.2

5.2.1 Acts as catalyst. / Increases the rate of reaction. / Act as dehydrating agent. ✓

(1)

5.2.2 Water / H₂O

(1)

5.2.3

$$\text{mol C} : \text{mol H} : \text{mol O}$$

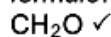
$$\frac{40}{12} : \frac{6,67}{1} : \frac{53,33}{16}$$

$$3,33 : 6,67 : 3,33$$

$$1 : 2 : 1$$

Empirical formula/Empiriese

formule:



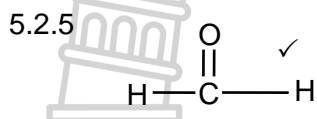
(5)

5.2.4 $M(\text{CH}_2\text{O}) = 30 \text{ g}\cdot\text{mol}^{-1}$ ✓
 Formula-units/*Formule-eenhede*:

$$\frac{60}{30} = 2 \quad \checkmark$$

Molecular formula/*Molekulêre formule*: $\text{C}_2\text{H}_4\text{O}_2$ ✓

(3)



(1)

5.2.6 Methyl ✓methanoate ✓

(2)

[19]

QUESTION 6

6.1

6.1.1 E ✓

(1)

6.1.2 C ✓

(1)

6.1.3 D ✓

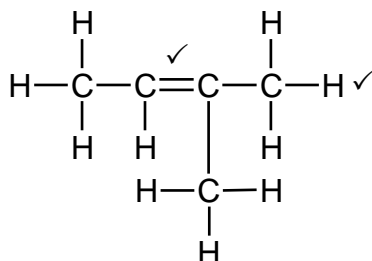
(1)

6.2

6.2.1 Pent-2✓-yne✓ OR 2-pentyne

(2)

6.2.2



Marking criteria:

- Whole structure correct: ✓✓
- Only functional group correct: ✓

(2)

6.2.3 2-methylbut-1-ene OR 3-methylbut-1-ene

Marking criteria:

- Correct stem i.e. but-1-ene/1-butene. ✓
- Only one type of substituent, methyl, correctly identified. ✓
- Entire name correct. ✓

(3)

6.3

6.3.1 Esters ✓

(1)

6.3.2 Sulphuric acid / H_2SO_4 ✓

(1)

6.3.3 Methyl ✓ propanoate ✓

(2)

[14]

QUESTION 7

7.1

7.1.1 A OR D ✓

(1)

7.1.2 B ✓

(1)

7.1.3 E ✓

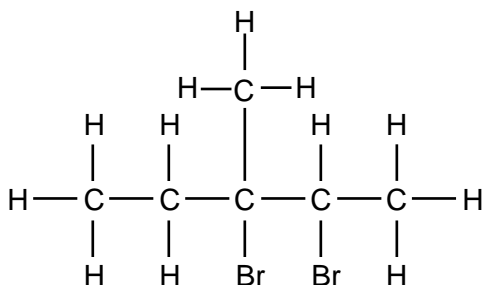
(1)

7.1.4 D ✓

(1)

7.2

7.2.1

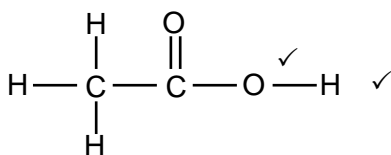


Marking criteria:

- Five C atoms in longest chain. ✓
- Two Br and one methyl substituents. ✓
- Whole structure correct. ✓

(3)

7.2.2

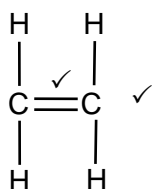


Marking criteria:

- Whole structure correct: ✓✓
- Only functional group correct: ✓

(2)

7.2.3



Marking criteria:

- Whole structure correct: ✓✓
- Only functional group correct: ✓

(2)

7.3

7.3.1 Hydrogen gas ✓

(1)

7.3.2 Addition / Hydrogenation ✓

(1)

[13]



ORGANIC MOLECULES: PHYSICAL PROPERTIES

QUESTION 1

1.1 Alkanes have ONLY single bonds. ✓

OR

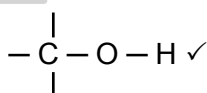
Alkanes have single bonds between C atoms.

OR

Alkanes have no double OR triple bonds OR multiple bonds. (1)

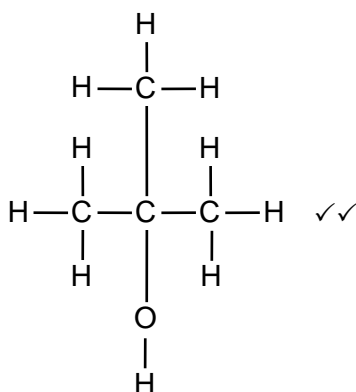
1.2

1.2.1



(1)

1.2.2



Marking criteria:

- - OH group on second C atom of longest chain. ✓
- Tertiary group consisting of four C atoms with methyl group on 2nd C atom. ✓

(2)

1.3

1.3.1

Criteria for investigative question:

The dependent and independent variables are stated. ✓

Ask a question about the relationship between the independent and dependent variables. ✓

Examples:

- How does an increase in chain length influence boiling point?
- What is the relationship between chain length and boiling point?

(2)

1.3.2

• **Structure:**

The chain length increases. ✓

• **Intermolecular forces:**

Increase in strength of intermolecular. ✓

• **Energy:**

More energy needed to overcome intermolecular forces. ✓

OR

• **Structure:**

From propane to methane the chain length decreases. ✓

• **Intermolecular forces:**

Decrease in strength of intermolecular. ✓

• **Energy:**

Less energy needed to overcome intermolecular forces. ✓

(3)

1.4

- Between propane molecules are London forces. ✓
- Between propan-1-ol molecules are London forces, dipole-dipole forces and hydrogen bonds. ✓
- Hydrogen bonds are stronger than London forces. ✓

(3)

[12]

QUESTION 2

- 2.1 C ✓ (1)
- 2.2 Chain length ✓ (1)
- 2.2.1 Chain length ✓ (1)
- 2.2.2 Boiling point ✓ (1)
- 2.3 London forces ✓ (1)
- 2.4 Higher than ✓ (1)
- 2.5 Lower than ✓ (1)
- Both compounds **D** and **E** have hydrogen bonding between molecules. ✓
 - Compound **D** has one site for hydrogen bonding whilst compound **E** has two sites for hydrogen bonding OR can form dimers. ✓
 - More energy needed to overcome intermolecular forces in compound **E**. ✓ (4)
- [9]**

QUESTION 3

- 3.1 Saturated/✓
ANY ONE:
- B has ONLY single bonds. ✓
 - B has single bonds between C atoms.
 - B has no double OR triple bonds OR multiple bonds. (2)
- 3.2 - 42 (°C) ✓ (1)
- 3.2.1 - 42 (°C) ✓ (1)
- 3.2.2 78 (°C) ✓ (1)
- 3.3
- Between molecules of C/propane are London forces / induced dipole forces. ✓
 - Between molecules of E/ethanol are (London forces / induced dipole forces and) hydrogen bonds. ✓
 - Hydrogen bonds / Forces between alcohol molecules are stronger. ✓ (3)
- 3.4 Decrease ✓
- From **A** to **D**:
- Chain length increases. ✓
 - Strength of intermolecular forces increases. ✓
 - More energy needed to overcome intermolecular forces. ✓ (4)
- 3.5 Higher than ✓ (1)
- [12]**

QUESTION 4

- 4.1 A bond OR an atom OR a group of atoms ✓ that determine(s) the (physical and chemical) properties of a group of organic compounds. ✓ (2)
- 4.2 **D** / ethanoic acid ✓ (2)
- 4.2.1 **D** / ethanoic acid ✓
Lowest vapour pressure. ✓ (2)
- 4.2.2 **A** / butane ✓ (1)

- 4.3
- Between molecules of **A** / butane are London forces. ✓
 - Between molecules of **B** / propan-2-one / ketones are dipole-dipole forces ✓ in addition to London forces.
 - Intermolecular forces in A are weaker than those in **B**. ✓
- (3)

- 4.4
- London forces / induced dipole forces / dipole-dipole forces. ✓
- OR**
- A** and **B** do not have hydrogen bonding. / **C** and **D** have hydrogen bonding.
- (1)

- 4.5 **OPTION 1**
- **D** has more sites for hydrogen bonding than **C**. / **D** forms dimers. ✓
 - **D** has stronger intermolecular forces. ✓
- (2)

4.6

OPTION 1

$$V(\text{CO}_2) = 4V(\text{C}_4\text{H}_{10}) = (4)(8) \checkmark = 32 \text{ cm}^3$$

$$V(\text{H}_2\text{O}) = 5V(\text{C}_4\text{H}_{10}) = (5)(8) \checkmark = 40 \text{ cm}^3$$

$$V(\text{O}_2 \text{ reacted}): V(\text{O}_2) = \frac{13}{2} V(\text{C}_4\text{H}_{10}) = \left(\frac{13}{2}\right)(8) \checkmark = 52 \text{ cm}^3$$

$$V(\text{O}_2 \text{ excess}): V(\text{O}_2) = 60 - 52 \checkmark = 8 \text{ cm}^3$$

$$V_{\text{tot}} = 32 + 40 + 8 = 80 \text{ cm}^3 \checkmark$$

OPTION 2

	C ₄ H ₁₀	O ₂	CO ₂	H ₂ O
Initial V (cm ³)	8	60	0	0
Change in V (cm ³)	8	52 ✓	32 ✓	40 ✓
Final V (cm ³)	0	8 ✓	32	40

Total volume = 8 + 32 + 40 = 80 cm³ ✓

(5)
[16]

QUESTION 5

- 5.1 Temperature ✓ at which the vapour pressure equals atmospheric pressure. ✓
- (2)

- 5.2 The stronger the intermolecular forces, the higher the boiling point. / The boiling point is proportional to the strength of intermolecular forces. ✓

NOTE: NOT DIRECTLY proportional!

(1)

- 5.3
- 5.3.1
- Between molecules of **A** / propane are London forces. ✓
 - Between molecules of **B** / propan-2-one are dipole-dipole forces ✓ in addition to London forces.
 - Intermolecular forces in A are weaker than those in **B**. ✓
- (3)

- 5.3.2
- Both **C** and **D**: hydrogen bonding ✓
 - **D** has two / more sites for hydrogen bonding. / **D** forms dimers. ✓
 - **D** has stronger intermolecular forces than **C**. ✓
- (3)

- 5.4 Liquid ✓
- (1)

[10]

QUESTION 6

6.1 Temperature ✓ at which the vapour pressure equals atmospheric pressure. ✓ (2)

6.2	Criteria for conclusion:	
	Dependent and independent variables correctly identified.	✓
	Relationship between the independent and dependent variables correctly stated.	✓

Examples:

- Boiling point increases with increase in chain length.
- Boiling point decreases with decrease in chain length.
- Boiling point is proportional to chain length.

NOTE: Boiling point is **NOT DIRECTLY** proportional to chain length. (2)

6.3

6.3.1 P ✓ (1)

6.3.2 R ✓ (1)

- 6.4
- Between alkane molecules are London forces. ✓
 - In addition to London forces and dipole-dipole forces each alcohol molecule has one site for hydrogen bonding. ✓
 - In addition to London forces and dipole-dipole forces each carboxylic acid molecule has two sites for hydrogen bonding. ✓
 - Intermolecular forces in carboxylic acids are stronger than intermolecular forces in alkanes and alcohols. ✓
 - More energy is needed to overcome intermolecular forces in carboxylic acids than in the other two compounds. ✓
- (5)
[11]

QUESTION 7

7.1 Compounds with the same molecular formula ✓ but different structural formulae. ✓ (2)

7.2 Chain ✓ (1)

- 7.3 **From A to C:**
- **Structure:**
Less branched / less compact / larger surface area (over which intermolecular forces act). ✓
 - **Intermolecular forces:**
Stronger intermolecular forces. ✓
 - **Energy:**
More energy needed to overcome intermolecular forces. ✓
- (3)

7.4 A ✓
Lowest boiling point. ✓ (2)

7.5 $C_5H_{12} + 8O_2 \rightarrow 5CO_2 + 6H_2O$ ✓ Bal ✓ (3)
[11]

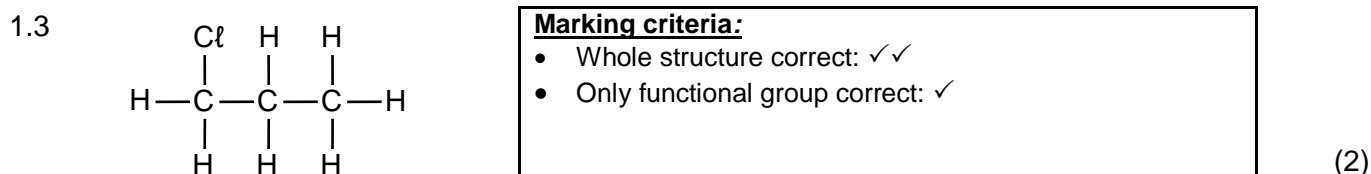
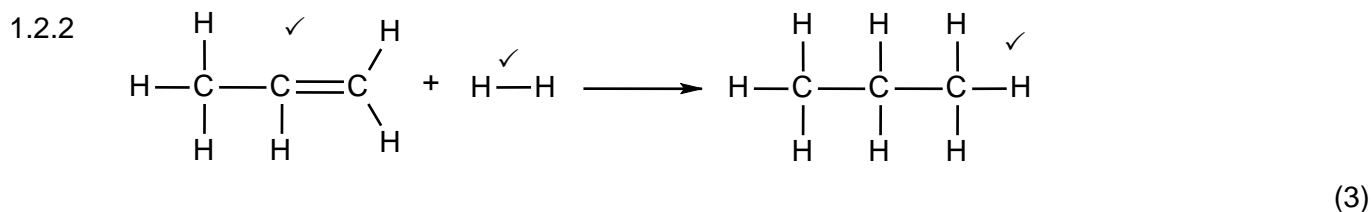
ORGANIC MOLECULES: ORGANIC REACTIONS

QUESTION 1

1.1 Substitution / chlorination / halogenation ✓ (1)

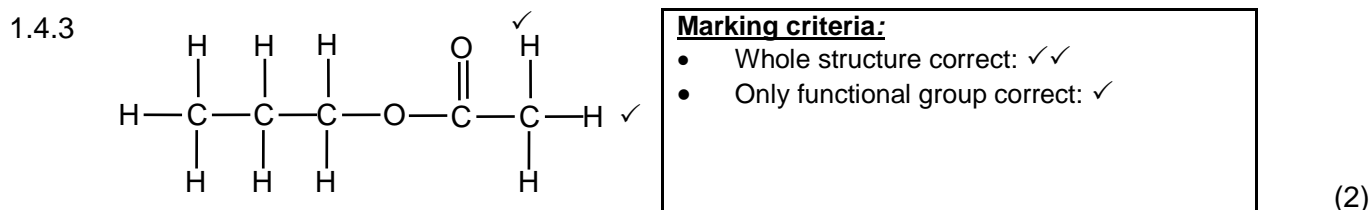
1.1.2 Substitution / hydrolysis ✓ (1)

1.2 Hydrogenation ✓ (1)



1.4 Esterification ✓ (1)

1.4.2 Concentrated H₂SO₄ / sulphuric acid ✓ (1)



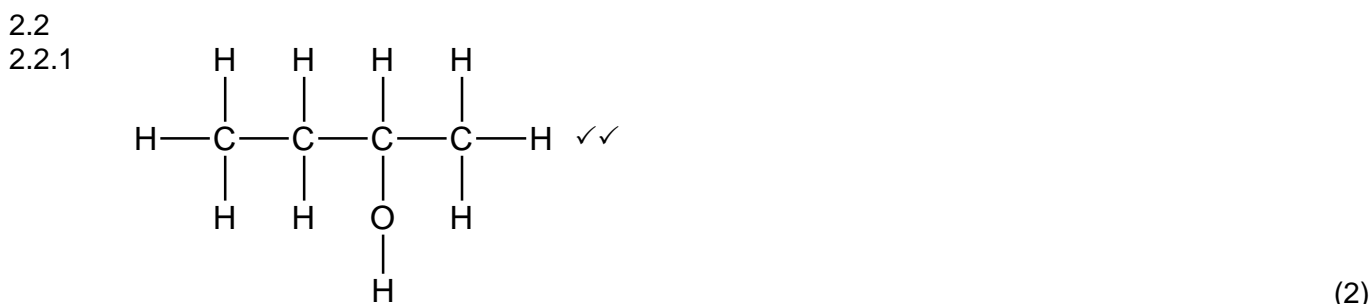
1.4.4 Propyl ✓ ethanoate ✓ (2)

1.5 Sulphuric acid / H₂SO₄ ✓ (1)

[15]

QUESTION 2

2.2 Unsaturated ✓
Contains a double bond / multiple bond between C atoms. ✓ (2)

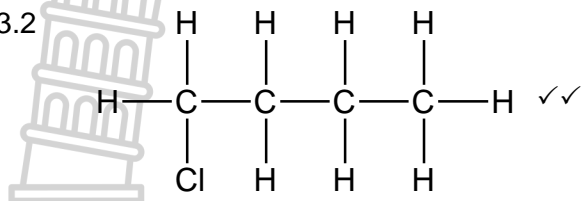


2.2.2 Addition / hydration ✓ (1)

2.3

2.3.1 2-chloro✓butane ✓ (2)

2.3.2 (1)



2.4

2.4.1 H₂O **OR** dilute NaOH / KOH ✓ (1)
Mild heat ✓

2.4.2 Substitution / hydrolysis ✓ (1)

2.4.3 C₄H₉Cl + NaOH ✓ → C₄H₁₀O + NaCl ✓ bal. ✓
OR
C₄H₉Cl + H₂O ✓ → C₄H₁₀O + HCl ✓ bal. ✓ (3)

[15]

QUESTION 3

3.1

3.1.1 Hydrolysis ✓ (1)

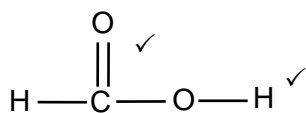
3.1.2 (Mild) heat ✓ (2)
Dilute (strong) base/aqueous base ✓

3.1.3 Ethanol ✓ (1)

3.2

3.2.1 Esterification ✓ (1)

3.2.2



Marking criteria:

- Functional group correct: ✓
- Whole structure correct: ✓✓

(2)

3.2.3 Ethyl✓ methanoate ✓ (2)

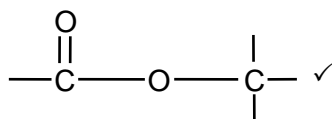
[9]

QUESTION 4

4.1

4.1.1 Esterification / Condensation ✓ (1)

4.1.2

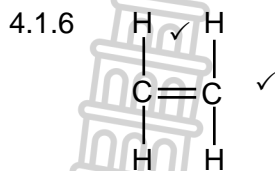


(1)

4.1.3 Propanoic acid ✓ (1)

4.1.4 Dehydration / elimination ✓ (1)

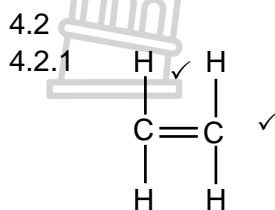
4.1.5 (Concentrated) H_2SO_4 / sulphuric acid / H_3PO_4 / phosphoric acid ✓ (1)



Marking criteria:

- Only functional group correct: ✓
- Whole structure correct: ✓✓

(2)



Marking criteria:

- Functional group correct: ✓
- Whole structure correct: ✓✓

(2)

4.2.2 Addition ✓ (1) [10]

QUESTION 5

5.1
5.1.1 Addition ✓ (1)

5.1.2 Polyethene ✓ (1)

5.2
5.2.1 Chloro✓ethane ✓ (2)

5.2.2 Hydrohalogenation/hydrochlorination ✓ (1)



(2)

5.3.2 HCl / hydrogen chloride ✓ (1)

5.4
5.4.1 Saturated ✓
There are no double / multiple bonds between C atoms. ✓ (2)

5.4.2 H_2 / hydrogen gas ✓ (1)

5.4.3 $2C_2H_6 + 7O_2 \rightarrow 4CO_2 + 6H_2O$ ✓ Bal. ✓ (3) [14]



QUESTION 6

6.1
6.1.1 Addition / Hydrogenation ✓ (1)

6.1.2 Elimination / Dehydrohalogenation / Dehydrobromination ✓ (1)

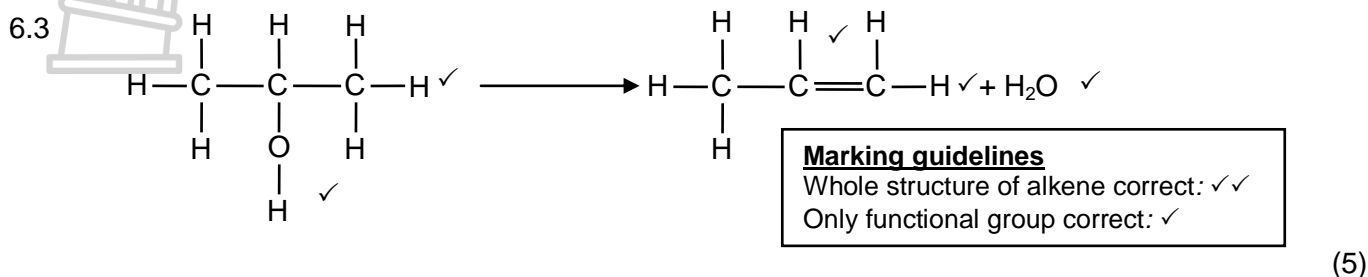
6.1.3 Substitution / Halogenation / Bromination ✓ (1)

6.2
6.2.1 Pt / platinum ✓ (1)

6.2.2 H₂SO₄ / sulphuric acid ✓ (1)

6.2.3 Hydration ✓ (1)

6.2.4 2-bromopropane ✓ (2)



6.4 Higher temperature ✓
Concentrated base ✓ (2)
[15]

QUESTION 7

7.1
7.1.1 High temperature / heat / high energy / high pressure ✓ (1)

7.1.2 C₆H₁₂ ✓ (1)

7.1.3 Alkenes ✓ (1)

7.2 **X** / C₆H₁₂ / Alkene ✓

OPTION 1

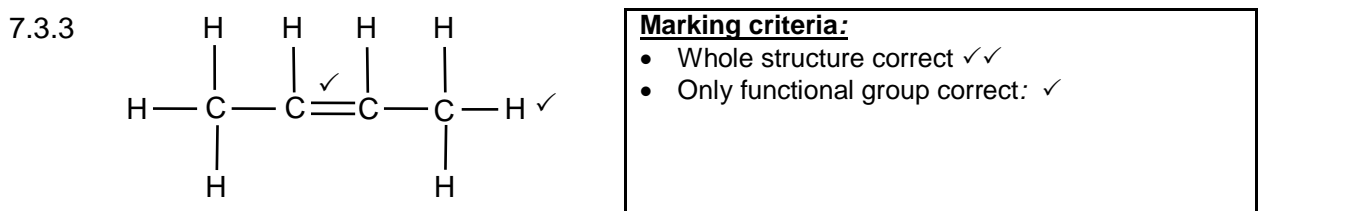
- **X** has a double bond. / **X** is unsaturated. ✓
- **X** can undergo addition. ✓
- **X** will react without light / heat. ✓

OPTION 2

- Butane is an alkane **OR** butane is saturated. ✓
- Butane can only undergo substitution. ✓
- Butane will only react in the presence of light or heat. ✓ (4)

7.3
7.3.1 2-chloro ✓ butane ✓ (2)

7.3.2 Substitution / Hydrolysis ✓ (1)



7.3.4 Hydration ✓ (1)
[13]

REACTION RATE AND ENERGY IN CHEMICAL REACTIONS

QUESTION 1

1.1 Change in concentration of products / reactants ✓ per unit time. ✓ (2)

1.2
1.2.1 Temperature ✓ (1)

1.2.2 Rate of reaction / Volume of gas formed per unit time ✓ (1)

1.3

- Larger mass / amount / surface area. ✓
- More effective collisions per unit time. / More particles collide with sufficient kinetic energy & correct orientation per unit time. ✓✓ (3)

1.4 Marking criteria

Compare Exp. 1 with Exp. 2:	The reaction in <u>exp. 1 is faster</u> than in <u>exp. 2</u> due to the <u>higher acid concentration</u> .	✓
	Therefore the <u>gradient</u> of the graph representing <u>exp. 1</u> is <u>greater / steeper</u> than that of <u>exp. 2</u> .	✓
Compare Exp. 1 with Exp 3 & 4:	The reaction in <u>exp. 3 is faster</u> than that in <u>exp. 1</u> due to the <u>higher temperature</u> .	✓
	The reaction in <u>exp. 4 is faster</u> than that in <u>exp. 1</u> due to the <u>higher temperature / larger surface area</u> . OR Graph <u>A</u> represents <u>exp. 4</u> due to the <u>greater mass</u> of CaCO_3 - <u>greater yield of CO_2</u> at a faster rate.	✓
	Therefore the <u>gradient</u> of the graphs of <u>exp. 3 & 4</u> are <u>greater/steeper</u> than that of <u>exp. 1</u> .	✓
Final answer	C	✓

(6)

1.5

<p>Marking criteria:</p> <ul style="list-style-type: none"> • Divide volume by: 25,7 ✓ • Use ratio: $n(\text{CO}_2) = n(\text{CaCO}_3) = 1:1$ ✓ • Substitute 100 in $n = \frac{m}{M}$. ✓ • Subtraction. ✓ • Final answer: 7,00 g to 7,5 g ✓
$n(\text{CO}_2) = \frac{V}{V_m}$ $= \frac{4,5}{25,7} \checkmark$ $= 0,18 \text{ mol}$ $n(\text{CaCO}_3) = n(\text{CO}_2) = 0,18 \text{ mol} \checkmark$ $n(\text{CaCO}_3) = \frac{m}{M}$ $0,18 = \frac{m}{100} \therefore m(\text{CaCO}_3) = 18 \text{ g}$ $m(\text{CaCO}_3) \text{ not reacted: } 25 - 18 \checkmark = 7,00 \text{ g} \checkmark$

(5)
[18]

QUESTION 2

2.1 Exothermic ✓
Temperature increases during reaction. / $T_i < T_f$ ✓ (2)

2.2 Larger surface area in experiment 2. ✓ (1)

2.3 More than one independent variable. ✓
OR Different concentrations and state of division. (1)

2.4 Faster than ✓
 A catalyst is used in experiment 5. ✓
 • A catalyst provides an alternative pathway of lower activation energy. ✓
 • More molecules have sufficient kinetic energy. ✓
 • More effective collisions per unit time. ✓ (5)

2.5 $n(\text{Zn}) = \frac{m}{M} \checkmark = \frac{1,2}{65} \checkmark = 0,018 \text{ mol}$
 $n(\text{HCl})_{\text{reacted}} = 2n(\text{Zn}) = 2(0,018) = 0,037 \text{ mol} \checkmark$
 $\text{Rate} = \frac{\Delta n}{\Delta t} = \frac{0,037}{8} \checkmark = 4,63 \times 10^{-3} \text{ mol} \cdot \text{s}^{-1} \checkmark$ (6)
[15]

QUESTION 3

3.1 Exothermic ✓
 $\Delta H < 0$ / Energy is released. ✓ (2)

	OPTION 1	OPTION 2	
3.2	$n(\text{HCl}) = cV$	Ave rate = $-\frac{\Delta c}{\Delta t}$	
3.2.1	$= (1,5) \checkmark (30 \times 10^{-3}) \checkmark$ $= 0,045 \text{ mol}$ Ave rate = $-\frac{\Delta n}{\Delta t}$ $= -\frac{(0 - 45) \checkmark}{(60 - 0) \checkmark}$ $= 7,5 \times 10^{-4} \text{ mol} \cdot \text{s}^{-1} \checkmark$	$= -\frac{(0 - 1,5) \checkmark}{(60 - 0) \checkmark}$ $= 0,025 \text{ mol} \cdot \text{dm}^{-3} \cdot \text{s}^{-1}$ $\therefore \text{average rate} = (0,025)(30 \times 10^{-3}) \checkmark \checkmark$ $= 7,5 \times 10^{-4} (\text{mol} \cdot \text{s}^{-1}) \checkmark$	(5)

3.2.2(a) Increases ✓
 • The reaction is exothermic, resulting in an increase in temperature. ✓
 • More molecules have sufficient kinetic energy. ✓
 • More effective collisions per unit time. ✓ (4)

3.2.2(b) Decreases ✓
 Concentration of acid decreases. ✓
OR The surface area of magnesium decreases. (2)

3.3 **ANY TWO**
 • Higher temperature ✓
 • Larger surface area/state of division/contact area of Mg. / Use Mg powder. ✓
 • Addition of a catalyst. (2)
[15]

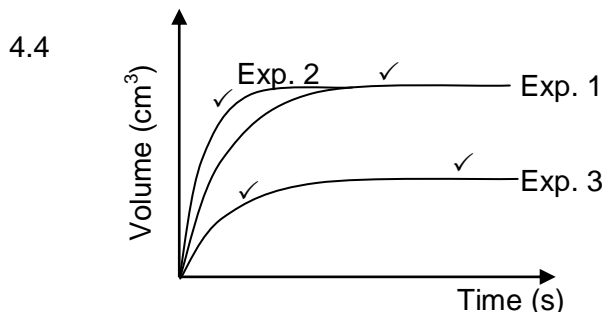
QUESTION 4

4.1 Time: Stop watch ✓
Volume: Gas syringe / burette ✓ (2)

4.2
4.2.1 t_1 ✓ (1)

4.2.2 t_3 ✓ (1)

4.3 Between t_1 and t_2 ✓ (1)



Marking criteria		
Exp. 2	Initial gradient higher than that of Exp. 1.	✓
2	Curve reaches same constant volume as for Exp. 1 (but earlier).	✓
Exp. 3	Initial gradient lower than that of Exp. 1.	✓
3	Curve reaches a smaller constant volume as for Exp. 1 (later stage).	✓

(4)

4.5.1

Marking criteria	
• $n(\text{HCl}) = (0,1)(100 \times 10^{-3})$	
• Use mole ratio: $n(\text{Zn}) = \frac{1}{2}n(\text{HCl})$	
• Substitute 65 into $n = \frac{m}{M}$.	
• $n(\text{Zn}_{\text{final}}) = n(\text{Zn}_{\text{initial}}) - n(\text{Zn}_{\text{used}})$ $m(\text{Zn}_{\text{final}}) = m(\text{Zn}_{\text{initial}}) - m(\text{Zn}_{\text{used}})$	
• Final answer: 0,33 g – 0,48 g	
OPTION 1 $n(\text{HCl}) = cV$ $= (0,1)(100 \times 10^{-3})$ ✓ $= 0,01 \text{ mol}$ n(Zn reacted): $n(\text{Zn}) = \frac{1}{2}n(\text{HCl})$ $= \frac{1}{2}(0,01)$ ✓ $= 5 \times 10^{-3} \text{ mol}$ $m(\text{Zn reacted}) = (5 \times 10^{-3})(65)$ ✓ $= 0,325 \text{ g}$ $m(\text{Zn}_f) = 0,8 - 0,325$ ✓ $= 0,48 \text{ g}$ ✓ (0,475 g)	OPTION 2 $n(\text{HCl}) = cV$ $= (0,1)(100 \times 10^{-3})$ ✓ $= 0,01 \text{ mol}$ $n(\text{Zn reacted}) = \frac{1}{2}n(\text{HCl})$ $= \frac{1}{2}(0,01)$ ✓ $= 5 \times 10^{-3} \text{ mol}$ $n(\text{Zn})_i = \frac{m}{M}$ $= \frac{0,8}{65}$ ✓ $= 1,23 \times 10^{-2} \text{ mol}$ $n(\text{Zn})_f = 1,23 \times 10^{-2} - 5 \times 10^{-3}$ ✓ $= 7,3 \times 10^{-3} \text{ mol}$ $m(\text{Zn}) = nM = (7,3 \times 10^{-3})(65) = 0,47 \text{ g}$ ✓

(5)

4.5.2 Smaller than ✓

(1)

[15]

QUESTION 5

- 5.1 **ANY TWO:**
 Increase temperature. ✓
 Increase concentration of acid. ✓
 Add a catalyst. ✓ (2)

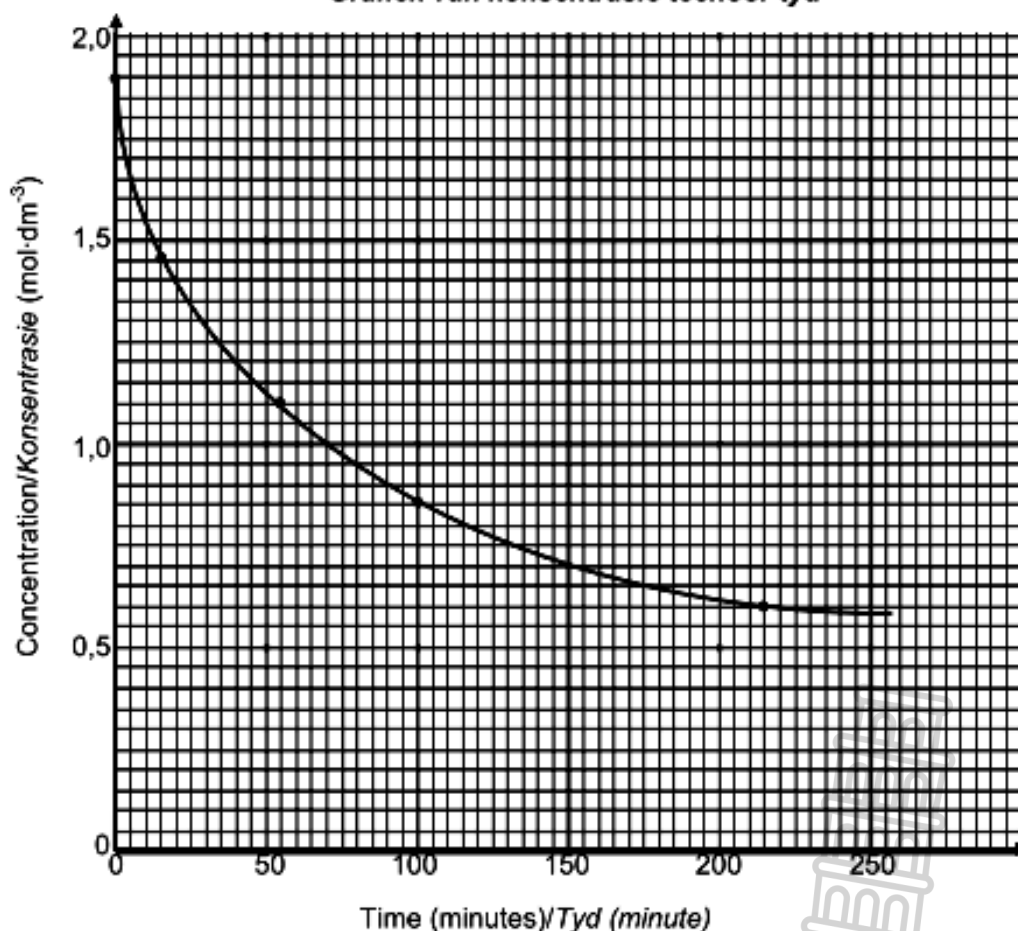
5.2 Change in concentration of products / reactants ✓ per unit time. ✓ (2)

5.3

5.3.1 average rate / *gemiddelde tempo* = $-\frac{\Delta c}{\Delta t}$
 $= -\frac{(1,45 - 1,90)}{(15 - 0)}$ ✓
 $= 0,03 \text{ (mol} \cdot \text{dm}^{-3}) \cdot \text{min}^{-1}$ ✓ (3)

5.3.2

Graph of concentration versus time
Grafiek van konsentrasie teenoor tyd



Marking criteria/Nasienriglyne	
Four points correctly plotted./Vier punte korrek gestip.	✓✓
Curve drawn as shown./Kurwe getrek soos getoon.	✓

5.3.3 1,15 to 1,25 mol·dm⁻³ ✓ (1)

- 5.3.4
- Concentration of reactants decreases. ✓
 - Less particles per unit volume. ✓
 - Less effective collisions per unit time. ✓ (3)

5.3.5

Marking criteria/Nasienriglyne	
<ul style="list-style-type: none"> Use $n = cV$ to calculate $\Delta n/n(\text{initial})$ & $n(\text{final})$. <i>Gebruik $n = cV$ om $\Delta n/n(\text{aanvanklik})$ & $n(\text{finaal})$ te bereken.</i> $\Delta n(\text{HCl}) = n(\text{final/finaal}) - n(\text{initial/aanvanklik})$. OR/OF $\Delta c(\text{HCl}) = c(\text{final/finaal}) - c(\text{initial/aanvanklik})$ Use ratio/<i>Gebruik verhouding</i> $n(\text{CH}_3\text{Cl}) : n(\text{HCl}) = 1 : 1$ Substitute/<i>Vervang</i> $50,5 \text{ g}\cdot\text{mol}^{-1}$ in $n = \frac{m}{M}$. Final answer/<i>Finale antwoord</i>: 3,54–4,0 g. 	
<p>OPTION 1/OPSIE 1</p> <p>Mol initially/<i>begin</i>: $n(\text{HCl}) = cV$ ✓ $= (1,9)(60 \times 10^{-3})$ ✓ $= 0,11 \text{ mol (0,114)}$</p> <p>Mol final/<i>finaal</i>: $n(\text{HCl}) = cV$ $= (0,6)(60 \times 10^{-3})$ $= 0,04 \text{ mol (0,036)}$</p> <p>$\Delta n(\text{HCl}) = 0,04 - 0,11$ ✓ $= -0,07 \text{ mol (0,078 mol)}$ $\Delta n(\text{HCl}) = 0,07 \text{ mol (0,078)}$</p> <p>$n(\text{formed/gevorm}) = n(\text{reacted/reageer})$ $n(\text{CH}_3\text{Cl}) = n(\text{HCl})$ ✓ $= 0,07 \text{ mol}$</p> <p>$m(\text{CH}_3\text{Cl}) = nM$ $= (0,07)(50,5)$ ✓ $= 3,54 \text{ g}$ ✓</p> <p>Accept range/<i>Aanvaar gebied</i>: 3,54 – 4,0 g</p>	<p>OPTION 2/OPSIE 2</p> <p>$\Delta c(\text{HCl}) = 0,6 - 1,9$ ✓ $= -1,3$ $= 1,3 \text{ mol}\cdot\text{dm}^{-3}$</p> <p>$\Delta n(\text{HCl}) = \Delta cV$ $= (1,3)(60 \times 10^{-3})$ ✓ $= 0,08 \text{ mol (0,078)}$</p> <p>$n(\text{formed/gevorm}) = n(\text{reacted/reageer})$ $n(\text{CH}_3\text{Cl}) = n(\text{HCl})$ ✓ $= 0,08 \text{ mol}$</p> <p>$m(\text{CH}_3\text{Cl}) = nM$ $= (0,08)(50,5)$ ✓ $= 4 \text{ g}$ ✓</p> <p>Accept range/<i>Aanvaar gebied</i>: 3,54 – 4,0 g</p>

(5)
[19]

QUESTION 6

6.1 **ANY TWO:**

- Temperature of reaction mixture ✓
- Addition of a catalyst ✓
- Concentration of reactants

(2)

6.2 Sulphur / S ✓

(1)

6.3 Water is used to dilute/change the concentration of the $\text{Na}_2\text{S}_2\text{O}_3(\text{aq})$ ✓

(1)

6.4	Criteria for investigative question:	
	The <u>dependent</u> and <u>independent</u> variables are stated correctly.	✓
	Asks a question about the relationship between <u>dependent</u> and <u>independent</u> variables.	✓

Dependent variable: reaction rate
Independent variable: concentration

Examples:

- What is the relationship between concentration and reaction rate?
- How does the reaction rate change with change in concentration? (2)

6.5 A ✓ (1)

6.6 **Experiment B:**

- The concentration of Na₂S₂O₃(aq) is higher. / More Na₂S₂O₃ particles per unit volume. ✓
- More particles with correct orientation. ✓
- More effective collisions per unit time. ✓ (3)

6.7

<p>OPTION 1</p> $c = \frac{m}{MV} \checkmark$ $= \frac{62,5}{\checkmark(158)(0,25)} \checkmark$ $= 1,58 \text{ mol} \cdot \text{dm}^{-3}$	<p>OPTION 2</p> $n = \frac{m}{M} \checkmark$ $= \frac{62,5}{158} \checkmark$ $= 0,396 \text{ mol}$ $c = \frac{n}{V}$ $= \frac{0,396}{0,25} \checkmark$ $= 1,58 \text{ mol} \cdot \text{dm}^{-3}$	<p>OPTION 3</p> <p>250 cm³ 62,5 g</p> <p>10 cm³ $\frac{10}{250} \checkmark \times 62,5 \checkmark$</p> <p>∴ m(Na₂S₂O₃ in D)</p> <p>= 2,5 g</p> $n(\text{Na}_2\text{S}_2\text{O}_3 \text{ in D}) = \frac{m}{M}$ $= \frac{2,5}{158} \checkmark \} \checkmark$ $= 0,0158 \text{ mol}$
$n(\text{Na}_2\text{S}_2\text{O}_3 \text{ in D}) = cV$ $= (1,58)(0,01) \checkmark$ $= 0,0158 \text{ mol}$		$n(\text{S}) = (\text{Na}_2\text{S}_2\text{O}_3) \checkmark$ $= 0,0158$ $n = \frac{m}{M}$ $0,0158 = \frac{m}{32} \checkmark$ $m = 0,51 \text{ g} \checkmark$

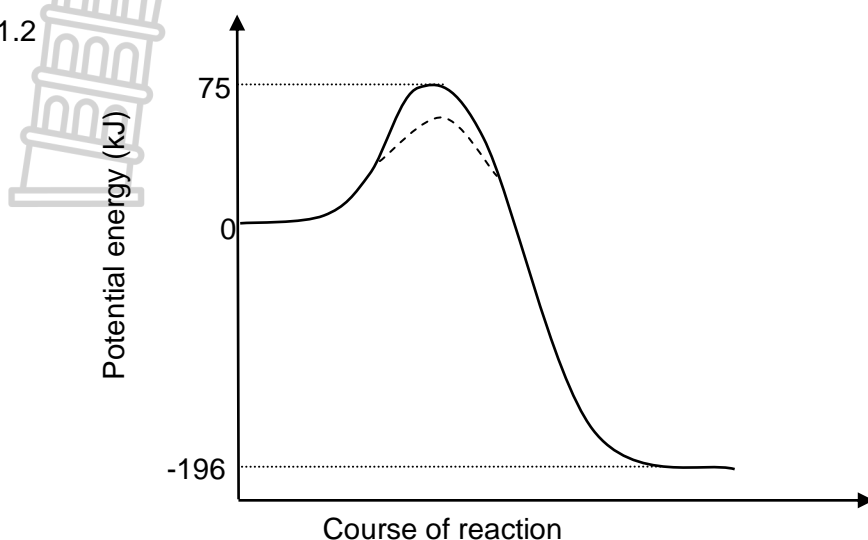
(7)
[17]

QUESTION 7

7.1

7.1.1 The minimum energy needed for a reaction to take place. ✓✓ (2)

7.1.2



Marking criteria:	
Shape of curve for exothermic reaction as shown.	✓
Energy of activated complex shown as 75 kJ in line with the peak.	✓
Energy of products shown as - 196 kJ below the zero.	✓
IF: Wrong shape, e.g. straight line.	0/3

(3)

7.1.3

Marking criteria

- Dotted line (---) on graph in QUESTION 7.1.2 showing lower energy for activated complex. ✓
- Dotted curve starts at/above energy of reactants and ends at/above energy of products on the inside of the original curve. ✓

(2)

7.1.4

- A catalyst provides an alternative pathway of lower activation energy. ✓
- More molecules have sufficient kinetic energy. ✓
- More effective collisions per unit time. ✓

(3)

7.2

7.2.1

$$\begin{aligned}
 \text{Ave rate} &= \frac{\Delta V}{\Delta t} \\
 &= \frac{52 - 16}{40 - 10} \\
 &= 1,2 \text{ (dm}^3 \cdot \text{s}^{-1})
 \end{aligned}$$

(3)



7.2.2

<p>Marking criteria:</p> <ul style="list-style-type: none"> • $V(O_2) = 60 \text{ dm}^3$ AND divide volume by 24. ✓ • Use ratio: $n(H_2O_2) = 2n(O_2) = 1:2$ ✓ • Use $34 \text{ g} \cdot \text{mol}^{-1}$ in $n = \frac{m}{M}$ or in ratio calculation. ✓ • Final answer: 170 g ✓ 		
<p>OPTION 1</p> $n(O_2) = \frac{V}{V_M}$ $= \frac{60}{24} \checkmark$ $= 2,5 \text{ mol}$ $n(H_2O_2) = 2n(O_2)$ $= 2(2,5) \checkmark$ $= 5 \text{ mol}$ $n(H_2O_2) = \frac{m}{M}$ $\therefore 5 = \frac{m}{34} \checkmark$ $\therefore m = 170 \text{ g} \checkmark$	<p>OPTION 2</p> $24 \text{ dm}^3 : 1 \text{ mol}$ $60 \text{ dm}^3 : 2,5 \text{ mol} \checkmark$ $n(H_2O_2) = 2n(O_2)$ $= 2(2,5) \checkmark$ $= 5 \text{ mol}$ $34 \text{ g} \checkmark : 1 \text{ mol}$ $x : 5 \text{ mol}$ $x = 170 \text{ g} \checkmark$	<p>OPTION 3</p> $n(O_2) = \frac{V}{V_M}$ $= \frac{60}{24} \checkmark$ $= 2,5 \text{ mol}$ $n(O_2) = \frac{m}{M}$ $\therefore 2,5 = \frac{m}{32}$ $\therefore m = 80 \text{ g}$ \checkmark $2(34) \text{ g} \checkmark H_2O_2 \dots\dots\dots 32 \text{ g } O_2$ $x \text{ g } H_2O_2 \dots\dots\dots 80 \text{ g } O_2$ $m(H_2O_2) = 170 \text{ g} \checkmark$

(4)

7.2.3 Equal to ✓

(1)

7.3

7.3.1 Q ✓

(1)

7.3.2 P ✓

(1)

[20]



CHEMICAL EQUILIBRIUM

QUESTION 1

1.1 The stage in a chemical reaction when the rate of forward reaction equals the rate of reverse reaction. ✓✓

OR

The stage in a chemical reaction when the concentrations of reactants and products remain constant. (2)

1.2 **CALCULATIONS USING NUMBER OF MOLES**

- Correct K_c expression (formulae in square brackets). ✓
- Substitution of concentrations into K_c expression. ✓
- Substitution of K_c value. ✓
- Equilibrium concentration of both NO_2 & N_2O_4 multiplied by $0,08 \text{ dm}^3$. ✓
- Change in $n(\text{N}_2\text{O}_4) = \text{equilibrium } n(\text{N}_2\text{O}_4) - \text{initial } n(\text{N}_2\text{O}_4)$ ✓
- **USING** ratio: $\text{NO}_2 : \text{N}_2\text{O}_4 = 2 : 1$ ✓
- Initial $n(\text{NO}_2) = \text{equilibrium } n(\text{NO}_2) + \text{change } n(\text{NO}_2)$. ✓
- Final answer: $1,11 \text{ (mol)}$ ✓

OPTION 1

$$K_c = \frac{[\text{N}_2\text{O}_4]}{[\text{NO}_2]^2} \quad \checkmark$$

$$171 \checkmark = \frac{[\text{N}_2\text{O}_4]}{[\text{NO}_2]^2} \quad \checkmark$$

$$\therefore [\text{N}_2\text{O}_4] = 171 \times (0,2)^2 = 6,84 \text{ mol}\cdot\text{dm}^{-3}$$

	NO_2	N_2O_4	
Initial quantity (mol)	$1,11 \checkmark$	0	
Change (mol) subtract ✓	$1,094$	$0,55 \checkmark$	ratio ✓
Quantity at equilibrium (mol)/	$0,016$	$0,55$	
Equilibrium concentration ($\text{mol}\cdot\text{dm}^{-3}$)	$0,2$	$6,84$	$\times 0,08 \checkmark$

OPTION 2

$$K_c = \frac{[\text{N}_2\text{O}_4]}{[\text{NO}_2]^2} \quad \checkmark$$

$$171 \checkmark = \frac{[\text{N}_2\text{O}_4]}{(0,2)^2} \quad \checkmark$$

$$\therefore [\text{N}_2\text{O}_4] = 171 \times (0,2)^2 = 6,84 \text{ mol}\cdot\text{dm}^{-3}$$

Equilibrium moles:

$$\left. \begin{aligned} n(\text{N}_2\text{O}_4) &= (6,84)(0,080) \\ &= 0,55 \text{ mol} \\ n(\text{NO}_2) &= (0,2)(0,080) \\ &= 0,016 \text{ mol} \end{aligned} \right\} \checkmark \times 0,08 \text{ dm}^3$$

$$n(\text{N}_2\text{O}_4 \text{ formed}) = 0,55 - 0 = 0,55 \text{ mol} \quad \checkmark$$

Ratio:

$$n(\text{NO}_2 \text{ reacted}) = 2n(\text{N}_2\text{O}_4 \text{ formed}) = 2(0,55) = 1,094 \text{ mol} \quad \checkmark$$

$$\text{Initial } n(\text{NO}_2) = 0,016 + 1,094 \checkmark = 1,11 \text{ (mol)} \quad \checkmark$$



(8)

1.3

1.3.1 Concentration (of the gases) increases. / Molecules become more condensed or move closer to each other. ✓

(1)

- 1.3.2
- Increase in pressure favours the reaction that leads to smaller number of moles of gas. ✓
 - Forward reaction is favoured. ✓
 - Number of moles/amount of N_2O_4 / colourless gas increases. ✓
 - **OR** Number of moles/amount of NO_2 / brown gas decreases. ✓
- (3)
- 1.4
- 1.4.1 Darker ✓ (1)
- 1.4.2 Decreases ✓ (1)
- [16]**

QUESTION 2

- 2.1 A reaction is reversible when products can be converted back to reactants. ✓ (1)
- 2.2 No change ✓ (1)
- 2.3
- 2.3.1 Temperature decreases ✓ (1)
- 2.3.2
- Decrease in temperature decreases the rate of both forward and reverse reactions. ✓
 - Decrease in temperature favours the exothermic reaction. ✓
 - The rate of the reverse (exothermic) reaction is faster or the reverse reaction is favoured. ✓
- (3)

2.4

OPTION 1/OPSIE 1

At equilibrium/by ewewig: $[H_2] = [I_2]$ ✓

$$K_c = \frac{[H_2][I_2]}{[HI]^2} \checkmark$$

$$\therefore 0,02 \checkmark = \frac{(x)(x)}{\left(\frac{0,04}{2}\right)^2} \checkmark \quad \begin{array}{l} \text{Divide by } 2 \text{ dm}^3 \checkmark \\ \text{Deel deur } 2 \text{ dm}^3 \end{array}$$

$$\therefore x = [H_2] = 2,83 \times 10^{-3} \text{ mol}\cdot\text{dm}^{-3} \checkmark \quad (0,0028 \text{ mol}\cdot\text{dm}^{-3})$$

OPTION 2/OPSIE 2

	HI	H ₂	I ₂	
Initial quantity (mol) <i>Aanvangshoeveelheid (mol)</i>	x	0	0	
Change (mol) <i>Verandering (mol)</i>	x - 0,04	$\frac{x - 0,04}{2}$	$\frac{x - 0,04}{2}$	ratio ✓ <i>verhouding</i>
Quantity at equilibrium (mol)/ <i>Hoeveelheid by ewewig (mol)</i>	0,04	$\frac{x - 0,04}{2}$	$\frac{x - 0,04}{2}$	
Equilibrium concentration (mol·dm ⁻³) <i>Ewewigskonsentrasie (mol·dm⁻³)</i>	0,02	$\frac{x - 0,04}{4}$	$\frac{x - 0,04}{4}$	Divide by 2 dm ³ ✓ <i>Deel deur 2 dm³</i>

$K_c = \frac{[H_2][I_2]}{[HI]^2} \checkmark$

$$\therefore 0,02 \checkmark = \frac{\left(\frac{x - 0,04}{4}\right)\left(\frac{x - 0,04}{4}\right)}{(0,02)^2} \checkmark$$

$$\therefore x = 0,05$$

$$[H_2] = \frac{x - 0,04}{2}$$

$$= \frac{0,05 - 0,04}{2}$$

$$= 2,83 \times 10^{-3} \text{ mol}\cdot\text{dm}^{-3} \checkmark$$

No K_c expression, correct substitution/Geen K_c -uitdrukking, korrekte substitusie: Max./Maks. $\frac{5}{6}$

Wrong K_c expression/Verkeerde K_c -uitdrukking: Max./Maks. $\frac{2}{6}$

(6)

2.5 $K_c = \frac{1}{0,02} = 50 \checkmark$ (1)

2.6 Increases \checkmark (1)
[14]

QUESTION 3

3.1 When the equilibrium (in a closed system) is disturbed, the system will re-instate a new equilibrium \checkmark by favouring the reaction that will cancel the disturbance. \checkmark (2)

3.2
3.2.1 Remains the same \checkmark (1)

3.2.2 Increases \checkmark (1)

3.3

<p>OPTION 1 $K_c = [\text{NH}_3][\text{H}_2\text{S}] \checkmark$ $\therefore 1,2 \times 10^{-4} \checkmark = [\text{NH}_3][\text{H}_2\text{S}]$ $\therefore [\text{NH}_3] = [\text{H}_2\text{S}]$ $= 0,011 \text{ mol}\cdot\text{dm}^{-3}$ $n(\text{NH}_3) = cV$ $= (0,011)(5) \checkmark$ $= 0,06 \text{ mol (0,06 mol)}$ $n(\text{NH}_4\text{HS}) = n(\text{NH}_3) = 0,06 \text{ mol} \checkmark$ $m(\text{NH}_4\text{HS}) = nM$ $= (0,06)(51) \checkmark$ $= 2,81 \text{ g} \checkmark$</p>	<p>Marking criteria:</p> <ul style="list-style-type: none"> • K_c expression. • Substitution of $1,2 \times 10^{-4}$. • x volume (5 dm^3) • <u>Use</u> mole ratio: 1:1 • Substitution $51 \text{ g}\cdot\text{mol}^{-1}$ • Final answer: 2,81 g
---	--

OPTION 2

	NH_4HS	NH_3	H_2S
Initial quantity (mol)		0	0
Change (mol)	x	x	x
Quantity at equilibrium (mol)	-	x	x
Equilibrium concentration ($\text{mol}\cdot\text{dm}^{-3}$)	-	$\frac{x}{5}$	$\frac{x}{5}$

Divide by 5 \checkmark

$K_c = [\text{NH}_3][\text{H}_2\text{S}] \checkmark$
 $\therefore 1,2 \times 10^{-4} \checkmark = \left(\frac{x}{5}\right)\left(\frac{x}{5}\right)$
 $\therefore x = 0,0547 \text{ mol}$
 $m(\text{NH}_4\text{HS}) = nM$
 $= (0,0547) \checkmark (51) \checkmark$
 $= 2,79 \text{ g} \checkmark$

3.4 Decreases \checkmark (6)

- Increase in pressure favours the reaction that leads to the smaller number of moles of gas. $\checkmark\checkmark$
- The reverse reaction is favoured. \checkmark

(3)
[13]

QUESTION 4

4.1 Equal to ✓ (1)

4.2
$$K_c = \frac{[X_3]^2}{[X_2]^3} \checkmark$$

$$= \frac{(0,226)^2}{(0,06)^3} \checkmark$$

$$= 236,46 \checkmark$$
 (4)

4.3 (1)

4.3.1 Increases ✓

- 4.3.2
- The increase in $[X_3]$ is opposed. / Change is opposed. ✓
 - The reverse reaction is favoured. / X_3 is used. / $[X_3]$ decreases. ✓
- (2)

4.4 Higher than ✓ (1)

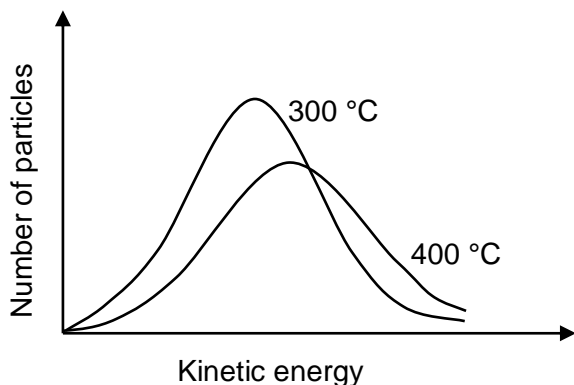
4.5 Exothermic ✓

- The concentration of the product / $X_3(g)$ is lower. / The concentration of the reactant / $X_2(g)$ is higher. ✓
- The increase in temperature favoured the reverse reaction. ✓
- According to Le Chatelier's principle an increase in temperature favours the endothermic reaction. ✓

OR

- $[X_3]$ decreases and $[X_2]$ increases. ✓
 - K_c decreases if temperature increases. ✓
 - Decrease in temperature favoured the forward reaction. ✓
- (4)

4.6



Marking criteria		
Peak of curve at 400 °C lower than at 300 °C and shifted to the right.		✓
Curve at 400 °C has larger area at the higher E_k .		✓

(2)
[15]

QUESTION 5

5.1

OPTION 1/OPSIE 1	OPTION 2/OPSIE 2
$c = \frac{m}{MV} \checkmark$ $= \frac{2,2}{(44)(5)} \checkmark$ $= 0,01 \text{ mol} \cdot \text{dm}^{-3} \checkmark$	$n = \frac{m}{M}$ $= \frac{2,2}{44} \checkmark$ $= 0,05 \text{ mol}$ $c = \frac{n}{V}$ $= \frac{0,05}{5} \checkmark$ $= 0,01 \text{ mol} \cdot \text{dm}^{-3} \checkmark$
Both formulae/ albei formules	

(4)

5.2 For equilibrium, a forward and a reverse reaction are needed. ✓

OR Without CaO(s), the reverse reaction is not possible.

OR If only CO₂ is present, the reverse reaction cannot take place.

(1)

5.3 CO₂ is a gas and will escape if the container is not sealed. ✓

(1)

5.4

OPTION 1 (POSITIVE MARKING FROM Q6.1)	
$K_c = [\text{CO}_2] \checkmark$ $= 0,0108$ $\therefore [\text{CO}_2] = 0,0108 \text{ (mol} \cdot \text{dm}^{-3}) \checkmark$	\swarrow
$n(\text{CO}_2 \text{ at equilibrium/by ewewig}) = cV$ $= (0,0108)(5) \checkmark$ $= 0,054 \text{ mol}$	\swarrow
$n(\text{CO}_2 \text{ formed/gevorm}) = n(\text{CO}_2 \text{ at equilibrium/by ewewig})$ $n(\text{CO}_2 \text{ initially/begin})$ $= 0,054 - 0,05 \checkmark$ $= 0,004 \text{ mol}$	\swarrow
$n(\text{CaCO}_3) = n(\text{CO}_2 \text{ formed}) = 0,004 \text{ mol} \checkmark$	\swarrow
$m(\text{CaCO}_3) = nM$ $= (0,004)(100) \checkmark$ $= 0,4 \text{ g} \checkmark$	\swarrow



OPTION 2 (POSITIVE MARKING FROM Q6.1)

$K_c = [\text{CO}_2] \checkmark$
 $= 0,0108 \checkmark$
 $\therefore [\text{CO}_2] = 0,0108 \text{ (mol}\cdot\text{dm}^{-3}\text{)}$

	CaCO ₃	CaO	CO ₂	
Initial quantity (mol) <i>Aanvangshoeveelheid (mol)</i>	0	0	0,05	
Change (mol) <i>Verandering (mol)</i>	0,004	x	0,004 \checkmark	\checkmark Ratio/ Verhouding
Quantity at equilibrium (mol) <i>Hoeveelheid by ewewig (mol)</i>			0,054 \checkmark	
Equilibrium concentration (mol·dm ⁻³) <i>Ewewigskonsentrasie (mol·dm⁻³)</i>			0,0108	

$m(\text{CaCO}) = nM$
 $= (0,004)(100) \checkmark$
 $= 0,4 \text{ g} \checkmark$

(7)

5.5

5.5.1 Remains the same \checkmark

(1)

5.5.2 Decreases \checkmark

(1)

5.6 Endothermic \checkmark

- K_c decreases with decrease in temperature. \checkmark
- Therefore the product of the concentration of products decreases./The reverse reaction is favoured. \checkmark
- A decrease in temperature favours the exothermic reaction. \checkmark

OR

- K_c increases with increase in temperature. \checkmark
- Increase in temperature favours the forward reaction. \checkmark
- Increase in temperature favours the endothermic reaction. \checkmark

(4)

[19]

QUESTION 6

6.1 Reversible reaction ✓ (1)

6.2 Endothermic ✓
 ΔH is positive. / $\Delta H > 0$ ✓ (2)

6.3 Larger than ✓
 $K_c > 1$ ✓ (2)

6.4

OPTION 1

$$n = \frac{m}{M}$$

$$= \frac{168}{28} \checkmark$$

$$= 6 \text{ mol}$$

	CO ₂	CO	
Initial quantity (mol)	x	0	
Change (mol)	3	6 ✓	ratio ✓
Quantity at equilibrium (mol)	x - 3 ✓	6	
Equilibrium concentration (mol·dm ⁻³)	$\frac{x-3}{2}$	3	Divide by 2 ✓

$$K_c = \frac{[\text{CO}]^2}{[\text{CO}_2]} \checkmark$$

$$14 \checkmark = \frac{(3)^2}{\frac{x-3}{2}} \checkmark$$

$$\therefore x = 4,29 \text{ mol} \checkmark$$

OPTION 2

$$n = \frac{m}{M} \qquad c = \frac{n}{V}$$

$$= \frac{168}{28} \checkmark \qquad = \frac{6}{2} \text{ Divide by 2 } \checkmark$$

$$= 6 \text{ mol} \qquad = 3 \text{ mol} \cdot \text{dm}^{-3}$$

	CO ₂	CO	
Initial concentration (mol·dm ⁻³)	x	0	
Change (mol·dm ⁻³)	1,5	3 ✓	ratio ✓
Equilibrium concentration (mol·dm ⁻³)	x - 1,5 ✓	3	

$$K_c = \frac{[\text{CO}]^2}{[\text{CO}_2]} \checkmark$$

$$14 \checkmark = \frac{[3]^2}{x-1,5} \checkmark$$

$$\therefore x = 2,14 \text{ mol} \cdot \text{dm}^{-3}$$

$$n(\text{CO}_2) = cV = (2,14)(2) = 4,29 \text{ mol} \checkmark$$

OPTION 3

$$n = \frac{m}{M}$$

$$= \frac{168}{28} \checkmark$$

$$= 6 \text{ mol}$$

	CO ₂	CO
Initial quantity (mol)	4,28 ✓	0
Change (mol)	3	6
Quantity at equilibrium (mol)	1,28 ✓	6 ✓
Equilibrium concentration (mol·dm ⁻³)	0,64	3

ratio ✓

x 2 ✓

$$K_c = \frac{[\text{CO}]^2}{[\text{CO}_2]} \checkmark$$

$$14 \checkmark = \frac{[3]^2}{[\text{CO}_2]} \checkmark$$

$$\therefore [\text{CO}_2] = 0,64 \text{ mol}\cdot\text{dm}^{-3}$$

(9)

6.5

6.5.1 Remains the same ✓

(1)

6.5.2 Decreases ✓

(1)

6.5.3 Increases ✓

(1)

[17]**QUESTION 7**7.1 The stage in a chemical reaction when the rate of forward reaction equals the rate of reverse reaction. ✓✓**OR**The stage in a chemical reaction when the concentrations / quantities of reactants and products remain constant.

(2)

7.2

7.2.1 Remains the same ✓

(1)

7.2.2 Decreases ✓

- When the temperature is increased the reaction that will oppose this increase / decrease the temperature will be favoured. ✓

OR

The forward reaction is exothermic.

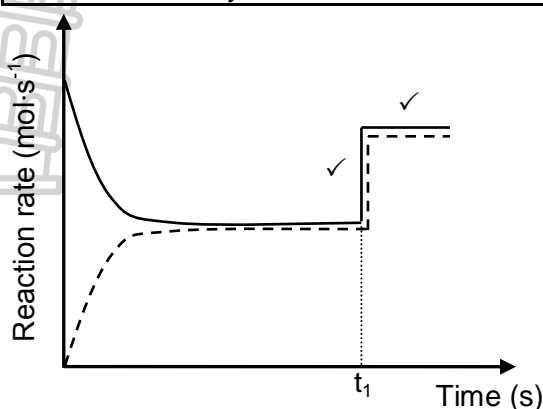
- An increase in temperature favours the endothermic reaction. ✓
- The reverse reaction is favoured. ✓

(4)

7.3

Marking criteria:

- Vertical parallel lines show a sudden increase in rate of both forward and reverse reactions. ✓
- Horizontal parallel lines showing a constant higher rate for both forward and reverse catalysed reactions after time t_1 . ✓



(2)

7.4

CALCULATIONS USING NUMBER OF MOLES

- Use $M(\text{PbS}) = 239 \text{ g}\cdot\text{mol}^{-1}$ in $n = \frac{m}{M}$ or in ratio calculation. ✓
- Use ratio: $n(\text{H}_2\text{S})_{\text{equil}} = n(\text{PbS})$ ✓
- $n(\text{H}_2\text{S})_{\text{formed}} = n(\text{H}_2\text{S})_{\text{equilibrium}}$ ✓
- **USING** ratio: $\text{H}_2 : \text{H}_2\text{S} = 1 : 1$ ✓
- $n(\text{H}_2)_{\text{equilibrium}} = n(\text{H}_2)_{\text{initial}} - n(\text{H}_2)_{\text{formed}}$ ✓
- Divide equilibrium $n(\text{H}_2\text{S})$ & $n(\text{H}_2)$ by 2 dm^3 . ✓
- Correct K_c expression ✓
- Substitution of concentrations into K_c expression. ✓
- Final answer: 0,07 ✓

OPTION 1

$$n(\text{PbS}) = \frac{m}{M}$$

$$= \frac{2,39}{239} \checkmark$$

$$= 0,01 \text{ mol}$$

$$n(\text{H}_2\text{S})_{\text{equilibrium}} = n(\text{PbS}) \checkmark = 0,01 \text{ mol}$$

	H_2	H_2S
Initial quantity (mol)	0,16	0
Change (mol)	0,01	0,01 ✓
Quantity at equilibrium (mol)	0,15 ✓	0,01
Equilibrium concentration ($\text{mol}\cdot\text{dm}^{-3}$)	0,075	0,005

ratio ✓

divide by 2 ✓

$$K_c = \frac{[\text{H}_2\text{S}]}{[\text{H}_2]} \checkmark$$

$$= \frac{0,005}{0,075} \checkmark$$

$$= 0,067 \approx 0,07 \checkmark$$

OPTION 2

$$n(\text{PbS}) = \frac{m}{M}$$

$$= \frac{2,39}{239} \checkmark$$

$$= 0,01 \text{ mol}$$

$$n(\text{H}_2\text{S})_{\text{reacted}} = n(\text{PbS}) \checkmark = 0,01 \text{ mol}$$
$$= n(\text{H}_2\text{S})_{\text{equilibrium}}$$

$$n(\text{H}_2\text{S})_{\text{formed}} = n(\text{H}_2\text{S})_{\text{equilibrium}} - n(\text{H}_2\text{S})_{\text{initial}}$$
$$= 0,01 - 0 \checkmark$$
$$= 0,01 \text{ mol}$$

$$n(\text{H}_2)_{\text{reacted}} = n(\text{H}_2\text{S})_{\text{formed}} \checkmark = 0,01 \text{ mol}$$

$$n(\text{H}_2)_{\text{equilibrium}} = n(\text{H}_2)_{\text{initial}} - n(\text{H}_2)_{\text{reacted}}$$
$$= 0,16 - 0,01 \checkmark$$
$$= 0,15 \text{ mol}$$

$$c(\text{H}_2) = \frac{n}{V}$$
$$= \frac{0,15}{2} \checkmark$$

$$= 0,075 \text{ mol} \cdot \text{dm}^{-3}$$

$$c(\text{H}_2\text{S}) = \frac{n}{V}$$
$$= \frac{0,01}{2}$$

$$= 0,005 \text{ mol} \cdot \text{dm}^{-3}$$

$$K_c = \frac{[\text{H}_2\text{S}]}{[\text{H}_2]} \checkmark$$

$$= \frac{0,005}{0,075} \checkmark$$

$$= 0,067 \approx 0,07 \checkmark$$

(9)
[18]



ACIDS AND BASES

QUESTION 1

- 1.1
- 1.1.1 Ionises / dissociates completely in water. ✓ (1)
- 1.1.2 NO_3^- / Nitrate ion ✓ (1)
- 1.1.3 $\text{pH} = -\log[\text{H}_3\text{O}^+] / -\log[\text{H}^+] \checkmark = -\log(0,3) \checkmark = 0,52 \checkmark$ (3)
- 1.2
- 1.2.1 $c = \frac{n}{V} \checkmark$
 $2 = \frac{n}{0,1} \checkmark$
 $\therefore n(\text{HCl}) = 0,2 \text{ mol} \checkmark$ (3)
- 1.2.2 Burette ✓ (1)
- 1.2.3 B ✓
 Titration of strong acid and strong base. ✓✓ (3)
- 1.2.4 The number of moles of acid in the flask remains constant. ✓ (1)
- 1.2.5 $c = \frac{n}{V} \checkmark \therefore 0,2 = \frac{n}{0,021} \checkmark \therefore n = 4,2 \times 10^{-3} \text{ mol} \checkmark$
 $n(\text{HCl})_{\text{excess}} = n(\text{NaOH}) = 4,2 \times 10^{-3} \text{ mol}$ (3)
- 1.2.6 **POSITIVE MARKING FROM QUESTION 1.2.1 AND 1.2.5.**

Marking criteria:

- $n(\text{HCl reacted}) = \text{initial (from Q1.2.1)} - \text{excess (from Q1.2.5)} \checkmark$
- Use mol ratio of acid: base = 2 : 1. ✓
- Substitute 40 into $n = \frac{m}{M} \checkmark$
- $\frac{m(\text{MgO})_{\text{reacted}}}{4,5} \times 100 \checkmark$
- Final answer: 87,11 % ✓

OPTION 1

$n(\text{HCl reacted}):$
 $0,2 - 4,2 \times 10^{-3} \checkmark = 0,196 \text{ mol}$

$n(\text{MgO reacted}):$
 $\frac{1}{2}n(\text{HCl}) = \frac{1}{2}(0,196)$
 $= 9,8 \times 10^{-2} \text{ mol} \checkmark$

$n(\text{MgO reacted}) = \frac{m}{M}$
 $\therefore 0,098 = \frac{m}{40} \checkmark$
 $\therefore m = 3,92 \text{ g}$
 $\% \text{ purity} = \frac{3,92}{4,5 \times 100} \checkmark$
 $= 87,11\% \checkmark$

OPTION 2

$n(\text{HCl reacted}):$
 $0,2 - 4,2 \times 10^{-3} \checkmark = 0,196 \text{ mol}$

$n(\text{HCl reacted}) = \frac{m}{M}$
 $0,196 = \frac{m}{36,5}$
 $\therefore m(\text{HCl reacted}) = 7,154 \text{ g}$
 $40 \text{ g MgO} \checkmark \dots\dots\dots 73 \text{ g HCl} \checkmark$
 $x \text{ g MgO} \dots\dots\dots 7,154 \text{ g}$
 $\therefore x = 3,92 \text{ g}$
 $\% \text{ purity} = \frac{3,92}{4,5 \times 100} \checkmark$
 $= 87,11\% \checkmark$

(5)
[21]

QUESTION 2

2.1

2.1.1 An acid is a proton (H^+ ion) donor. ✓ (1)

2.1.2 It ionises to form 2 protons for each H_2SO_4 molecule. / It ionises to form 2 moles of H^+ ions. ✓

OR

It donates 2 H^+ ions per H_2SO_4 molecule. (1)

2.2

2.2.1 Amphiprotic substance / Ampholyte ✓ (1)

2.2.2 H_2CO_3 ✓ (1)

2.3

2.3.1

$$n(NaHCO_3) = \frac{m}{M} \checkmark$$

$$= \frac{27}{84} \checkmark$$

$$= 0,32 \text{ mol} \quad (0,0321485 \text{ mol})$$

$$n(H_2SO_4) = \frac{1}{2}n(NaHCO_3) = \frac{1}{2}(0,32) \checkmark = 0,16 \text{ mol} \quad (0,01607142 \text{ mol})$$

$$c = \frac{n}{V} \checkmark$$

$$6 = \frac{0,16}{V} \checkmark$$

$$\therefore V = 0,03 \text{ dm}^3 \checkmark \quad (30 \text{ cm}^3/0,027 \text{ dm}^3/27 \text{ cm}^3)$$

(6)

2.3.2

$$n_a(\text{initial/aanvanklik}) = n_a(\text{final/finaal})$$

$$c_a v_a(\text{initial/aanvanklik}) = c_a v_a(\text{final/finaal})$$

$$\therefore (6)v_a = (0,1)(1) \checkmark$$

$$\therefore v_a = 0,02 \text{ dm}^3 \checkmark \quad (20 \text{ cm}^3/0,0167 \text{ dm}^3/16,7 \text{ cm}^3)$$

(2)

2.3.3

Shows end point of titration. / Shows when neutralisation occurs. ✓ (1)



2.3.4

$$n_a(\text{initial/aanvanklik}) = c_a V_a \\ = (0,1)(25 \times 10^{-3}) \checkmark \\ = 2,5 \times 10^{-3} \text{ mol}$$

$$n_b(\text{reacted/gereageer}) = c_b V_b \\ = (0,1)(30 \times 10^{-3}) \checkmark \\ = 3 \times 10^{-3} \text{ mol}$$

$$\frac{n_a}{n_b} = \frac{1}{2}$$

$$\therefore n_a(\text{neutralised/geneutraliseer}) = \frac{1}{2}n_b = \frac{1}{2}(3 \times 10^{-3}) \checkmark = 1,5 \times 10^{-3} \text{ mol}$$

$$n_a(\text{left/oorgebly}) = n_a(\text{initial/aanvanklik}) - n_a(\text{neutralised/geneutraliseer}) \\ = 2,5 \times 10^{-3} - 1,5 \times 10^{-3} \checkmark \\ = 1 \times 10^{-3} \text{ mol}$$

$$c_a = \frac{n}{V} \\ = \frac{1 \times 10^{-3}}{(25 \times 10^{-3} + 30 \times 10^{-3})} \\ = 0,018 \text{ mol} \cdot \text{dm}^{-3}$$

$$\text{pH} = -\log[\text{H}_3\text{O}^+] \checkmark \\ = -\log(2 \times 0,018) \checkmark \\ = 1,44 \checkmark$$

(8)
[22]**QUESTION 3**

3.1

3.1.1 Diprotic ✓ (1)

3.1.2 H_2O ✓
 $(\text{COO})_2^{2-}$ ✓ (2)3.1.3 $\text{HC}_2\text{O}_4^- / \text{H}(\text{COO})_2^-$ ✓

It acts as base (in reaction I) and as acid (in reaction II). ✓ (2)

3.2 Ionises / dissociates incompletely / partially. ✓ (1)

3.3

OPTION 1	OPTION 2	Marking guidelines:
$c = \frac{m}{MV} \checkmark$ $\therefore 0,2 = \frac{m}{0,25 \times 90} \checkmark$ $\therefore m = 4,5 \text{ g} \checkmark$	$c = \frac{n}{V} \checkmark$ $0,2 = \frac{n}{0,25} \checkmark$ $n = 0,05 \text{ mol}$ $n = \frac{m}{M}$ $0,05 = \frac{m}{90} \checkmark$ $m = 4,5 \text{ g} \checkmark$	<ul style="list-style-type: none"> Any formula of: $c = \frac{m}{MV} / c = \frac{n}{V} / n = \frac{m}{M}$ Substitution of V as $0,25 \text{ dm}^3$. Substitution of $90 \text{ g} \cdot \text{mol}^{-1}$. Final answer: 4,5 g

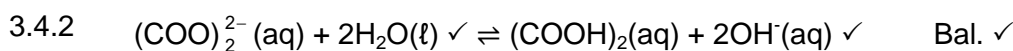
(4)

3.4

3.4.1

<p>OPTION 1</p> $\frac{c_a V_a}{c_b V_b} = \frac{n_a}{n_b} \checkmark$ $\frac{0,2 \times 25}{c_b \times 36} = \frac{1}{2} \checkmark$ $\therefore c_b = 0,28 \text{ mol} \cdot \text{dm}^{-3} \checkmark$	<p>Marking guidelines:</p> <ul style="list-style-type: none"> • Formula • Substitution of 0,2 x 25. • Use $V_b = 36 \text{ cm}^3$. • Use mol ratio 1:2. • Final answer: $0,28 \text{ mol} \cdot \text{dm}^{-3}$
<p>OPTION 2</p> $n((\text{COOH})_2) = cV \checkmark$ $= (0,2)(0,025) \checkmark$ $= 0,005 \text{ mol}$ $n(\text{NaOH}) = 2(0,005) \checkmark$ $= 0,01 \text{ mol}$ $c = \frac{n}{V}$ $= \frac{0,01}{0,036} \checkmark$ $= 0,28 \text{ mol} \cdot \text{dm}^{-3} \checkmark$	<p>Marking guidelines:</p> <ul style="list-style-type: none"> • Any ONE of formulae. • Substitution of 0,2 x 0,025. • Use mol ratio 1:2. • Use $V_b = 0,036 \text{ dm}^3$. • Final answer: $0,28 \text{ mol} \cdot \text{dm}^{-3}$

(5)



(3)

[18]

QUESTION 4

4.1

4.1.1 Hydrolysis \checkmark

(1)

4.1.2 Acidic \checkmark

Forms H_3O^+ ions during hydrolysis. \checkmark

OR

Salt of strong acid and weak base.

(2)

4.2

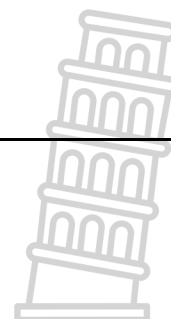
4.2.1 $n = cV \checkmark$
 $= (0,1)(0,1) \checkmark$
 $= 0,01 \text{ mol} \checkmark$

(3)

4.2.2 **POSITIVE MARKING FROM QUESTION 7.2.1.**

Marking criteria

- Substitute volume and concentration to calculate $n(\text{HCl}) \checkmark$
- Use mole ratio: $n(\text{NaOH}) = n(\text{HCl}) = 1:1 \checkmark$
- $n(\text{NaOH}) \times 4$ **OR** $V(\text{HCl}) \times 4$ **OR** $n(\text{HCl}) \times 4 \checkmark$
- Subtraction: $n(\text{NaOH}_{\text{initial}}) - n(\text{NaOH}_{\text{excess}}) \checkmark$
- Use mole ratio: $n(\text{NaOH}) = n(\text{NH}_4\text{Cl}) = 1:1 \checkmark$
- Substitute $53,5 \text{ g} \cdot \text{mol}^{-1}$ in $n = \frac{m}{M} \checkmark$
- Percentage calculation \checkmark
- Final answer: $0,11 \text{ g} - 0,21 \text{ g} \checkmark$



OPTION 1

$$n(\text{HCl}) = c_a V_a = (0,11)(14,55 \times 10^{-3}) \checkmark = 1,6 \times 10^{-3} \text{ mol}$$

$$n(\text{NaOH}) = n(\text{HCl}) = 1,6 \times 10^{-3} \text{ mol} \checkmark$$

$$n(\text{NaOH excess}) \text{ in } 100 \text{ cm}^3 = 1,6 \times 10^{-3} \times 4 \checkmark \\ = 6,4 \times 10^{-3} \text{ mol}$$

$$n(\text{NaOH reacted}) = 0,01 - 6,4 \times 10^{-3} \checkmark \\ = 3,6 \times 10^{-3} \text{ mol}$$

$$n(\text{NH}_4\text{Cl}) = n(\text{NaOH}) = 3,6 \times 10^{-3} \text{ mol} \checkmark \text{ (0,003598 mol)}$$

$$m(\text{NH}_4\text{Cl}) = nM \\ = (3,6 \times 10^{-3})(53,5) \checkmark \\ = 0,193 \text{ g}$$

$$92\% : 0,193 \text{ g} \\ 100\% : x$$

$$\therefore x = \frac{0,193 \times 100}{92} \checkmark \\ = 0,21 \text{ g} \checkmark$$

$$n(\text{NH}_4\text{Cl}) = 0,92 \frac{x}{53,5} \checkmark$$

$$\therefore 3,6 \times 10^{-3} = 0,92 \frac{x}{53,5}$$

$$\therefore x = 0,21 \text{ g} \checkmark$$

$$n(\text{NH}_4\text{Cl}) = \frac{m}{53,5} \checkmark$$

$$\therefore 3,6 \times 10^{-3} = \frac{m}{53,5} \\ n(\text{NH}_4\text{Cl}) = 0,192 \text{ g}$$

$$m(\text{fertiliser}): \\ m = \frac{0,192 \times 100}{92} \checkmark \\ = 0,21 \text{ g} \checkmark$$

OPTION 2

V(HCl) to neutralise 100 cm³ NaOH:

$$V(\text{HCl}) = 14,55 \times 4 \checkmark \\ = 58,2 \text{ cm}^3$$

$$n(\text{HCl}) = cV \\ = (0,11)(0,0582) \checkmark \\ = 0,006402 \text{ mol}$$

$$n(\text{NaOH}) = n(\text{HCl}) \\ = 0,006402 \text{ mol} \checkmark$$

$$n(\text{NaOH}_{\text{reacted}}): \\ n(\text{NaOH}) = 0,01 - 0,006402 \checkmark \\ = 0,003598 \text{ mol}$$

$$n(\text{NH}_4\text{Cl}) = n(\text{NaOH}) \\ = 0,003598 \text{ mol} \checkmark$$

$$m(\text{NH}_4\text{Cl}) = nM \\ = (0,003598)(53,5) \checkmark \\ = 0,192 \text{ g}$$

$$92\% : 0,192 \text{ g}$$

$$100\% : \frac{0,192 \times 100}{92} \checkmark = 0,21 \text{ g} \checkmark$$

OPTION 3

n(HCl) to neutralise 100 cm³ NaOH:

$$n(\text{HCl}) = cV \\ = (0,11)(0,01455 \times 4) \checkmark \\ = 0,006402 \text{ mol (6,4} \times 10^{-3} \text{ mol)}$$

$$n(\text{NaOH}_{\text{excess}}): \\ n(\text{NaOH}) = n(\text{HCl}) = 6,4 \times 10^{-3} \text{ mol} \checkmark$$

$$n(\text{NaOH}_{\text{reacted}}): \\ n(\text{NaOH}) = 0,01 - 0,006402 \checkmark \\ = 0,003598 \text{ mol}$$

$$n(\text{NH}_4\text{Cl}) = n(\text{NaOH}) \\ = 0,003598 \text{ mol} \checkmark$$

$$m(\text{NH}_4\text{Cl}) = nM \\ = (0,003598)(53,5) \checkmark \\ = 0,192 \text{ g}$$

$$92\% : 0,192 \text{ g}$$

$$100\% : 0,192 \times \frac{100}{92} \checkmark = 0,21 \text{ g} \checkmark$$

4.3

<p>OPTION 1 $[\text{OH}^-] = [\text{NaOH}] = 0,5 \text{ mol}\cdot\text{dm}^{-3}$ $K_w = [\text{H}_3\text{O}^+][\text{OH}^-]$ $1 \times 10^{-14} = [\text{H}_3\text{O}^+]0,5 \checkmark$ $\therefore [\text{H}_3\text{O}^+] = 2 \times 10^{-14} \text{ mol}\cdot\text{dm}^{-3}$ $\text{pH} = -\log[\text{H}^+] \checkmark$ $= -\log(2 \times 10^{-14}) \checkmark$ $= 13,7 \checkmark$</p>	<p>OPTION 2 $\text{pOH} = -\log[\text{OH}^-] \checkmark$ $= -\log(0,5) \checkmark$ $= 0,301$ $\text{pH} + \text{pOH} = 14$ $\text{pH} = 14 - 0,301 \checkmark$ $= 13,7 \checkmark \quad (13,699)$</p>
--	--

(4)
[18]

QUESTION 5

5.1 It is a proton donor. $\checkmark\checkmark$

(2)

5.2

5.2.1 $\text{CO}_3^{2-}(\text{aq}) \checkmark$

(1)

5.2.2 $\text{H}_2\text{CO}_3 + \text{H}_2\text{O} \checkmark \rightleftharpoons \text{HCO}_3^-(\text{aq}) + \text{H}_3\text{O}^+(\text{aq}) \checkmark$ Bal. \checkmark

(3)

5.2.3

<p>OPTION/OPSIE 1 $\text{pH} = -\log[\text{H}^+] \checkmark$ $3,4 = -\log[\text{H}^+] \checkmark$ $[\text{H}^+] = 10^{-3,4} / 3,98 \times 10^{-4} \text{ mol}\cdot\text{dm}^{-3}$ $[\text{H}^+][\text{OH}^-] = 10^{-14} \checkmark$ $\therefore [\text{OH}^-] = \frac{1 \times 10^{-14}}{3,98 \times 10^{-4}} \checkmark$ $= 2,51 \times 10^{-11} \text{ mol}\cdot\text{dm}^{-3} \checkmark$</p>	<p>OPTION/OPSIE 2 $\text{pH} + \text{pOH} = 14 \checkmark$ $3,4 + \text{pOH} = 14 \checkmark$ $\text{pOH} = 11,6$ $\text{pOH} = -\log[\text{OH}^-] \checkmark$ $11,6 = -\log[\text{OH}^-] \checkmark$ $[\text{OH}^-] = 10^{-11,6} / 2,51 \times 10^{-11} \text{ mol}\cdot\text{dm}^{-3} \checkmark$</p>
---	--

(5)

5.3

5.3.1 An acid that donates ONE proton / H^+ ion / H_3O^+ ion. \checkmark

(1)

OR

An acid of which ONE mol ionises to form ONE mol of protons / H^+ ions / H_3O^+ ions.

(1)

5.3.2

<p>OPTION/OPSIE 1 $\frac{c_a \times V_a}{c_b \times V_b} = \frac{n_a}{n_b} \checkmark$ $\frac{c_a \times 25}{0,1 \times 27,5} = \frac{1}{1} \checkmark$ $c_a = 0,11 \text{ mol}\cdot\text{dm}^{-3} \checkmark$</p>	<p>OPTION/OPSIE 2 $n(\text{NaOH}) = cV \checkmark$ $= 0,1 \times 0,0275 \checkmark$ $= 0,00275 \text{ mol}$ $n(\text{acid X}) = n(\text{NaOH})$ $= 0,00275 \text{ mol} \checkmark$ $c(\text{acid X}) = \frac{n}{V}$ $= \frac{2,75 \times 10^{-3}}{0,025} \checkmark$ $= 0,11 \text{ mol}\cdot\text{dm}^{-3} \checkmark$</p>
---	--

(5)

5.3.3 Weak \checkmark

The $[\text{H}^+]$ OR $[\text{H}_3\text{O}^+]$ is lower than the concentration of acid X. \checkmark

Therefore the acid is incompletely ionised. \checkmark

(3)

[20]

QUESTION 6

6.1

6.1.1 An acid is a proton/ H^+ donor. ✓✓**NOTE: NOT H_3O^+**

(2)

6.1.2 H_2O ✓ H_2CO_3 ✓

(2)

6.1.3 H_2O ✓ **OR** HCO_3^-

(1)

6.2

$$6.2.1 \quad n(HCl) = cV \quad \checkmark$$

$$= (0,1)(0,5) \quad \checkmark$$

$$= 0,05 \text{ mol}$$

$$n(NaHCO_3) = cV$$

$$= (0,25)(0,8) \quad \checkmark$$

$$= 0,2 \text{ mol}$$

$$n(NaHCO_3)_{\text{reacted}} = n(HCl)$$

$$= 0,05 \text{ mol} \quad \checkmark$$

$$n(NaHCO_3)_{\text{excess}} = \underline{0,2 - 0,05} \quad \checkmark$$

$$= 0,15 \text{ mol}$$

$$n(OH^-) = n(NaHCO_3) \quad \checkmark$$

$$= 0,15 \text{ mol}$$

$$c(OH^-) = \frac{n}{V}$$

$$= \frac{0,15}{1,3} \quad \checkmark$$

$$= 0,12 \text{ mol} \cdot \text{dm}^{-3} \quad \checkmark$$

(8)

6.2.2 **POSITIVE MARKING FROM QUESTION 6.2.1****OPTION 1**

$$K_w = [H_3O^+][OH^-] = 1 \times 10^{-14}$$

$$1 \times 10^{-14} = [H_3O^+](0,12) \quad \checkmark$$

$$[H_3O^+] = 8,33 \times 10^{-14} \text{ mol} \cdot \text{dm}^{-3}$$

$$pH = -\log [H_3O^+] \quad \checkmark$$

$$= -\log(8,33 \times 10^{-14}) \quad \checkmark$$

$$= 13,08 \quad \checkmark$$

OPTION 2

$$pOH = -\log[OH^-] \quad \checkmark$$

$$= -\log(0,12) \quad \checkmark$$

$$= 0,92$$

$$pH + pOH = 14$$

$$pH + 0,92 = 14 \quad \checkmark$$

$$pH = 13,08 \quad \checkmark$$

(4)

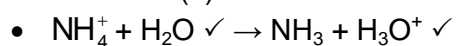
[17]**QUESTION 7**

7.1

7.1.1 Hydrolysis is the reaction of a salt with water. ✓✓

(2)

7.1.2 Smaller than (7) ✓



(3)

7.2

7.2.1

<p>Marking guidelines:</p> <ul style="list-style-type: none"> • Substitution of $98 \text{ g}\cdot\text{mol}^{-1}$. ✓ • Final answer: $0,08 \text{ mol}$ ✓ 	
<p>OPTION 1</p> $n = \frac{m}{M}$ $= \frac{7,35}{98}$ $= 0,08 \text{ mol} \checkmark \quad (0,075 \text{ mol})$	<p>OPTION 2</p> $98 \text{ g} \checkmark : 1 \text{ mol}$ $7,35 : 0,08 \text{ mol} \checkmark$
	<p>OPTION 3</p> $c = \frac{m}{MV}$ $= \frac{7,35}{98 \times 0,5}$ $= 0,15 \text{ mol}\cdot\text{dm}^{-3}$ $n = cV$ $= 0,15 \times 0,5$ $= 0,08 \text{ mol} \checkmark$

(2)

7.2.2

POSITIVE MARKING FROM QUESTION 7.2.1.

<p>OPTION 1</p> $\text{pH} = -\log[\text{H}_3\text{O}^+] \checkmark$ $1,3 \checkmark = -\log[\text{H}_3\text{O}^+]$ $[\text{H}_3\text{O}^+] = 0,05 \text{ mol}\cdot\text{dm}^{-3}$ $[\text{H}_2\text{SO}_4] = \frac{1}{2}[\text{H}_3\text{O}^+]$ $= \frac{1}{2} \times 0,05 \checkmark$ $= 0,025 \text{ mol}\cdot\text{dm}^{-3}$ $n(\text{H}_2\text{SO}_4)_{\text{excess}} = cV \checkmark$ $= 0,025 \times 0,5 \checkmark$ $= 0,0125 \text{ mol}$ $n(\text{H}_2\text{SO}_4)_{\text{react}} = 0,075 - 0,0125 \checkmark$ $= 0,0625 \text{ mol}$ $n(\text{NaOH}) = 2n(\text{H}_2\text{SO}_4)$ $= 2 \times 0,0625 \checkmark$ $= 0,125 \text{ mol}$ $n(\text{NaOH}) = \frac{m}{M}$ $\therefore 0,125 = \frac{m}{40} \checkmark$ $m = 5 \text{ g} \checkmark$	<p>Marking guidelines:</p> <ul style="list-style-type: none"> • Formula: $\text{pH} = -\log[\text{H}_3\text{O}^+] \checkmark$ • Substitution of $1,3 \checkmark$ • Use $[\text{H}_2\text{SO}_4] : [\text{H}_3\text{O}^+] = 1 : 2 \checkmark$ • Formula: $c = \frac{n}{V} \checkmark$ • Multiply by $0,5 \text{ dm}^3 \checkmark$ • Subtract $n_{\text{initial}} - n_{\text{excess}} \checkmark$ • Use $n(\text{NaOH}) : n(\text{H}_2\text{SO}_4) = 2:1 \checkmark$ • Substitution of $40 \text{ g}\cdot\text{mol}^{-1} \checkmark$ • Final answer: $m = 5 \text{ g} \checkmark$ • Range: $4,8 - 5,6 \text{ g}$
<p>Q7.2.1</p>	<p>OR</p> $1 \text{ mol} : 40 \text{ g} \checkmark$ $0,125 \text{ mol} : 5 \text{ g} \checkmark$

OPTION 2

Q7.2.1

$$[\text{H}_2\text{SO}_4]_{\text{in}} = \frac{n}{V} \checkmark$$

$$= \frac{0,075}{0,5} \checkmark$$

$$= 0,15 \text{ mol}\cdot\text{dm}^{-3}$$

$$[\text{H}_3\text{O}^+]_{\text{initially}} = 2[\text{H}_2\text{SO}_4]$$

$$= 2 \times 0,15 \checkmark$$

$$= 0,3 \text{ mol}\cdot\text{dm}^{-3}$$

$$\text{pH} = -\log[\text{H}_3\text{O}^+] \checkmark$$

$$1,3 \checkmark = -\log[\text{H}_3\text{O}^+]$$

$$[\text{H}_3\text{O}^+] = 0,05 \text{ mol}\cdot\text{dm}^{-3}$$

$$[\text{H}_3\text{O}^+]_{\text{react}} = 0,3 - 0,05 \checkmark$$

$$= 0,25 \text{ mol}\cdot\text{dm}^{-3}$$

$$[\text{H}_2\text{SO}_4]_{\text{react}} = \frac{1}{2}[\text{H}_3\text{O}^+]$$

$$= \frac{1}{2} \times 0,25$$

$$= 0,125 \text{ mol}\cdot\text{dm}^{-3} \text{ (0,14)}$$

OR

$$n(\text{H}_2\text{SO}_4)_{\text{react}} = cV$$

$$= (0,125)(0,5)$$

$$= 0,0625 \text{ mol}$$

$$n(\text{NaOH}) = 2n(\text{H}_2\text{SO}_4)$$

$$= 2 \times 0,0625 \checkmark$$

$$= 0,125 \text{ mol}$$

$$n(\text{NaOH}) = \frac{m}{M}$$

$$0,125 = \frac{m}{40} \checkmark$$

$$\therefore m = 5 \text{ g} \checkmark$$

Marking guidelines:

- Formula: $c = \frac{n}{V} \checkmark$
 - Divide by $0,5 \text{ dm}^3 \checkmark$
 - Use $[\text{H}_3\text{O}^+] : [\text{H}_2\text{SO}_4] = 2:1 \checkmark$
 - Formula: $\text{pH} = -\log[\text{H}_3\text{O}^+] \checkmark$
 - Substitution of $1,3 \checkmark$
 - Subtract $[\text{H}_3\text{O}^+]_{\text{initial}} - [\text{H}_3\text{O}^+]_{\text{excess}} \checkmark$
 - Use $n(\text{NaOH}) : n(\text{H}_2\text{SO}_4) = 2:1 \checkmark$
- OR**
- Use $[\text{H}_2\text{SO}_4] : [\text{NaOH}] = 1 : 2 \checkmark$
 - Substitution of $40 \text{ g}\cdot\text{mol}^{-1} \checkmark$
 - Final answer: $m = 5 \text{ g} \checkmark$
- Range: $4,8 - 5,6 \text{ g}$

(9)
[16]

GALVANIC CELLS

QUESTION 1

- 1.1
- Pressure: 1 atmosphere (atm) / 101,3 kPa / $1,013 \times 10^5$ Pa ✓
 - Temperature: 25 °C / 298 K ✓
- (2)
- 1.2
- Platinum is inert / does not react with the H^+ ions OR acid. ✓
 - Platinum is a conductor (of electricity). ✓
- (2)
- 1.3
- 1.3.1 Salt bridge ✓ (1)
- 1.3.2 -0,31 V ✓ (1)
- 1.3.3 $2H^+ + 2e^- \rightarrow H_2$ ✓✓ (2)
- 1.4
- 1.4.1 **POSITIVE MARKING FROM QUESTION 1.3.2.**
- $E_{\text{cell}}^{\theta} = E_{\text{reduction}}^{\theta} - E_{\text{oxidation}}^{\theta}$ ✓
- $2,05 \checkmark = -0,31 \checkmark - E_{M/M^{2+}}^{\theta}$
- $E_{M/M^{2+}}^{\theta} = -2,36$ (V) ✓
- M is magnesium/ Mg. ✓ (5)
- 1.4.2 Exothermic ✓ (1)
- 1.5 The cell reaction reaches equilibrium. ✓ (1)
- [15]**

QUESTION 2

- 2.1 A substance that is being reduced. ✓✓
- OR**
- A substance that gains/accepts electrons.
- OR**
- A substance whose oxidation number decreases. (2)
- 2.2 Ag^+ is a stronger oxidising agent ✓ than Cu^{2+} ✓ and will oxidise Cu ✓ to (blue) Cu^{2+} ions. ✓
- OR**
- Cu^{2+} is a weaker oxidising agent ✓ than Ag^+ ✓ and Cu will be oxidised ✓ to Cu^{2+} ions. ✓ (4)
- 2.3 Chemical energy to electrical energy ✓ (1)
- 2.4 A ✓ (1)
- 2.5 $E_{\text{cell}}^{\theta} = E_{\text{reduction}}^{\theta} - E_{\text{oxidation}}^{\theta}$ ✓
- $= +0,80 \checkmark - 0,34 \checkmark$
- $= +0,46$ V ✓ (4)
- 2.6 $Cu + 2Ag^+(aq) \checkmark \rightarrow Cu^{2+}(aq) + 2Ag(s) \checkmark$ Balancing ✓ (3)
- 2.7 Remains the same ✓ (1)
- [16]**

QUESTION 3

3.1 Redox reaction ✓ (1)

3.2 P ✓
Negative electrode. / Mg is a stronger reducing agent/is oxidized/release electrons.
✓ (2)

3.3
3.3.1 (Temperature): 25 °C/298 K ✓
(Concentration): 1 mol·dm⁻³ ✓ (2)

3.3.2 Mg(s) | Mg²⁺(aq) ✓ || Pb²⁺(aq) | Pb(s) ✓
OR
Mg | Mg²⁺ || Pb²⁺ | Pb (3)

3.3.3 Pb²⁺ / Pb(NO₃)₂ / lead(II) ions ✓ (1)

3.4 $E_{\text{cell}}^{\ominus} = E_{\text{reduction}}^{\ominus} - E_{\text{oxidation}}^{\ominus}$ ✓
= -0,13 ✓ - (-2,36) ✓
= 2,23 V ✓ (4)

3.5
3.5.1 Remains the same ✓ (1)

3.5.2 Increases ✓ (1)

[15]**QUESTION 4**

4.1 Temperature: 25 °C / 298 K ✓
Pressure: 101,3 kPa / 1,013 x 10⁵ Pa / 1 atm / 100 kPa ✓
Concentration: 1 mol·dm⁻³ ✓ (3)

4.2
4.2.1 Cd(s) / Cadmium / Cd | Cd²⁺ / Cd²⁺ | Cd ✓ (1)

4.2.2 $E_{\text{cell}}^{\ominus} = E_{\text{cathode}}^{\ominus} - E_{\text{anode}}^{\ominus}$ ✓
0,13 ✓ = $E_{\text{cathode}}^{\ominus} - (-0,40)$ ✓
 $E_{\text{cathode}}^{\ominus} = -0,27$ V ✓
Q is Ni / nickel ✓ (5)

4.3
4.3.1 Cd(s) → Cd²⁺(aq) + 2e⁻ ✓ ✓ (2)

4.3.2 Pt / Platinum ✓ (1)

4.4 **OPTION 1**

Compare Q^{2+} & Cd^{2+}	Q^{2+} is reduced / Cd is oxidised and therefore <u>Q^{2+} is a stronger oxidising agent than Cd^{2+}.</u>	✓
Compare R_2 & Cd^{2+}	R_2 is reduced / Cd is oxidised and therefore <u>R_2 is a stronger oxidising agent than Cd^{2+}.</u> ✓	✓
Compare R_2 & Q^{2+}	The cell potential of combination II is higher than that of combination I, therefore <u>R_2 is a stronger oxidising agent than Q^{2+}.</u>	✓
Final answer	$Cd^{2+}; Q^{2+}; R_2$ OR $Cd^{2+}; Ni^{2+}; Cl_2$	✓

OPTION 2

- The reduction potential of $Cl | Cl_2 = 1,36 V$ ✓ because the cell potential of combination II is 1,76 V and the reduction potential of $Cd | Cd^{2+}$ is 0,4 V.
OR
 R_2 is Cl_2 because the cell potential of combination II is 1,76 V and the reduction potential of $Cd | Cd^{2+}$ is 0,4 V.
- $Cd | Cd^{2+}$ has the lowest reduction potential (-0,4 V) and therefore Cd^{2+} is the weakest oxidising agent. ✓
- $Cl | Cl_2$ has the highest reduction potential and therefore Cl_2 is the strongest oxidising agent. ✓
- Final answer: $Cd^{2+}; Q^{2+}; R_2$ ✓ OR $Cd^{2+}; Ni^{2+}; Cl_2$

(4)
[16]**QUESTION 5**

5.1 B ✓

(1)

5.2

5.2.1 $Cl_2(g) + 2e^- \rightarrow 2Cl^-(aq)$ ✓✓

(2)

5.2.2 Cl_2 / chlorine ✓

(1)

5.3 **OPTION 1/OPSIE 1**

$$E_{cell}^{\ominus} = E_{cathode}^{\ominus} - E_{anode}^{\ominus} \quad \checkmark$$

$$= 1,36 \checkmark - (-2,36) \checkmark$$

$$= 3,72 V \checkmark$$

(4)

- 5.4
- The Mg electrode becomes smaller. / The mass of the Mg electrode decreases. / The Mg electrode being corroded. ✓
 - Magnesium is oxidised. / $Mg \rightarrow Mg^{2+} + 2e^-$ ✓

(2)
[10]**QUESTION 6**6.1 Electrons are transferred. ✓

OR

The oxidation number of Mg / H changes.

OR

Mg is oxidised / H^+ is reduced.

(1)

6.2 H^+ ions / HCl / $H^+(aq)$ / $HCl(aq)$ ✓

(1)

6.3 Ag is a weaker reducing agent ✓ than H_2 and will not be oxidised ✓ to Ag^+ ✓

OR

 H_2 is a stronger reducing agent ✓ than Ag and will be oxidised ✓ to H^+ . ✓

(3)

- 6.4 Electrode / Conductor of electrons in hydrogen half-cell ✓ (1)
- 6.5
- 6.5.1 Chemical energy to electrical energy ✓ (1)
- 6.5.2 Provides path for movement of ions./Completes the circuit./Ensures electrical neutrality in cell. ✓ (1)
- 6.5.3 $2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$ ✓✓ (2)
- 6.5.4 $\text{Mg(s)} \mid \text{Mg}^{2+}(\text{aq}) \parallel \text{H}^+(\text{aq}) \mid \text{H}_2(\text{g}) \mid \text{Pt}$
- OR**
- $\text{Mg(s)} \mid \text{Mg}^{2+}(1 \text{ mol}\cdot\text{dm}^{-3}) \parallel \text{H}^+(1 \text{ mol}\cdot\text{dm}^{-3}) \mid \text{H}_2(\text{g}) \mid \text{Pt}$ (3)
- 6.6 $E_{\text{cell}}^{\theta} = E_{\text{reduction}}^{\theta} - E_{\text{oxidation}}^{\theta}$ ✓
 $= 0,00 \text{ ✓} - (-2,36) \text{ ✓}$
 $= 2,36 \text{ V } \text{ ✓}$ (4)
- 6.7 Increases ✓ (1)
- [18]**

QUESTION 7

- 7.1
- 7.1.1 AgNO_3 / Silver nitrate ✓ (1)
- 7.1.2 $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ ✓✓ (2)
- 7.1.3 $\text{Ni} + 2\text{Ag}^+ \text{ ✓} \rightarrow \text{Ni}^{2+} + 2\text{Ag} \text{ ✓}$ Bal ✓
- OR**
- $\text{Ni} + 2 \text{AgNO}_3 \rightarrow \text{Ni}(\text{NO}_3)_2 + 2\text{Ag}$ (3)
- 7.2
- 7.2.1 Ni ✓
 Ni is a stronger reducing agent. / Ni is the anode. / Ni loses electrons. / Ni is oxidised. ✓ (2)
- 7.2.2 $\text{Ni (s)} \mid \text{Ni}^{2+}(\text{aq}) \parallel \text{Ag}^+(\text{aq}) \mid \text{Ag(s)}$
- OR**
- $\text{Ni (s)} \mid \text{Ni}^{2+}(1 \text{ mol}\cdot\text{dm}^{-3}) \parallel \text{Ag}^+(1 \text{ mol}\cdot\text{dm}^{-3}) \mid \text{Ag(s)}$ (3)
- 7.2.3 $E_{\text{cell}}^{\theta} = E_{\text{reduction}}^{\theta} - E_{\text{oxidation}}^{\theta}$ ✓
 $= 0,80 \text{ ✓} - (-0,27) \text{ ✓}$
 $= 1,07 \text{ V } \text{ ✓}$ (4)
- 7.2.4 Increases ✓ (1)
- [16]**

ELECTROLYTIC CELLS

QUESTION 1

- 1.1 Electrolytic ✓ (1)
- 1.2 Q ✓ & T ✓
- $\text{Cu}^{2+} + 2\text{e}^- \rightarrow \text{Cu}$ ✓✓ (4)
- 1.3
- 1.3.1 Cl_2 / chlorine gas ✓ (1)
- 1.3.2 Cu^{2+} ions / copper(II) ions / CuCl_2 / copper(II) chloride ✓ (1)
- 1.4 Cu is a stronger reducing agent ✓ than Cl^- ions ✓ and Cu will be oxidised ✓ to Cu^{2+} .
OR
 Cl^- ions is a weaker reducing agent ✓ than Cu ✓ and Cu will be oxidised ✓ to Cu^{2+} . (3)
[10]

QUESTION 2

- 2.1 A solution that conducts electricity through the movement of ions. ✓✓ (2)
- 2.2
- 2.2.1 $2\text{H}_2\text{O} + 2\text{e}^- \rightarrow \text{H}_2 + 2\text{OH}^-$ ✓✓ (2)
- 2.2.2 Chlorine gas / Cl_2 ✓ (1)
- 2.3 H_2O is a stronger oxidising agent ✓ than Na^+ and will be reduced ✓ to H_2 . (2)
[7]

QUESTION 3

- 3.1 A solution that conducts electricity through the movement of ions. ✓✓ (2)
- 3.2 Plastic is a non-conductor of electricity. / Graphite is a conductor. ✓ (1)
- 3.3
- 3.3.1 $\text{Ni}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Ni}(\text{s})$ ✓✓ (2)
- 3.3.2 Ni / nickel ✓
 Ni is oxidised. ✓
OR Ni loses electrons.
OR Ni is the anode.
OR Ni is the positive electrode. (2)
- 3.4 Ring ✓
 Reduction takes place at the cathode. ✓
OR
 Negative electrode. (2)
- 3.5 Decreases ✓
 Ni^{2+} ions from the electrolyte will be reduced (to Ni). ✓
OR
 Ni^{2+} changes to Ni (2)
[11]

QUESTION 4

- 4.1 **ANY ONE:**
- The chemical process in which electrical energy is converted to chemical energy. ✓✓
 - The use of electrical energy to produce a chemical change.
 - The process during which an electric current passes through a solution / molten ionic compound. (2)

- 4.2 **ANY ONE:**
- To keep the polarity of the electrodes constant. ✓
 - To prevent the anode and cathode from swapping.
 - DC provides a one way flow of electrons ensuring that the same chemical reaction occurs all the time at the same electrodes. (1)



- 4.4
- Cu^{2+} is a stronger oxidising agent ✓ than Zn^{2+} . ✓
 - Cu^{2+} will be reduced to Cu. ✓

OR

- Zn is a stronger reducing agent than Cu.
- Cu^{2+} will be reduced to Cu. (3)

QUESTION 5

5.1 Electrolytic cell ✓ (1)

5.2 The substance/species which loses electrons. ✓✓ (2)

5.3 P ✓ (1)



5.5 A ✓
 Cl^- ions move to the positive electrode / anode where they are oxidised to Cl_2 . ✓✓ (3)

[12]

QUESTION 6

6.1
 6.1.1 Electrolyte ✓ (1)

6.1.2 Electrolytic cell ✓ (1)

6.2 A to B ✓ (1)

6.3
 6.3.1 B ✓ (1)

6.3.2 A ✓ (1)

6.4 Decreases ✓

Copper (Cu) is oxidised to Cu^{2+} / Oxidation takes place at A. ✓ (2)

[7]

QUESTION 7

- 7.1 Endothermic ✓ (1)
- 7.2 Anode ✓
Connected to the positive terminal of the battery. ✓ (2)
- 7.3 Chlorine gas / Cl_2 ✓ (1)
- 7.3.1 Chlorine gas / Cl_2 ✓ (1)
- 7.3.2 Hydrogen gas / H_2 ✓ (1)
- 7.3.3 $2\text{H}_2\text{O}(\ell) + 2\text{e}^- \rightarrow \text{H}_2(\text{g}) + 2\text{OH}^-(\text{aq})$ ✓✓ (2)
- 7.4 Basic / alkaline ✓
 OH^- ions form / NaOH forms ✓ (2)
- [9]**



FERTILISERS

QUESTION 1

- 1.1
 1.1.1 Nitrogen / N₂ ✓
 Hydrogen / H₂ ✓ (2)
- 1.1.2 NH₃ + HNO₃ ✓ → NH₄NO₃ ✓ Bal. ✓ (3)

1.2 **Marking criteria:**

- Use ratio: $\frac{3}{9}$ ✓
- x 20 kg ✓
- x 36 / 36 % ✓
- Final answer: 2,4 kg ✓

<p>OPTION 1:</p> $\% N = \frac{3}{9} \checkmark (x 36) \checkmark$ $= 12 \%$ $\therefore m(N) : \frac{12}{100} (x 20 \text{ kg})$ $= 2,4 \text{ kg} \checkmark$	<p>OPTION 2:</p> <p>m(nutrients):</p> $\frac{36}{100} \checkmark (x 20) = 7,2 \text{ kg}$ $\therefore m(N) = \frac{3}{9} \checkmark x 7,2$ $= 2,4 \text{ kg} \checkmark$	<p>OPTION 3:</p> <p>m(N):</p> $\frac{3 \checkmark}{9} \times (x 20) (x \frac{36 \checkmark}{100}) = 2,4 \text{ kg}$
--	---	--

(4)
[9]

QUESTION 2

- 2.1 Contact process ✓ (1)
- 2.2
 2.2.1 Vanadium pentoxide / Vanadium(V) oxide / V₂O₅ ✓ (1)
- 2.2.2 H₂S₂O₇(l) + H₂O(l) ✓ → 2H₂SO₄(l) ✓ Bal. ✓ (3)
- 2.3
 2.3.1 H₂SO₄ + 2NH₃ ✓ → (NH₄)₂SO₄ ✓ Bal. ✓ (3)
- 2.3.2 Ammonium sulphate ✓ (1)
- 2.4
 2.4.1 Total percentage of fertiliser ✓ (1)
- 2.4.2
- Mass of fertiliser in P/Massa kunsmis in P: $\frac{25}{100} \times 50 = 12,5 \text{ kg}$

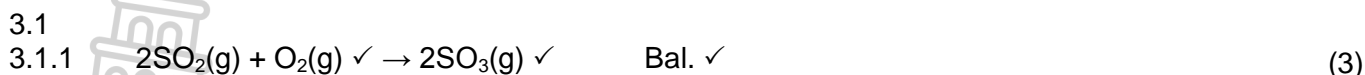
Mass of fertiliser in Q/Massa kunsmis in Q: $\frac{20}{100} \times 50 = 10 \text{ kg}$

}

✓
- ↓
- Amount of potassium in P/Massa kalium in P: $\frac{3}{10} \times 12,5 = 3,75 \text{ kg} \checkmark$
- Amount of potassium in Q/Massa kalium in Q: $\frac{4}{8} \times 10 = 5 \text{ kg} \checkmark$
- Fertiliser Q has more potassium per mass than fertiliser P. ✓
 Kunsmis Q het meer kalium per massa as kunsmis P.

(4)
[14]

QUESTION 3



3.1.2 Catalyst \checkmark
OR
Increase the reaction rate (1)

3.2 Exothermic \checkmark
The temperature increases \checkmark (2)

- 3.3
- An exothermic reaction is favoured by a decrease in temperature. \checkmark
 - The forward reaction is favoured. \checkmark
 - Higher yield (of SO_3). \checkmark
- OR**
- An endothermic reaction is favoured by an increase in temperature. \checkmark
 - The reverse reaction is favoured. \checkmark
 - Lower yield (of SO_3). \checkmark (3)

3.4
3.4.1 $\text{H}_2\text{S}_2\text{O}_7 \checkmark$ (1)

3.4.2 A mist will form (which is difficult to collect). \checkmark
OR
The reaction is too exothermic. (1)



QUESTION 4

4.1
4.1.1 Haber process \checkmark (1)



4.1.3 Air \checkmark (1)

4.2
4.2.1 40% \checkmark (1)

- 4.2.2
- High yield \checkmark
 - High rate due to higher concentration. \checkmark (2)

4.2.3 Low reaction rate \checkmark (1)

OPTION 1	OPTION 2
$\% \text{ N in } \text{NH}_4\text{NO}_3 = \frac{28}{80} \checkmark \times 100$ $= 35\%$ <p>m(N) in 50 kg:</p> $\frac{35}{100} \times 50 \checkmark = 17,5 \text{ kg } \checkmark$	$m(\text{N in } \text{NH}_4\text{NO}_3) = \frac{28}{80} \checkmark \times 50 \checkmark$ $= 17,5 \text{ kg } \checkmark$
	(3)

QUESTION 5

5.1 Ostwald process ✓ (1)

5.2 NO / nitrogen monoxide ✓
Water / H₂O ✓ (2)

5.3 NH₃ + HNO₃ ✓ → NH₄NO₃ ✓ bal ✓ (3)

<p>OPTION 1/OPTION 1</p> $n(\text{NH}_3) = \frac{m}{M}$ $= \frac{6,8 \times 10^7}{17}$ $= 4 \times 10^6 \text{ mol}$ <p style="text-align: center;">↓</p> $n(\text{NH}_4\text{NO}_3) = n(\text{NH}_3)$ $= 4 \times 10^6 \text{ mol}$ $m(\text{NH}_4\text{NO}_3) = nM$ $= (4 \times 10^6)(80) \checkmark$ $= 3,2 \times 10^8 \text{ g}$ $= 3,2 \times 10^5 \text{ kg} \checkmark$	<p>OPTION 2/OPSIE 2</p> $m(\text{NH}_4\text{NO}_3) = \frac{6,8 \times 10^4}{17} \times 80 \checkmark \checkmark$ $= 3,2 \times 10^5 \text{ kg} \checkmark$ <hr/> <p>OPTION 3/OPSIE 3</p> <p>17 g ✓ NH₃ forms/vorm 80 g ✓ NH₄NO₃ 6,8 x 10⁴ kg forms/vorm x g NH₄NO₃</p> $x = 6,8 \times 10^4 \times \frac{80}{17}$ $= 3,2 \times 10^5 \text{ kg} \checkmark$
---	--

(3)

5.5 To make a NPK fertiliser . / To make fertilisers which contain all three primary nutrients. ✓ (1)
[10]

QUESTION 6

6.1
6.1.1 Air ✓ (1)

6.1.2 Natural gas / methane / oil / coal ✓ (1)

6.1.3 Sulphur / iron pyrite / iron sulphide ✓ (1)

6.2
6.2.1 Haber process ✓ (1)

6.2.2 Ammonia ✓ (1)

6.2.3 H₂SO₄ ✓ (1)

6.2.4 SO₃ + H₂SO₄ ✓ → H₂S₂O₇ ✓ Bal. ✓ (3)

6.3
6.3.1 $\left. \begin{aligned} \%N[\text{NH}_4\text{NO}_3] &= \frac{28}{80} \checkmark \times 100 = 35\% \\ \%N[(\text{NH}_4)_2\text{SO}_4] &= \frac{28}{132} \checkmark \times 100 = 21,21\% \end{aligned} \right\} \checkmark$

Ammonium nitrate (has the highest percentage of nitrogen). ✓ (4)

6.3.2 Ostwald process ✓ (1)
[14]

QUESTION 7

7.1

7.1.1 Haber process ✓ (1)

7.1.2 Contact process / Catalytic oxidation of SO₂ ✓ (1)7.1.3 Sulphur trioxide / SO₃ ✓ (1)7.1.4 SO₃ + H₂SO₄ ✓ → H₂S₂O₇ ✓ Bal. ✓ (3)7.1.5 H₂SO₄ ✓ + 2NH₃ ✓ → (NH₄)₂SO₄ ✓ Bal. ✓ (4)

7.2

Marking guidelines:

- Calculate the mass of fertiliser.
- Add %N and %P OR mass N and mass P.
- Subtraction: 100 – (%N + %P)
OR m(fertiliser) – [m(N) + m(P)]
OR %fertiliser – [%N + %P]
- Final answer: 8:1:5

OPTION 1

$$m(\text{fertiliser}) = \frac{36}{100} \times 20 \checkmark$$

$$= 7,2 \text{ kg}$$

$$\%N = \frac{4,11}{7,2} \times 100$$

$$= 57,08\%$$

$$\%P = \frac{0,51}{7,2} \times 100$$

$$= 7,08\%$$

$$\%K = \frac{100 - \checkmark}{\checkmark} (57,08 + 7,08) \checkmark$$

$$= 35,84\%$$

$$57,08 : 7,08 : 35,84$$

$$8 : 1 : 5 \checkmark$$

OPTION 2

$$m(\text{fertiliser}) = \frac{36}{100} \times 20 \checkmark$$

$$= 7,2 \text{ kg}$$

$$m(K) = \frac{7,2 - \checkmark}{\checkmark} (4,11 + 0,51) \checkmark$$

$$= 2,58 \text{ kg}$$

$$4,11 : 0,51 : 2,58$$

$$8 : 1 : 5 \checkmark$$

OPTION 3

$$\%N = \frac{4,11}{20} \times 100 = 20,55\%$$

$$\%P = \frac{0,51}{20} \times 100 = 2,55\%$$

$$\%K = \frac{36 - \checkmark}{\checkmark} (20,55 + 2,55) \checkmark = 12,9\%$$

$$20,55 : 2,55 : 12,9$$

$$8 : 1 : 5 \checkmark$$

(4)
[14]**BIBLIOGRAPHY**

Department of Basic Education, *National Senior Certificate Physical Sciences Question Papers*, 2014 – 2016, Pretoria